

## CHAPTER 4

### THE DETERMINATION OF THE STRUCTURE

The structure of  $Nb_3As$  was calculated by using various computer programs as shown in Table 4.0.

Table 4.0 The crystallographic calculations were performed on a CDC 3100 computer using these programs.

Program	Calculation
FFCRL	Film factors.
LPARL	Correction for Lorentz-polarization effects and absorption for cylindrical or spherical crystals.
XSFRL	Interpolation of the X-ray scattering factors.
SFCRL	Calculation of structure factors, the temperature factor may be isotropic and or anisotropic and the anomalous dispersion correction terms may be included.
EXPRL	Expansion of reflection data for Fourier summations.
FORRL	Fourier summation.
LSQRL	Full matrix least-squares refinements.
DISTAN	Calculation of interatomic distances and angles.

These programs were obtained from the Institute of Chemistry, University of Uppsala, Sweden and were modified to be used with the CDC 3100 computer at Bangkok Data Center Co., Ltd. by Professor R. Liminga of the University of Uppsala and the Crystallography Group, Faculty of Science, Chulalongkorn University.

#### 4.1 Structure factor calculation

The collected observed intensity data as previously described constituted the material from which the crystal structure was derived.

The structure factor,  $|F_{hkl}|$ , is the most important quantity derived from these intensities. The relationship between the observed structure factor,  $|F_o|$ , and the intensity depends on the Lorentz-polarization factor,

$$|F_o|_{hkl} = \sqrt{\frac{KI}{Lp}}_{hkl} \quad \dots (4.1.1)$$

where  $L$  is the Lorentz-factor,  
 $p$  is the polarization factor,  
 $K$  is a constant for any given set of measurements.

The results of  $|F_{\text{relative}}|$ 's,  $|F_{\text{rel}}|$ , are defined by

$$\begin{aligned} |F_{\text{rel}}| &= k' |F_o| \\ &= \sqrt{\frac{I_{hkl}}{Lp}} \quad \dots (4.1.2) \end{aligned}$$

In the course of this crystal structure analysis, structure factors were computed many times by using the program adapted from those used in the University of Uppsala, Sweden.

Some essential input of the computer program consists of raw intensity data from the five layers of the Weissenberg photographs, each reflection intensity being identified by its indices,  $h$ ,  $k$  and  $l$ , and the cell parameters, and a list of atomic

positions, which needed to be estimated first.

Before starting the computation of the structure factors, it was necessary to derive some constant functions for correcting the raw intensity data and we had to know the atomic positions in the unit cell of  $\text{Nb}_3\text{As}$ . These atomic positions could be obtained from Patterson functions and included in the  $F_0$  synthesis.

The input of the LPARL program were the observed intensities, the cell parameters and the absorption correction values of the crystal,  $A^*$ , and the output of this program was data corrected for the Lorentz-polarization factor.

The input data of the scattering factor values of both Nb and As at the appropriate  $\frac{\sin \theta}{\lambda}$  values were the atomic scattering data at fixed interval as shown in Table 4.1.1. The input information is available in tabular form in the International Tables for X-ray Crystallography, Vol. III.

The anomalous dispersion is the effect when the wavelength of the incident beam lies near an absorption edge in the scattering element. This effect causes the scattering factor to be complex, it can be represented by

$$f_0^{\text{anom.}} = f_0 + \Delta f' + i \Delta f'' \dots (4.1.3)$$

where  $f_0$  is the normal scattering factor,  
 $\Delta f'$  is a real correction term and  
 $\Delta f''$  is the imaginary component.

Table 4.1.1 The scattering factor of Nb and As atom, for MoK<sub>α</sub> radiation,  $\lambda_{K_{\alpha}}^{\circ} = 0.71068 \text{ \AA}$ .

$\frac{\sin \theta}{\lambda}$	$f_{\text{Nb}}$	$f_{\text{As}}$
0.00	41.00	33.00
0.05	40.04	32.32
0.10	37.82	30.61
0.15	35.23	28.53
0.20	32.64	26.49
0.25	30.19	24.63
0.30	27.95	22.95
0.35	25.97	21.39
0.40	24.24	19.93
0.45	22.74	18.54
0.50	21.43	17.21
0.55	20.28	15.95
0.60	19.23	14.76
0.65	18.26	13.66
0.70	17.34	12.65
0.75	16.47	11.73
0.80	15.62	10.91
0.85	14.80	10.18
0.90	14.00	9.54
0.95	13.24	8.97
1.00	12.51	8.48
1.05	11.82	8.05
1.10	11.17	7.68
1.15	10.56	7.35
1.20	10.00	7.07

The correction terms  $\Delta f'$  and  $\Delta f''$  which arise from anomalous scattering are independent of  $\sin\theta$ .

For Nb atom,  $\Delta f' = -2.19$  and  $\Delta f'' = 0.86$ .

For As atom,  $\Delta f' = 0.12$  and  $\Delta f'' = 2.17$ .

For structure factor calculation, the space group  $P4_2/n$  is classified as centrosymmetric.

The correction for thermal motion was applied to structure factor calculations. All of the atoms were assumed to be vibrating with the same amplitude about their rest positions, i.e. the vibration is isotropic. For the first structure factor calculation, we set the values of B's for Nb and As atoms as 0.25, where B is thermal parameter related to the magnitude of vibration.

The structure factor program usually puts out some summary information about the agreement between the observed and calculated data. The general agreement in magnitude is given by some sort of rescale factor.

The linear rescale factor is

$$\text{LRS} = \frac{\sum |F_c|}{\sum |F_o|} \quad \dots (4.1.4)$$

This rescale factors for individual levels were then used as the input scale factors for the corresponding levels in the next cycle of calculations. The over-all scale should be adjusted to bring  $\sum |F_o|$  and  $\sum |F_c|$  into agreement.

## 4.2 Determination of atomic positions

It is not possible to determine atomic positions directly from the intensities of X-ray diffraction. The interatomic vectors were determined from the Patterson synthesis.

$\text{Nb}_3\text{As}$  is a tetragonal crystal, space group  $P4_2/n$  whose co-ordinates of equivalent positions are :

$$\left. \begin{array}{l}
 x, y, z \\
 \frac{1-x}{2}, \frac{1-y}{2}, z \\
 \bar{y}, \frac{1+x}{2}, \frac{1+z}{2} \\
 \frac{1+y}{2}, \bar{x}, \frac{1+z}{2} \\
 \bar{x}, \bar{y}, \bar{z} \\
 \frac{1+x}{2}, \frac{1+y}{2}, \bar{z} \\
 y, \frac{1-x}{2}, \frac{1-z}{2} \\
 \frac{1-y}{2}, x, \frac{1-z}{2}
 \end{array} \right\} \text{for origin at } \bar{1}$$

and

$$\left. \begin{array}{l}
 x, y, z \\
 \bar{x}, \bar{y}, z \\
 \frac{1+x}{2}, \frac{1+y}{2}, \frac{1-z}{2} \\
 \frac{1-x}{2}, \frac{1-y}{2}, \frac{1-z}{2} \\
 \bar{y}, x, \bar{z} \\
 y, \bar{x}, \bar{z} \\
 \frac{1-y}{2}, \frac{1+x}{2}, \frac{1+z}{2} \\
 \frac{1+y}{2}, \frac{1-x}{2}, \frac{1+z}{2}
 \end{array} \right\} \text{for origin at } \bar{4}$$

The interatomic vectors of this space group, for the origin at  $\bar{1}$ , are shown in Table 4.2.1.

The Patterson function was calculated three dimensionally as projections about the c-axis.

The cell parameters of  $\text{Nb}_3\text{As}$  are  $a = 10.294$ ,  $b = 10.294$ ,  $c = 5.1958 \text{ \AA}$ , and there are 24 Nb atoms and 8 As atoms in each unit cell.

The Patterson function was projected along the c-axis at an interval of ca.  $0.25 \text{ \AA}$  from  $W = 0.00$  to  $0.50$ . All of the significant peaks were found only on three sections at  $W = 0$ ,  $0.25$  and  $0.50$  as shown in Fig.4.2.1. Therefore it was concluded that the niobium atoms must lie in the planes with the difference of  $\frac{1}{4}c$  in height.

The peaks on the Harker line, as caused by the n-glide perpendicular to c, at  $\frac{1}{2} \frac{1}{2} W$  where  $W (=2z) = 0$  and  $0.50$  clearly indicate that the niobium atoms are at the heights of 0 (or  $0.50$ ) and  $0.25$  (or  $0.75$ ) respectively.

The first two sets of niobium atoms were found from the Harker plane UVO.5. The peak at  $0.40, 0.02, 0.5$ , assumed to be corresponding to the vector  $(x+y, -\frac{1}{2}(x-y), \frac{1}{2})$ , yields the values of  $x = -0.06$  and  $y = 0.46$ . This information and that obtained from the Harker line led to the likely combination of certain co-ordinates of the "possible" set of atoms, which were then tested for their validity in generating the vectors which had to correspond to the actual Patterson peaks. An atom belonging to the first set of niobium atoms, Nb(1), was consequently assigned at  $-0.06, 0.46, 0.75$ . Another atom representing the second set of niobium atoms, Nb(2), was



similarly deduced from the peak at 0.00, 0.18, 0.50 as the vector  $\frac{-1+(x-y)}{2}, (x+y), \frac{1}{2}$  which in turn gave its co-ordinates in real space at 0.34, -0.16, 0.75. The third and last set of niobium atoms, Nb(3), were expected to lie in the planes  $z = 0.00$  and  $0.50$ , and an atom of this set was located at  $x = 0.10, y = 0.26, z = 0.50$  from the peak 0.34, 0.36, 0.50 found in the P(UVO.5) section corresponding to the vector  $\frac{1+(x-y)}{2}, (x+y), \frac{1}{2}$ . These co-ordinates of the three crystallographically non-equivalent types of niobium atoms are tabulated in Table 4.2.2.

Table 4.2.1 Interatomic vectors derived from a general set of atoms in the space group  $P_{4/n}$  with the origin at  $\bar{1}$ .

Atomic position	$x, y, z$	$\frac{1-x}{2}, \frac{1-y}{2}, z$	$\bar{y}, \frac{1+x}{2}, \frac{1+z}{2}$	$\frac{1+y}{2}, \bar{x}, \frac{1+z}{2}$	$\bar{x}, \bar{y}, \bar{z}$	$\frac{1+x}{2}, \frac{1+y}{2}, \bar{z}$	$y, \frac{1-x}{2}, \frac{1-z}{2}$	$\frac{1-y}{2}, x, \frac{1-z}{2}$
$x, y, z$	0, 0, 0	$\frac{1-2x}{2}, \frac{1-2y}{2}, 0$	$-(x+y), \frac{1+(x-y)}{2}, \frac{1}{2}$	$\frac{1-(x-y)}{2}, -(x+y), \frac{1}{2}$	$2\bar{x}, 2\bar{y}, 2\bar{z}$	$\frac{1}{2}, \frac{1}{2}, 2\bar{z}$	$-(x-y), \frac{1-(x+y)}{2}, \frac{1}{2}, -2z$	$\frac{1-(x+y)}{2}, x-y, \frac{1}{2}, -2z$
$\frac{1-x}{2}, \frac{1-y}{2}, z$	$-\frac{1+2x}{2}, -\frac{1+2y}{2}, 0$	0, 0, 0	$-\frac{1+(x-y)}{2}, x+y, \frac{1}{2}$	$x+y, -\frac{1-(x-y)}{2}, \frac{1}{2}$	$-\frac{1}{2}, -\frac{1}{2}, 2\bar{z}$	$2x, 2y, 2\bar{z}$	$-\frac{1+(x+y)}{2}, -(x-y), \frac{1}{2}, -2z$	$x-y, -\frac{1+(x+y)}{2}, \frac{1}{2}, -2z$
$\bar{y}, \frac{1+x}{2}, \frac{1+z}{2}$	$x+y, -\frac{1-(x-y)}{2}, -\frac{1}{2}$	$\frac{1-(x-y)}{2}, -(x+y), -\frac{1}{2}$	0, 0, 0	$\frac{1+2y}{2}, -\frac{1-2x}{2}, 0$	$-(x-y), -\frac{1-(x+y)}{2}, -\frac{1-2z}{2}$	$\frac{1+(x+y)}{2}, -(x-y), -\frac{1-2z}{2}$	$2y, 2\bar{x}, 2\bar{z}$	$\frac{1}{2}, -\frac{1}{2}, 2\bar{z}$
$\frac{1+y}{2}, \bar{x}, \frac{1+z}{2}$	$-\frac{1+(x-y)}{2}, x+y, -\frac{1}{2}$	$-(x+y), \frac{1+(x-y)}{2}, -\frac{1}{2}$	$-\frac{1-2y}{2}, \frac{1+2x}{2}, 0$	0, 0, 0	$-\frac{1-(x+y)}{2}, x-y, -\frac{1-2z}{2}$	$x-y, \frac{1+(x+y)}{2}, -\frac{1-2z}{2}$	$-\frac{1}{2}, \frac{1}{2}, 2\bar{z}$	$2\bar{y}, 2x, 2\bar{z}$
$\bar{x}, \bar{y}, \bar{z}$	$2x, 2y, 2z$	$\frac{1}{2}, \frac{1}{2}, 2z$	$x-y, \frac{1+(x+y)}{2}, \frac{1+2z}{2}$	$\frac{1+(x+y)}{2}, -(x-y), \frac{1+2z}{2}$	0, 0, 0	$\frac{1+2x}{2}, \frac{1+2y}{2}, 0$	$x+y, \frac{1-(x-y)}{2}, \frac{1}{2}$	$\frac{1+(x-y)}{2}, x+y, \frac{1}{2}$
$\frac{1+x}{2}, \frac{1+y}{2}, \bar{z}$	$-\frac{1}{2}, -\frac{1}{2}, 2z$	$2\bar{x}, 2\bar{y}, 2z$	$-\frac{1-(x+y)}{2}, x-y, \frac{1+2z}{2}$	$-(x-y), -\frac{1-(x+y)}{2}, \frac{1+2z}{2}$	$-\frac{1-2x}{2}, -\frac{1-2y}{2}, 0$	0, 0, 0	$-\frac{1-(x-y)}{2}, -(x+y), \frac{1}{2}$	$-(x+y), -\frac{1+(x-y)}{2}, \frac{1}{2}$
$y, \frac{1-x}{2}, \frac{1-z}{2}$	$x-y, -\frac{1+(x+y)}{2}, -\frac{1+2z}{2}$	$\frac{1-(x+y)}{2}, x-y, -\frac{1+2z}{2}$	$2\bar{y}, 2x, 2z$	$\frac{1}{2}, -\frac{1}{2}, 2z$	$-(x+y), -\frac{1-2y}{2}, -\frac{1}{2}$	$\frac{1+(x-y)}{2}, -(x+y), -\frac{1}{2}$	0, 0, 0	$\frac{1-2y}{2}, -\frac{1+2x}{2}, 0$
$\frac{1-y}{2}, x, \frac{1-z}{2}$	$-\frac{1+(x+y)}{2}, -(x-y), -\frac{1+2z}{2}$	$-(x-y), \frac{1-(x+y)}{2}, -\frac{1+2z}{2}$	$-\frac{1}{2}, \frac{1}{2}, 2z$	$2y, 2\bar{x}, 2z$	$-\frac{1-(x-y)}{2}, -(x+y), -\frac{1}{2}$	$x+y, \frac{1-(x-y)}{2}, -\frac{1}{2}$	$-\frac{1+2y}{2}, \frac{1-2x}{2}, 0$	0, 0, 0



Fig.4.2.1 Patterson sections for triniobium arsenide. (i) P(UVO),

(ii) P(UVO.25), (iii) P(UVO.5), (iv) P(UVO.75).

(i)

TYPE OF SYNTHESIS = PATT		.....																									
Rmax =	7.55	* Z = .000 *																									
		C = 132.3430542																									
		0	20	40	60	80	100	120	140	160	180	200	220	240	260	280	300	320	340	360	380	400	420	440	460	480	500
X= .0000	993	729	259	16	-26	-12	-5	-19	-18	7	31	39	29	22	48	71	50	11	-7	-3	11	9	-6	2	23	30	
X= .0201	729	519	160	-23	-57	-41	-25	-38	-43	-17	6	9	-3	4	49	80	51	3	-17	-6	28	36	2	-12	2	7	
X= .0401	259	160	15	-62	-84	-77	-52	-68	-64	-35	-19	-20	-27	-22	16	45	19	-13	-22	-10	29	35	-12	-31	-14	-12	
X= .0601	15	-3	-41	-67	-90	-92	-33	-86	-75	-44	-34	-40	-41	-34	2	31	-3	-37	-36	-28	-6	-3	-23	1	37	30	
X= .0801	-25	-32	-56	-74	-91	-89	-53	-89	-75	-41	-32	-40	-36	-13	50	92	40	-21	-24	-18	-18	-21	-2	2	24	90	
X= .1001	-12	-31	-67	-82	-85	-78	-70	-72	-57	-25	-18	-27	-25	9	91	140	78	1	-8	-0	-13	-28	0	90	145	114	
X= .1201	-5	-31	-68	-82	-83	-81	-74	-65	-44	-9	3	-13	-26	-6	67	94	45	-12	-17	-9	-15	-34	-25	37	87	70	
X= .1401	-19	-36	-64	-72	-76	-84	-53	-63	-18	44	65	24	-22	-31	-7	11	-7	-25	-19	-11	-5	-14	-24	3	32	22	
X= .1601	-18	-25	-47	-44	-39	-50	-12	-30	32	123	148	70	-7	-33	-29	-24	-31	-33	-22	-12	-4	-8	-20	-13	3	5	
X= .1801	7	-1	-30	-24	-4	-2	4	9	52	145	165	76	-2	-21	-14	-10	-23	-35	-34	-26	-19	-24	-36	-32	-6	28	
X= .2001	31	18	-23	-24	7	36	59	49	59	132	146	64	-1	-18	-10	6	-3	-24	-31	-28	-23	-26	-35	-38	4	56	
X= .2201	39	30	-14	-26	1	47	86	73	67	125	131	52	-9	-35	-33	-11	-14	-34	-39	-36	-32	-26	-27	-24	2	48	
X= .2401	28	27	-11	-20	-16	20	37	43	29	71	73	12	-34	-58	-62	-41	-37	-47	-35	-24	-29	-31	-33	-27	-0	15	
X= .2601	22	13	-16	-25	-13	17	37	45	2	2	3	-21	-34	-47	-54	-35	-26	-13	30	46	8	-28	-38	-22	19	26	
X= .2801	48	17	-14	-14	10	56	112	104	34	-7	-13	-15	-9	-12	-13	6	29	88	97	32	-22	-36	-10	45	53		
X= .3001	71	30	-6	-13	7	53	105	99	36	-4	-10	-8	-5	-2	20	50	5	39	71	68	12	-24	-26	13	85	94	
X= .3201	50	27	-5	-23	-21	2	35	29	-7	-22	-17	-14	-17	-4	51	105	89	37	17	3	-27	-36	-23	27	116	114	
X= .3401	11	13	-9	-31	-36	-29	-13	-15	-31	-31	-24	-25	-28	-6	67	107	82	14	-25	-34	-39	-37	-29	3	63	67	
X= .3601	-7	2	-22	-41	-35	-31	-28	-30	-35	-23	-21	-35	-49	-29	16	41	18	-18	-17	4	8	-1	-21	-27	0	13	
X= .3801	-3	-1	-29	-44	-29	-19	-22	-29	-31	-15	-3	-29	-57	-48	-7	-0	-15	-11	38	41	73	40	-6	-43	-30	-4	
X= .4001	11	-22	-33	-16	-7	-15	-28	-32	-13	-2	-32	-57	-33	-1	4	-11	-10	43	87	67	29	-12	-48	-36	-12		
X= .4201	9	-7	-24	-28	-13	-17	-35	-46	-50	-35	-25	-50	-64	-39	-10	1	-19	-48	-23	10	-10	-31	-38	-50	-35	-17	
X= .4401	-6	-19	-35	-33	-16	-25	-47	-51	-47	-27	-10	-33	-52	-39	-19	-1	-18	-66	-65	-45	-66	-76	-64	-58	-35	-10	
X= .4601	2	-11	-31	-20	7	4	-19	-27	-15	20	49	20	-23	-38	-17	7	3	-45	-54	-38	-59	-74	-64	-40	24	22	
X= .4801	23	2	-20	6	49	55	25	-2	2	43	82	53	2	-8	8	40	50	6	-18	-11	-27	-40	-30	31	179	273	
X= .5001	39	7	-12	30	58	114	70	22	5	23	56	48	15	26	53	94	114	67	13	-4	-12	-17	-10	73	273	296	

TYPE OF SYNTHESIS = PATT

R4R(= 7.55

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Z = .251  
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C = 132.3438542

	0	20	40	60	80	100	120	140	160	180	200	220	240	260	280	300	320	340	360	380	400	420	440	460	480	500
X= .0001	-13	-8	-9	-10	1	-12	-36	-38	-40	-40	-34	-35	-16	15	5	-16	-8	-11	-33	-39	-32	-26	-30	-49	-59	-57
X= .0211	-8	-7	-10	-1	19	5	-26	-32	-33	-32	-27	-26	-5	22	6	-9	5	-4	-27	-18	-10	-21	-28	-42	-57	-56
X= .0431	-9	-13	-15	-0	18	6	-19	-27	-28	-17	6	6	4	17	10	5	17	-8	-7	31	42	18	6	-13	-39	-44
X= .0651	-10	-14	-8	5	12	-1	-14	-12	-6	25	4	79	52	44	38	35	43	27	36	42	104	73	58	30	-11	-27
X= .0871	1	-3	18	16	12	-6	-11	-0	9	43	45	118	85	68	95	41	38	27	39	46	94	68	59	32	-7	-22
X= .1091	-12	-11	5	8	3	-6	-17	-11	-0	17	69	79	49	35	23	9	3	-5	7	41	42	25	21	3	-22	-24
X= .1231	-36	-38	-16	-11	-7	-7	-17	-19	-16	7	49	48	9	-11	-18	-24	-25	-29	-9	27	25	2	-6	-21	-36	-32
X= .1491	-35	-35	-27	-24	-11	-3	-13	-22	-16	13	48	41	-6	-34	-36	-39	-45	-50	-30	8	8	-17	-30	-44	-45	-25
X= .1601	-40	-41	-34	-26	1	24	7	-15	-14	0	13	8	-23	-39	-36	-40	-46	-52	-50	-30	-26	-39	-45	-50	-34	2
X= .1811	-40	-40	-26	-4	46	86	65	32	22	7	-13	-19	-38	-28	-24	-33	-35	-37	-43	-38	-33	-35	-29	-27	-22	4
X= .2001	-34	-36	-13	16	74	119	97	64	55	27	-5	-18	-25	-18	-17	-34	-37	-35	-38	-24	-11	-7	4	1	-12	-7
X= .2211	-35	-35	-21	5	54	87	63	37	32	11	-10	-16	-25	-22	-19	-31	-33	-34	-35	-8	21	26	32	38	15	16
X= .2411	-16	-27	-15	7	40	55	28	3	-4	-15	-18	-14	-24	-23	-14	-19	-20	-18	-18	8	48	40	35	30	18	26
X= .2611	15	6	20	32	44	46	14	-16	-21	-24	-20	-15	-24	-20	-9	15	-12	-1	-1	19	48	37	10	-6	-17	-9
X= .2801	5	7	22	28	26	22	-6	-33	-33	-33	-29	-17	-23	-23	-13	-18	10	33	34	55	79	50	2	-20	-31	-31
X= .3001	-16	-19	-9	-6	-10	-11	-23	-34	-33	-33	-34	-17	-22	-29	-19	-8	24	56	62	86	106	53	7	-16	-28	-32
X= .3211	-8	-31	-36	-38	-36	-38	-37	-42	-36	-39	-35	-24	-31	-34	-21	-14	-0	18	26	51	71	35	-10	-31	-42	-42
X= .3401	-11	-42	-55	-51	-48	-22	-29	-49	-50	-45	-40	-34	-39	-28	2	8	-10	-21	-18	1	15	-2	-28	-40	-45	-48
X= .3601	-33	-51	-49	-37	-19	15	16	-23	-41	-33	-33	-35	-35	-13	31	41	7	-18	-18	-11	-7	-14	-23	-29	-35	-32
X= .3811	-39	-46	-28	-11	2	35	41	-8	-22	-20	-24	-22	-15	1	37	42	5	-17	-18	-16	-13	-10	-6	-18	-27	-38
X= .4001	-32	-	-5	18	23	46	49	11	-2	3	10	17	28	48	64	55	10	-12	-14	-17	-11	0	13	12	-11	-11
X= .4211	-26	-17	22	52	61	88	88	43	30	44	44	48	55	72	101	88	35	8	0	-9	-4	9	17	16	0	4
X= .4411	-33	-20	17	48	66	98	92	42	31	45	39	34	35	43	70	65	22	8	-6	-12	0	13	9	1	-6	-7
X= .4611	-43	-45	-20	2	17	48	29	-4	5	23	13	17	19	5	8	5	-17	-22	-19	-16	11	27	8	-8	-10	-12
X= .4801	-59	-55	-40	-27	-18	-9	-19	-22	7	17	3	20	32	1	-21	-25	-35	-30	-23	-19	11	27	4	-12	-14	-18
X= .5001	-57	-56	-44	-27	-22	-24	-32	-25	2	4	-7	16	26	-9	-31	-32	-42	-40	-32	-30	-11	4	-7	-12	-18	-25

(ii)



Table 4.2.2 The determination of atomic positions from the Patterson map.

Atom	Patterson peak in vector space P(UVW)	Atomic co-ordinate in real space (xyz)
Nb(1)	0.40, 0.02, 0.5	-0.06, 0.46, 0.75
Nb(2)	0.00, 0.18, 0.5	0.34, -0.16, 0.75
Nb(3)	0.34, 0.36, 0.5	0.10, 0.26, 0.50

The positions of the remaining atoms can be obtained from the Fourier synthesis, that is the calculation of the electron density map. The phases assigned were obtained from the atomic co-ordinates shown in Table 4.2.2. New atomic parameters which were more accurate than before could then be assigned.

The electron density was calculated in two dimensions. The result was a map representing the projection of the electron density of the cell down the c axis onto the ab plane.

The Fourier map was calculated at four sections  $z = 0, 0.25, 0.50$  and  $0.75$  as shown in Fig.4.2.2.

Peaks from the  $F_0$  synthesis, by drawing closed curves through points of constant electron density. The atomic positions can be obtained from the points of the highest value of the electron density. We can get the position of As atom thus

As is at 0.04, 0.27, 0

The fractional and absolute co-ordinates of all atoms in the unit cell of  $Nb_3As$  are tabulated in Table 4.2.3.

Fig.4.2.2  $F_0$  synthesis, based on the three Nb atoms only,

(i)  $z=0$ , (ii)  $z=0.25$ , (iii)  $z=0.5$   
(i)

TYPE OF SYNTHESIS = F065

RMAX = 27.69

Z = .000

C = 36.8743589

	0	20	40	60	80	100	120	140	160	180	200	220	240	260	280	300	320	340	360	380	400	420	440	460	480	500
X= .0000	-80	-52	-32	-16	14	-16	-51	-38	-71	-94	-62	-64	-34	17	-19	-36	1	4	11	-1	-3	21	26	-4	-44	-68
X= .0201	-3	-73	-79	-40	38	32	-10	-17	-42	-62	-65	-47	121	351	322	105	-2	-15	-26	-18	3	39	66	48	2	-46
X= .0401	3	15	3	-8	59	91	32	21	12	25	-0	-27	208	604	607	228	-9	-13	3	18	19	-0	16	28	15	-10
X= .0601	10	36	13	-25	23	84	48	-11	-27	7	2	-65	9	329	421	105	-50	-47	-33	22	13	-28	-43	-15	12	5
X= .0801	-49	-61	-71	-54	-17	24	11	-49	-89	-45	16	-33	-12	140	103	-42	-38	-47	-62	26	44	-6	-4	-11	-26	-13
X= .1001	-66	-101	-109	-62	-0	31	18	-43	-74	-2	62	-4	-56	-8	-25	-44	38	28	-34	15	22	-12	15	-16	-82	-59
X= .1201	-43	-70	-116	-102	3	63	21	-55	-75	-21	-10	-73	-95	-78	-82	-52	15	1	-56	-17	13	-64	-43	-21	-58	-56
X= .1401	-57	-67	-81	-82	-10	40	-8	-28	-85	-69	-88	-110	-81	-61	-78	-74	-64	-88	-97	-26	10	-46	-103	-58	-24	6
X= .1601	-62	-65	-27	0	21	14	-50	-87	-39	-27	-81	-105	-65	-24	-36	-54	-51	-78	-99	-44	-1	-59	-125	-124	-78	-3
X= .1801	-67	-108	-66	-27	-5	-17	-89	-111	-57	-55	-105	-119	-104	-77	-70	-77	-50	-52	-82	-62	-64	-89	-73	-88	-114	-64
X= .2001	-91	-117	-67	-64	37	153	66	81	-120	-124	-108	-91	-105	-101	-94	-103	-77	-50	-53	-49	-76	-95	-20	20	-43	-66
X= .2201	-58	-39	15	64	302	698	632	71	-77	-73	-64	-23	-45	-50	-51	-50	-24	-25	-7	90	119	10	-30	41	19	-60
X= .2401	-16	1	27	147	613	999	711	177	-8	-2	1	3	-19	-23	-13	2	23	-2	89	451	631	954	62	9	11	-47
X= .2601	-43	7	6	64	350	629	453	108	-2	24	2	-12	-23	-20	2	-0	-4	-11	113	705	998	616	149	29	7	-13
X= .2801	-59	18	41	-29	117	50	-7	-26	-24	-50	-52	-50	-46	-23	-44	-81	-79	11	78	632	887	64	16	-34	-58	
X= .3001	-65	-44	20	-19	-94	-75	-46	-53	-50	-76	-103	-95	-102	-105	-90	-108	-125	-119	-80	69	150	38	-65	-67	-114	-90
X= .3201	-64	-114	-89	-73	-89	-63	-62	-83	-53	-49	-77	-70	-78	-105	-118	-165	-56	-56	-111	-90	-17	-3	-28	-66	-107	-67
X= .3401	-1	-78	-123	-125	-59	-1	-43	-100	-79	-51	-54	-36	-25	-65	-105	-82	-27	-39	-88	-53	18	22	0	-28	-65	-64
X= .3601	7	-24	-50	-101	-46	31	-24	-96	-87	-64	-74	-78	-61	-81	-109	-88	-69	-86	-79	-10	45	-10	-82	-80	-66	-58
X= .3801	-57	-60	-21	-42	-45	13	-17	-58	1	10	-54	-85	-78	-92	-69	-8	-20	-75	-94	21	64	3	-101	-117	-77	-46
X= .4001	-60	-85	-17	15	-13	21	13	-37	26	39	-47	-29	-8	-51	-0	63	-1	-73	-41	19	32	1	-60	-108	-100	-67
X= .4201	-42	-27	-11	-3	-7	44	26	-63	-47	-37	-44	100	141	-8	-31	-15	-45	-88	-48	11	24	-16	-54	-72	-61	-48
X= .4401	9	13	-14	-41	-27	33	22	-33	-47	-50	103	119	431	94	-65	1	7	-26	-12	40	83	23	-26	12	36	11
X= .4601	-7	15	28	17	1	19	17	1	-14	-11	225	605	601	210	-26	0	25	11	19	52	92	60	-9	3	16	1
X= .4801	-45	0	47	66	40	4	-19	-28	-16	-2	103	320	352	123	-46	-64	-62	-42	-18	-10	12	38	-40	-79	-73	-55
X= .5001	-68	-45	-9	25	21	-2	0	0	4	3	-35	-20	17	-34	-65	-62	-93	-71	-37	-50	-16	13	-17	-31	-51	-80
X= .5201	-45	-44	3	17	1	17	45	29	7	11	-18	-53	-43	-26	1	-5	-64	-55	-19		1	38	44	71	33	-82
X= .5401	-9	-20	30	84	27	13	2	-34	-28	-6	-1	-1	-12	0	32	17	-32	-36	0	33	71	85	72	-81	54	4
X= .5601	5	-34	-9	47	22	-5	-12	-40	-27	-10	-24	-23	-31	-22	5	-9	-27	-14	35	99	117	98	67	23	-1	18
X= .5801	-14	-21	-29	-26	-40	-21	16	4	-1	11	-27	53	-21	-0	-32	-79	-82	-36	35	92	96	55	34	-15	-49	-46
X= .6001	-57	-24	-36	-32	-52	-25	-7	-27	-36	-17	-31	-43	8	24	-42	-97	-104	-63	-1	20	22	36	39	18	-25	-67
X= .6201	-54	-38	-44	-9	29	39	24	-30	-70	-59	-29	15	57	45	4	-31	-52	-24	4	-15	23	81	49	7	-12	-48
X= .6401	7	13	-18	29	81	89	100	88	-1	-23	-3	58	85	37	15	-2	-35	1	29	11	53	70	-23	72	-52	-58
X= .6601	-5	17	-31	-10	39	44	96	121	49	8	3	18	18	-29	-45	-43	-65	-24	22	16	16	15	-63	-65	-36	-59
X= .6801	-65	-28	-61	-67	-32	-4	39	57	-7	-58	-49	-38	-24	-44	-91	-91	-69	-54	-36	-12	32	15	-54	-29	5	-66
X= .7001	-63	-25	-24	-48	-45	-18	1	-1	-47	-101	-81	-23	11	-13	-80	-80	-47	-78	-02	-16	32	-7	-86	-58	-14	-95









Table 4.2.3 The atomic positions in the unit cell of Nb<sub>3</sub>As.

Atom	Fractional co-ordinate from Patt.map and F <sub>o</sub> synthesis.	The equivalent positions	
		Fractional co-ordinate	Absolute co-ordinate
Nb(1)	-0.06, 0.46, 0.75	-0.06, 0.46, 0.75	9.676, 4.735, 3.897
		0.56, 0.04, 0.75	5.765, 0.412, 3.897
		0.54, 0.44, 0.25	5.559, 4.529, 1.299
		-0.04, 0.06, 0.25	9.882, 0.618, 1.299
		0.06, 0.54, 0.25	0.618, 5.559, 1.299
		0.44, -0.04, 0.25	4.529, 9.882, 1.299
		0.46, 0.56, 0.75	4.735, 5.765, 3.897
		0.04, -0.06, 0.75	0.412, 9.676, 3.897
Nb(2)	0.34, -0.16, 0.75	0.34, -0.16, 0.75	3.500, 8.647, 3.897
		0.16, 0.66, 0.75	1.647, 6.794, 3.897
		0.16, -0.16, 0.25	1.649, 8.647, 1.299
		0.34, 0.66, 0.25	3.500, 6.794, 1.299
		0.66, 0.16, 0.25	6.794, 1.647, 1.299
		-0.16, 0.34, 0.25	8.647, 3.500, 1.299
		-0.16, 0.16, 0.75	8.647, 1.647, 3.897
		0.66, 0.34, 0.75	6.794, 3.500, 3.897
Nb(3)	0.10, 0.26, 0.50	0.10, 0.26, 0.50	1.029, 2.676, 2.598
		0.40, 0.24, 0.50	4.118, 2.471, 2.598
		-0.26, 0.60, 0	7.618, 6.176, 0
		0.76, -0.10, 0	7.823, 9.265, 0
		-0.10, -0.26, 0.50	9.265, 7.618, 2.598
		0.60, 0.76, 0.50	6.176, 7.823, 2.598
		0.26, 0.40, 0	2.676, 4.118, 0
		0.24, 0.10, 0	2.471, 1.029, 0

Table 4.2.3

(continued)

Atom	Fractional co-ordinate from Patt.map and $F_o$ synthesis.	The equivalent positions	
		Fractional co-ordinate	Absolute co-ordinate
As	0.04 , 0.27 , 0	0.04, 0.27, 0	0.412, 2.779, 0
		0.46, 0.23, 0	4.735, 2.368, 0
		0.73, 0.54, 0.5	7.515, 5.559, 2.598
		0.77, -0.04, 0.5	7.926, 9.882, 2.598
		-0.04, 0.73, 0	9.882, 7.515, 0
		0.54, 0.77, 0	5.559, 7.926, 0
		0.27, 0.46, 0.5	2.779, 4.735, 2.598
		0.23, 0.04, 0.5	2.368, 0.412, 2.598

Fig.4.2.3 shows the positions of Nb atoms and As atoms in four neighbouring unit cells of  $Nb_3As$  in the three sections of  $z$ .

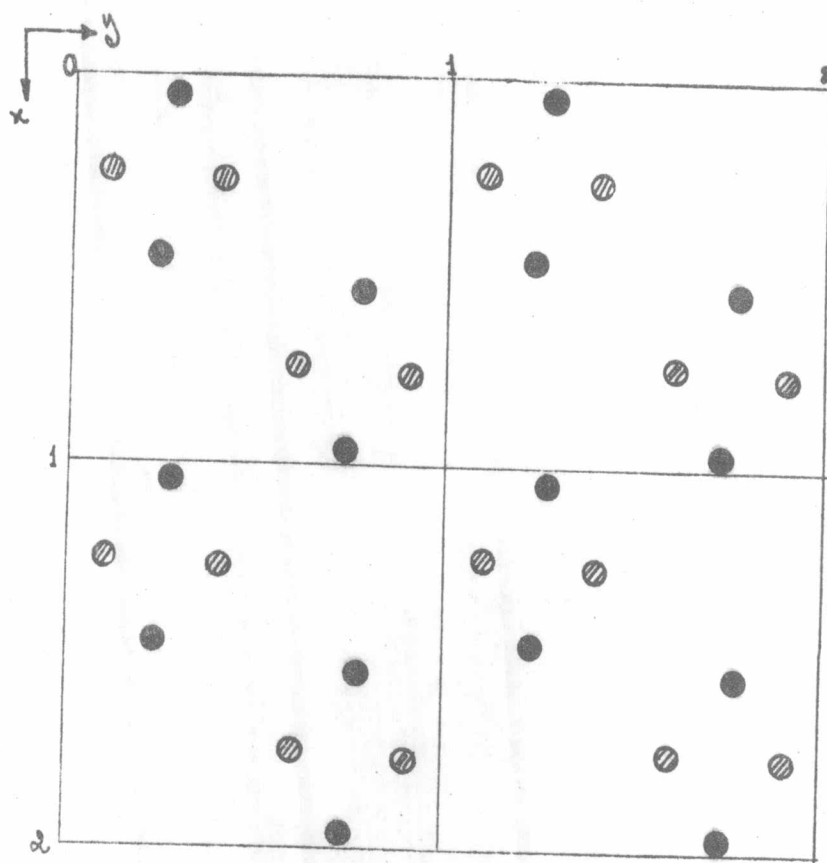


Fig.4.2.3.(i)

- Atom
- Nb(1)
  - ⊘ Nb(2)
  - ⊙ Nb(3)
  - As

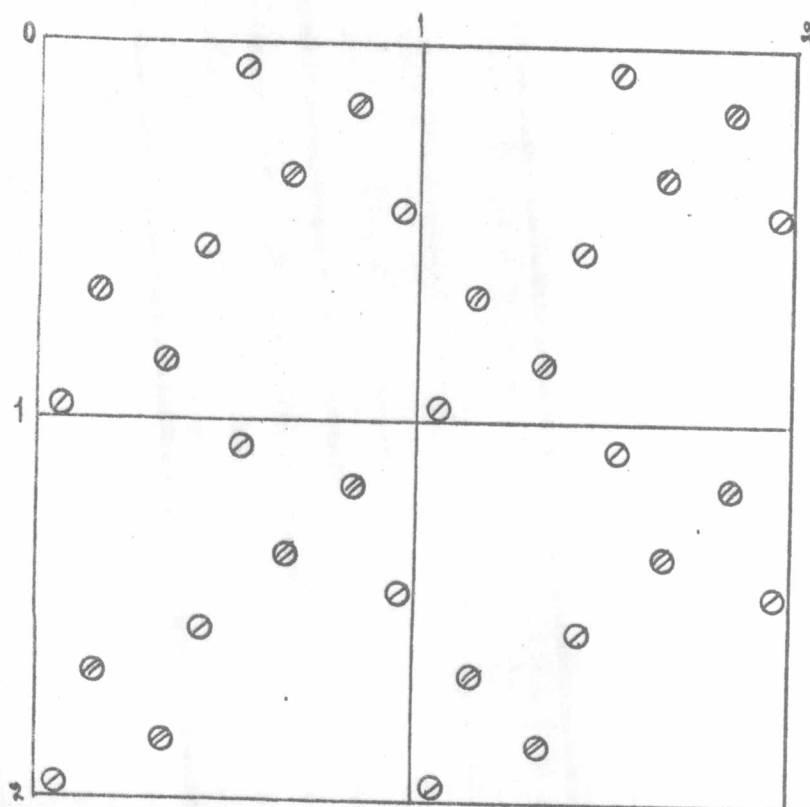


Fig.4.2.3.(ii)

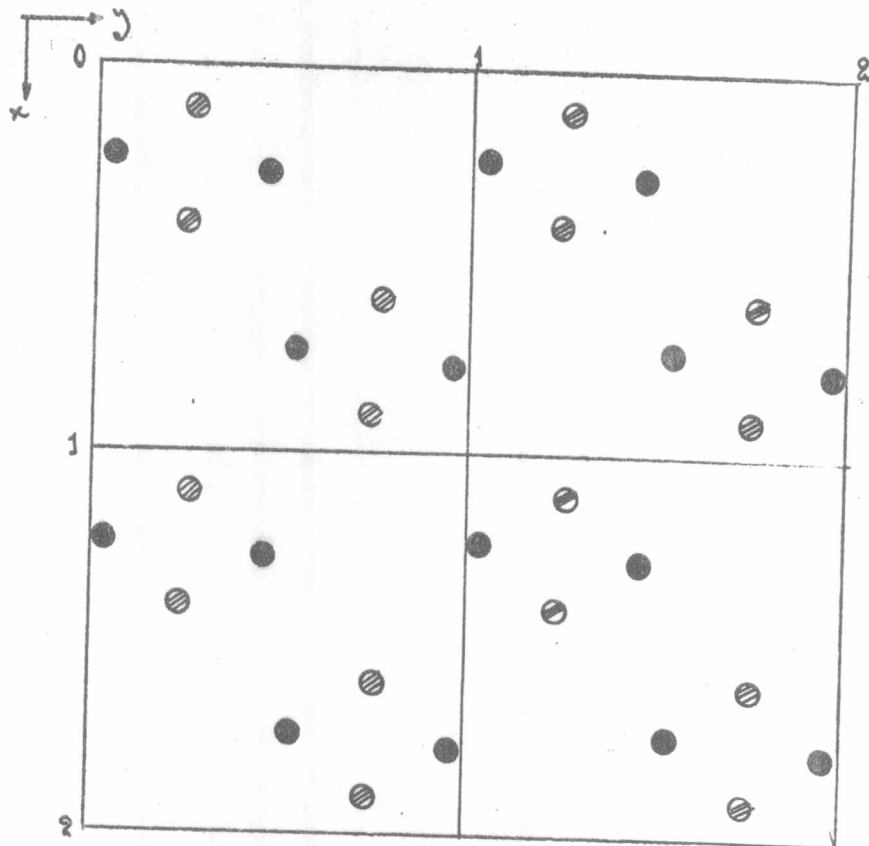
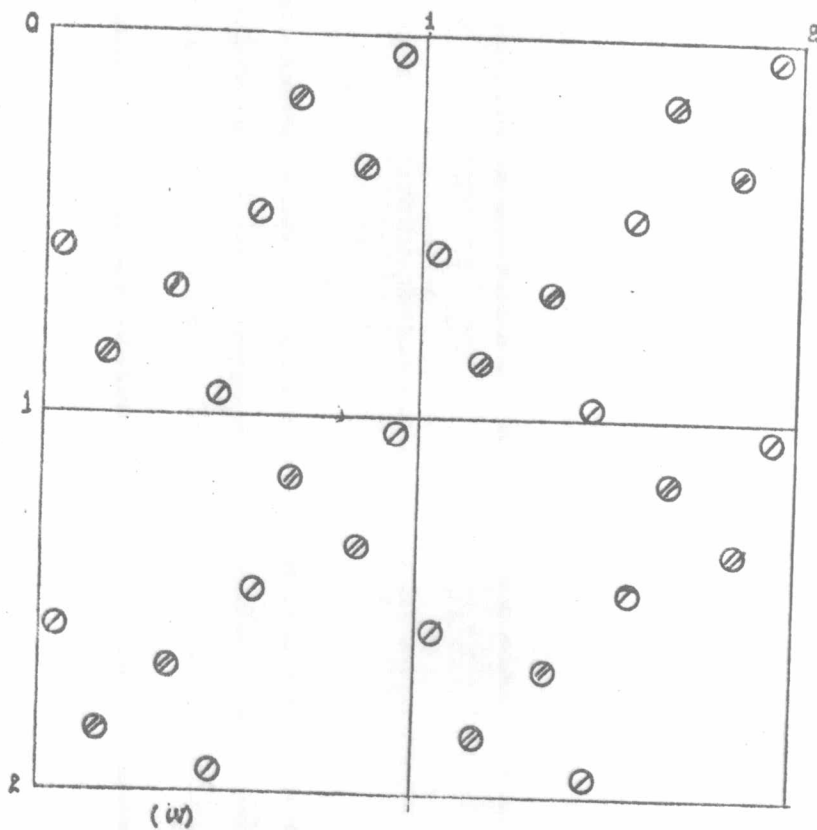


Fig.4.2.3. (iii)

- Atom
- Nb(1)
  - ◐ Nb(2)
  - ◑ Nb(3)
  - As



(iv)

Fig.4.2.3. The atomic positions in four unit cells. (i)

$z=0$ , (ii)  $z=0.25$ , (iii)  $z=0.5$ , (iv)  $z=0.75$ .

### 4.3 Refinement

#### 4.3.1 Refinement by $F_o$ synthesis

The accurate positional parameters of each atom can be determined from this  $F_o$  synthesis as given in section 4.2.

The Booth's method gives the accurate determination, provided the electron density is computed at equidistant points. The maximum in the xy section of an  $F_o$  synthesis for Nb(1) atom was determined as shown in Fig. 4.3.1 (i). As the peak near the line  $x = 0.56$ ,  $y = 0.04$  corresponded to Nb(1) was more conveniently observed than that at  $x = -0.06$ ,  $y = 0.46$ . The peak was therefore chosen and subjected to the Booth method, and these are tabulated in Table 4.3.1.

Using the equation

$$X_m = \frac{(P_2/P_1) - 4}{(2 P_2/P_1) - 4}$$

and

$$Y_m = \frac{(P_2/P_1) - 4}{(2 P_2/P_1) - 4}$$

the data in Table 4.3.1 gives  $X_m = 0.0255$  and  $Y_m = 0.0262$  and the actual x and y co-ordinates of the Nb(1) atom are 0.5655 and 0.0462.

Following the same criterion, the Nb(2) atom appearing at  $x = 0.66$ ,  $y = 0.16$ ,  $z = 0.25$  rather than that at  $x = 0.34$ ,  $y = -0.16$ ,  $z = 0.75$  was chosen to undergo the Booth treatment. The chosen peaks of Nb(2), Nb(3) and As atomic co-ordinates were similarly determined from the xy section shown in Fig. 4.3.1 (ii), (iii) and (iv) respectively. The tables of function  $\Delta\rho$  of the x and y co-ordinates were shown in Table 4.3.2, 4.3.3 and 4.3.4.

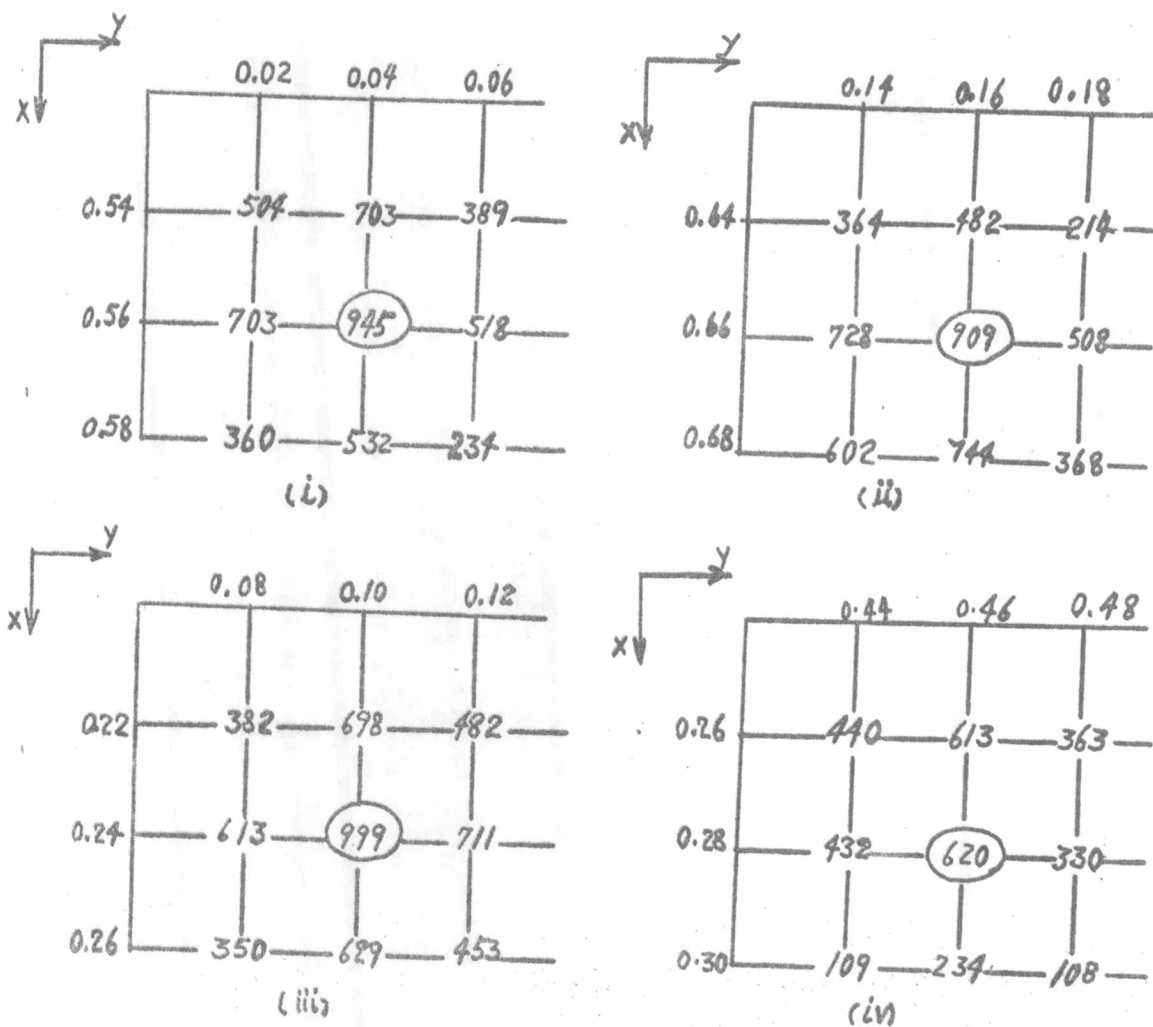


Fig.4.3.1 (i) Section through a Nb(1) atom in  $F_0$  synthesis at  $z = 0.75$ .  
(ii) Section through a Nb(2) atom in  $F_0$  synthesis at  $z = 0.25$ .  
(iii) Section through a Nb(3) atom in  $F_0$  synthesis at  $z = 0$ .  
(iv) Section through a As atom in  $F_0$  synthesis at  $z = 0.5$ .



Table 4.3.1 The function  $\Delta p$  of the x and y co-ordinates of Nb(1) atom.

x	0.54	0.56	0.58
$\Delta x$	0	0.02	0.04
$\Delta p$	0	242	171
y	0.02	0.04	0.06
$\Delta y$	0	0.02	0.04
$\Delta p$	0	242	185

Table 4.3.2 The function  $\Delta p$  of the x and y co-ordinates of Nb(2) atom.

x	0.64	0.66	0.68
$\Delta x$	0	0.02	0.04
$\Delta p$	0	427	262
y	0.14	0.16	0.18
$\Delta y$	0	0.02	0.04
$\Delta p$	0	181	220

Table 4.3.3 The function  $\Delta p$  of the x and y co-ordinates of Nb(3) atom.

x	0.22	0.24	0.26
$\Delta x$	0	0.02	0.04
$\Delta p$	0	301	69
y	0.08	0.10	0.12
$\Delta y$	0	0.02	0.04
$\Delta p$	0	386	98

Table 4.3.4 The function  $\Delta p$  of the x and y co-ordinates of As atom.

x	0.26	0.28	0.30
$\Delta x$	0	0.02	0.04
$\Delta p$	0	7	379
y	0.44	0.46	0.48
$\Delta y$	0	0.02	0.04
$\Delta p$	0	188	102

And from the other peaks, the co-ordinate of these atoms were given. The average value of the positional co-ordinates were given in the Table 4.3.5.

Table 4.3.5 The co-ordinates of Nb atoms and As atom were obtained from  $F_o$  synthesis.

Atom	x	y	z
Nb(1)	0.5465	0.4386	0.2500
Nb(2)	0.6644	0.1455	0.2500
Nb(3)	0.2401	0.0998	0
As	0.0402	0.2699	0

By assuming the atoms were at exactly  $z = 0$ , and  $0.25$ , it was inspected that the calculated values of the structure factors of a number of reflections were zero whereas the corresponding observed values were small. It was therefore concluded that the co-ordinates in  $z$  were approximately correct, however they needed be slightly shifted. The following set of atomic parameters with the newly adjusted values of  $z$  and the smaller values in the  $x$ -direction, were thus employed later on in the subsequent calculation which yielded a better agreement between the observed and calculated structure factors.

Table 4.3.6 The positional parameters of atoms in Nb<sub>3</sub>As.

Atom	x	y	z	Biso
Nb(1)	0.0614	0.5465	0.2400	0.2500
Nb(2)	0.1644	0.6755	0.7400	0.2500
Nb(3)	0.0988	0.2599	0.5100	0.2500
As	0.0402	0.2699	0.0100	0.2500

## 4.3.2 Least-squares refinement

After the structure factor calculation, the R value that is  $\frac{\sum_{hkl} |F_o| - |F_c|}{\sum_{hkl} |F_o|}$  was given. The output included the scale

factors for each layers.

In the initial stage, refinement was restricted to the scale factors, isotropic temperature factors, and the positional parameters of the four atoms.

The weighting function was a measure of the relative of the observation. According to Cruickshank this was calculated, thus,

$$W = \frac{1}{(a + |F_o| + c|F_o|^2)}$$

where  $a = 2 F_{o \text{ min}}$

and  $c = 2/F_{o \text{ max}}$

$F_{\text{min}}$  and  $F_{\text{max}}$  are the minimum and maximum values of the observed structure factors.

The atomic positions and thermal parameters were calculated for each atom to decrease R. Then the scale factor for each layer were introduced in the refinement. So the least shift of the parameters was given. The final R value being in the range of 0.1244, the parameters are shown in Table 4.3.7.

Table 4.3.7 The results from the last cycle of the least-square method for the isotropic temperature factor.

Atom	x	$\delta(x)$	y	$\delta(y)$	z	$\delta(z)$	Bisc
Nb(1)	0.0568	0.0005	0.5361	0.0005	0.239	0.001	0.5730
Nb(2)	0.1651	0.0005	0.4544	0.0005	0.730	0.001	0.8104
Nb(3)	0.1024	0.0005	0.2621	0.0005	0.516	0.001	0.6356
As	0.0419	0.0006	0.2714	0.0005	0.017	0.002	0.5166

From the last LSQR program, the observed structure factors were compared with the calculated structure factors for observed reflections of  $\text{Nb}_3\text{As}$  crystal in 5 layers as shown in Table 4.3.8.

Table 4.3.8 The observed and calculated structure factors for the observed reflections of Nb<sub>3</sub>As.

H	K	L	F <sub>o</sub>	F <sub>c</sub>	H	K	L	F <sub>o</sub>	F <sub>c</sub>
0	4	0	98.01	95.43	12	10	0	99.52	117.25
0	6	0	80.39	73.86	12	12	0	65.22	71.00
0	10	0	261.73	256.28	13	11	0	75.48	49.53
0	14	0	139.03	114.49	14	0	0	128.48	114.49
0	16	0	117.96	100.60	14	6	0	140.90	138.36
1	5	0	459.21	415.60	15	3	0	153.86	141.63
1	7	0	214.60	223.29	16	0	0	117.96	100.60
1	9	0	121.48	130.90	16	2	0	119.69	115.68
1	13	0	87.47	69.60	16	8	0	77.65	83.53
1	15	0	174.79	169.82	0	3	1	62.94	53.49
2	4	0	143.75	160.71	0	4	1	190.57	162.13
2	6	0	138.26	149.41	0	10	1	171.31	156.40
2	10	0	111.82	98.91	0	11	1	64.68	81.81
2	12	0	161.74	128.16	0	13	1	90.49	105.51
2	14	0	92.71	95.13	1	3	1	126.63	109.01
3	1	0	41.02	45.57	1	4	1	354.90	341.85
3	3	0	528.12	480.08	1	5	1	112.59	104.61
3	7	0	202.76	228.06	1	6	1	216.54	223.44
3	9	0	60.96	52.17	1	7	1	175.87	171.76
3	11	0	72.18	68.42	1	9	1	84.42	91.43
3	15	0	136.90	32.05	1	10	1	62.09	67.88
3	17	0	97.76	75.53	1	14	1	122.04	134.31
4	0	0	98.01	95.43	2	3	1	257.41	268.33
4	2	0	435.74	408.16	2	4	1	185.46	174.80
4	6	0	111.66	117.48	2	5	1	69.54	67.73
4	8	0	165.90	167.63	2	6	1	37.84	57.57
4	10	0	99.03	101.74	2	7	1	105.33	104.94
4	12	0	280.27	272.15	2	9	1	132.66	136.79
4	16	0	79.30	46.00	2	10	1	134.53	128.55
5	1	0	139.75	135.22	2	12	1	52.47	62.36
5	3	0	129.99	134.83	2	13	1	48.77	64.45
5	9	0	210.35	203.22	2	15	1	52.76	66.31
5	13	0	135.37	121.02	3	0	1	62.94	53.49
5	15	0	136.37	112.66	3	1	1	70.18	66.03
6	0	0	80.39	73.86	3	2	1	336.55	316.10
6	2	0	166.73	173.75	3	3	1	94.64	93.34
6	4	0	83.78	195.13	3	5	1	79.55	77.75
6	6	0	122.38	223.48	3	6	1	115.67	108.73
6	8	0	174.40	168.88	3	7	1	129.79	132.20
6	10	0	110.71	101.78	3	8	1	201.84	216.98
7	1	0	194.05	193.11	3	9	1	94.31	103.20
7	3	0	193.67	196.60	3	10	1	53.78	58.78
7	5	0	184.01	169.95	3	12	1	69.41	83.90
7	7	0	109.71	91.48	3	16	1	51.72	94.89
7	9	0	96.34	65.20	4	0	1	190.57	162.13
7	11	0	107.58	78.67	4	1	1	300.12	235.38
8	2	0	117.58	100.89	4	2	1	192.77	186.92
9	1	0	173.90	167.60	4	3	1	57.72	51.23
9	7	0	114.84	105.32	4	4	1	27.86	44.45
9	11	0	171.82	161.46	4	6	1	171.52	185.96
10	0	0	261.73	256.28	4	9	1	155.85	152.95
10	2	0	63.83	69.25	4	10	1	59.50	74.47
10	8	0	167.66	105.13	4	12	1	58.48	73.40
11	1	0	236.34	230.87	5	1	1	308.81	170.78
11	3	0	111.24	103.14	5	4	1	137.37	119.46
11	9	0	153.70	104.91	5	5	1	87.03	86.12
11	7	0	71.42	57.51	5	6	1	56.89	72.30
11	11	0	101.75	66.47	5	7	1	232.40	226.34
12	6	0	181.93	188.98	5	11	1	66.99	81.19
					5	13	1	76.20	71.36
					6	1	1	173.14	154.99
					6	4	1	123.54	118.60
					6	6	1	146.77	135.79
					6	7	1	72.45	58.69

Table 4.3.8

(continued)

H	K	L	F <sub>0</sub>	F <sub>a</sub>	H	K	L	F <sub>0</sub>	F <sub>a</sub>
5	9	1	176.38	159.67	2	6	2	131.77	136.26
7	1	1	93.28	90.11	2	8	2	78.35	80.80
7	2	1	205.96	167.00	2	10	2	124.03	137.26
7	3	1	233.47	217.82	2	12	2	59.05	69.25
7	4	1	77.02	80.97	2	14	2	144.34	167.20
7	5	1	98.21	98.89	3	8	2	119.68	105.15
7	8	1	79.33	90.32	3	1	2	257.18	245.14
7	9	1	190.61	168.57	3	3	2	204.16	183.86
7	10	1	66.80	66.18	3	5	2	116.05	114.47
7	14	1	108.03	99.23	3	9	2	97.30	100.89
8	1	1	42.25	54.05	4	0	2	191.05	180.13
8	2	1	100.69	100.06	4	2	2	63.83	70.84
8	3	1	184.05	165.43	4	4	2	163.77	158.80
8	4	1	249.57	226.83	4	5	2	82.20	78.84
8	5	1	92.96	109.78	4	6	2	243.71	227.73
8	7	1	226.85	183.87	4	10	2	151.60	161.03
8	8	1	116.94	130.10	4	15	2	166.90	169.78
8	10	1	129.23	128.20	5	0	2	44.80	33.87
8	11	1	144.18	119.06	5	3	2	218.40	209.04
8	13	1	63.58	75.40	5	5	2	411.74	375.53
8	14	1	62.95	77.43	5	6	2	39.62	43.98
9	1	1	119.14	117.53	5	13	2	66.58	67.93
9	4	1	264.78	186.41	6	0	2	353.31	327.05
9	8	1	55.36	61.88	6	1	2	41.64	55.88
9	9	1	143.61	123.47	6	2	2	47.57	41.54
10	0	1	171.31	156.40	6	14	2	84.58	96.32
10	1	1	51.15	45.93	7	0	2	65.49	66.09
10	2	1	87.44	94.64	7	1	2	225.29	218.99
10	3	1	74.14	76.51	7	3	2	85.24	93.14
10	6	1	206.14	197.25	7	7	2	166.33	165.61
10	9	1	82.18	83.03	7	11	2	56.86	81.96
10	12	1	59.51	66.88	7	13	2	136.52	158.62
11	0	1	64.68	81.81	8	0	2	226.61	201.51
11	1	1	79.43	83.78	8	2	2	150.70	148.72
11	2	1	97.06	107.50	8	18	2	71.94	67.00
11	4	1	130.44	129.49	9	1	2	59.68	65.62
11	5	1	108.18	110.54	9	3	2	118.47	109.35
11	6	1	48.10	64.61	9	5	2	70.68	71.52
11	7	1	127.85	124.08	9	7	2	92.13	98.03
11	10	1	76.04	25.18	10	0	2	175.65	143.75
11	11	1	105.80	108.24	10	2	2	139.72	124.04
12	6	1	56.21	77.00	10	4	2	94.24	95.40
13	0	1	90.49	105.51	10	8	2	127.50	128.22
13	2	1	135.04	138.16	10	10	2	164.85	141.31
13	9	1	74.37	88.37	10	15	2	67.86	28.71
13	10	1	76.22	100.89	11	5	2	101.81	96.79
14	4	1	83.60	41.82	11	7	2	67.43	72.32
14	8	1	75.36	94.85	11	13	2	130.11	133.13
15	4	1	91.67	89.77	12	2	2	42	191.44
15	8	1	51.83	30.96	12	5	2	114.19	103.78
0	3	2	130.88	105.15	14	4	2	95.91	110.69
0	4	2	178.65	180.13	15	6	2	104.84	28.06
0	5	2	34.79	33.87	15	8	2	74.07	9.54
0	6	2	353.31	327.05	16	12	2	59.71	83.96
0	7	2	60.36	66.09	17	1	2	248.26	216.82
0	8	2	194.25	201.51	19	13	2	58.25	22.28
0	10	2	150.28	143.75	0	3	3	36.51	43.99
1	3	2	45.52	52.51	0	4	3	152.82	139.11
1	7	2	74.82	93.24	0	6	3	32.55	50.95
1	9	2	148.32	148.97	0	7	3	93.45	99.05
1	11	2	107.32	106.07	0	10	3	124.78	116.17
1	13	2	166.68	177.04	1	4	3	249.97	212.05
2	2	2	317.98	405.83	1	5	3	55.36	64.67
2	5	2	78.14	15.50	1	6	3	157.88	147.61

Table 4.3.8

(continued)

H	K	L	F <sub>o</sub>	F <sub>c</sub>	H	K	L	F <sub>o</sub>	F <sub>c</sub>
1	7	3	72.52	82.71	8	14	3	74.16	21.63
1	9	3	83.60	86.86	9	1	3	97.00	101.41
1	10	3	101.87	94.25	9	3	3	71.62	80.59
1	13	3	57.17	69.18	9	4	3	237.07	193.47
2	1	3	31.63	36.48	9	5	3	63.76	75.17
2	2	3	43.30	47.79	9	7	3	73.92	72.53
2	3	3	154.35	191.64	9	9	3	89.71	85.07
2	5	3	225.16	210.71	9	13	3	70.18	39.87
2	5	3	55.12	59.90	10	0	3	114.63	116.17
2	6	3	60.50	66.67	10	2	3	92.15	107.95
2	7	3	150.58	144.41	10	6	3	174.24	178.71
2	8	3	59.07	57.85	10	9	3	124.55	124.29
2	9	3	76.86	81.47	10	11	3	76.91	38.68
2	10	3	57.80	75.31	11	2	3	99.64	119.06
2	12	3	61.52	30.55	11	3	3	53.66	20.99
2	19	3	94.33	56.73	11	5	3	110.86	112.47
3	0	3	36.51	43.99	11	7	3	75.84	80.33
3	1	3	131.87	116.19	12	2	3	51.55	69.55
3	2	3	262.11	240.98	12	3	3	70.88	56.03
3	3	3	21.71	8.74	12	4	3	70.78	81.34
3	4	3	32.65	43.59	13	2	3	111.56	129.51
3	6	3	146.83	134.46	13	3	3	74.63	87.77
3	7	3	81.11	88.78	13	6	3	89.66	89.30
3	8	3	192.27	159.78	13	9	3	108.49	123.55
3	9	3	191.93	163.86	14	2	3	59.65	74.76
3	10	3	78.23	74.55	14	3	3	78.29	86.47
4	0	3	151.44	139.11	18	5	3	79.51	38.24
4	1	3	211.03	198.34	0	3	4	53.54	63.56
4	2	3	38.01	62.00	0	4	4	61.55	66.90
4	3	3	95.77	93.67	0	6	4	53.56	55.97
4	4	3	42.74	43.35	0	10	4	182.33	179.46
4	5	3	45.36	50.37	0	11	4	55.92	73.25
4	6	3	137.81	129.00	7	14	4	77.51	86.12
4	8	3	63.33	71.35	1	4	4	81.58	82.50
4	9	3	91.82	81.07	1	5	4	321.81	248.29
5	1	3	157.68	133.60	1	6	4	71.58	71.74
5	5	3	37.80	54.33	1	7	4	149.46	142.02
5	5	3	63.13	76.90	1	8	4	51.88	45.84
5	7	3	175.97	168.70	1	9	4	97.23	96.20
5	12	3	72.11	44.77	1	15	4	109.07	126.70
6	0	3	50.29	50.95	2	3	4	41.36	57.04
6	1	3	191.34	161.97	2	4	4	92.87	94.37
6	4	3	96.37	112.82	2	5	4	82.90	85.90
6	5	3	107.91	139.52	2	8	4	101.88	107.22
6	7	3	86.17	80.22	2	10	4	74.91	78.10
6	9	3	214.69	188.06	2	12	4	93.48	98.12
6	11	3	57.17	55.25	2	14	4	48.70	67.50
6	14	3	71.58	33.18	3	0	4	5	63.56
7	0	3	95.17	99.05	3	3	4	319.29	286.02
7	1	3	106.10	100.81	3	4	4	60.53	54.46
7	2	3	69.28	77.72	3	5	4	66.63	67.91
7	3	3	237.28	225.39	3	7	4	161.21	150.96
7	4	3	54.46	55.47	3	8	4	52.67	49.88
7	5	3	102.69	104.93	4	0	4	52.70	66.90
7	6	3	56.63	52.63	4	1	4	49.78	49.58
7	8	3	61.96	87.97	4	2	4	279.43	240.63
7	9	3	113.02	103.79	4	6	4	68.12	74.03
8	2	3	47.55	62.27	4	7	4	39.03	17.91
8	3	3	137.55	136.17	4	8	4	139.60	121.25
8	4	3	97.67	102.55	4	10	4	58.36	64.74
8	5	3	138.20	135.97	4	12	4	223.80	196.85
8	7	3	99.50	88.94	5	1	4	64.09	70.86
8	8	3	141.51	143.24	5	2	4	141.90	121.99
8	18	3	148.67	132.57	5	3	4	74.56	98.84



Table 4.3.8

(continued)

H	K	L	F <sub>0</sub>	F <sub>0</sub>
5	5	4	129.77	133.13
5	5	4	69.25	71.43
5	13	4	71.83	88.51
5	15	4	83.36	83.86
6	0	4	47.83	55.97
6	2	4	86.63	106.69
6	3	4	36.27	46.52
6	4	4	132.92	124.60
6	5	4	44.24	44.16
6	6	4	173.52	146.51
6	7	4	70.79	64.06
6	8	4	119.09	117.62
7	1	4	126.37	126.91
7	2	4	48.43	49.33
7	3	4	137.86	137.33
7	5	4	147.38	126.13
7	7	4	55.33	63.60
8	2	4	67.75	64.82
8	6	4	53.26	20.19
9	1	4	134.93	116.15
9	7	4	66.67	76.82
10	0	4	202.07	179.46
10	1	4	88.13	76.96
10	2	4	51.35	55.21
10	8	4	86.26	80.22
11	0	4	53.88	73.25
11	1	4	161.39	159.37
11	3	4	63.68	75.46
11	5	4	62.87	73.23
12	1	4	52.56	16.53
12	6	4	167.78	137.77
13	1	4	55.37	14.03
13	5	4	65.95	44.90
14	0	4	57.97	86.12
15	4	4	61.71	13.76
15	2	4	82.06	83.64

#### 4.4 Determination of accurate lattice constants

The precise cell constants of  $Nb_3As$  can be obtained from the powder method by using the XDC 700 focusing camera.

The procedure for identifying a powder diffraction pattern involved measuring the positions of the diffraction lines in the photograph shown in Fig.3.6.3. In turn, the diffraction angles, the spacing  $d$  of these reflecting planes were calculated.

From the diffraction patterns, the six strongest reflections probably those of silicon were used to plot the graph of  $\theta/(S-S_0)$  against  $(S-S_0)$  shown in Fig.4.4.1, the values of  $\theta/(S-S_0)$  and  $\theta$  for each reflection were obtained. These lines were then confirmed as silicon.  $S_0$  is the reference primary line read in mm. and  $S$  is the reading in mm. of each diffraction lines. The values of  $\theta/(S-S_0)$  of silicon corresponding to  $(S-S_0)$  and the values of reflecting angles  $\theta$  were obtained in Table 4.4.1.

The Miller indices of the diffraction lines were obtained by the trial-and-error method. These angles  $\theta$  and the d-spacings have to be related to the unit cell parameters and Miller indices assigned to the individual reflections. For the tetragonal system, the value of  $\sin^2 \theta_{hkl}$  can be expressed, thus:

$$\sin^2 \theta_{hkl} = A(h^2 + k^2) + Cl^2 \quad \dots (4.4.1)$$

where  $a$  and  $c$  are values of rough cell constants,

$$A = \frac{\lambda^2}{4a^2}$$

$$C = \frac{\lambda^2}{4c^2}$$

All possible values of  $\sin^2 \theta_{hkl}$  based on the values of A, C and indices hkl are calculated and compared with the observed values of  $\sin^2 \theta_{hkl}$ . If the agreement is better than 0.0005, the indexing is probably correct.

Table 4.4.1 The standard silicon lines for  $\text{CuK}_{\alpha_1}$  radiation.

hkl	$\theta^\circ$	$(S-S_o)_{cal}$ (mm)	$(S-S_o)_o$ (mm)	$\theta / (S-S_o)_{cal}$	$\theta / (S-S_o)_o$
111	14.2214	49.725	49.68	0.28600	0.28626
220	23.6517	82.698	82.62	0.28600	0.28627
311	28.0616	98.117	98.06	0.28600	0.28617
400	34.5655	120.858	120.69	0.28600	0.28640
331	38.1887	133.527	133.42	0.28600	0.28603
422	44.0158	153.901	153.73	0.28600	0.28632

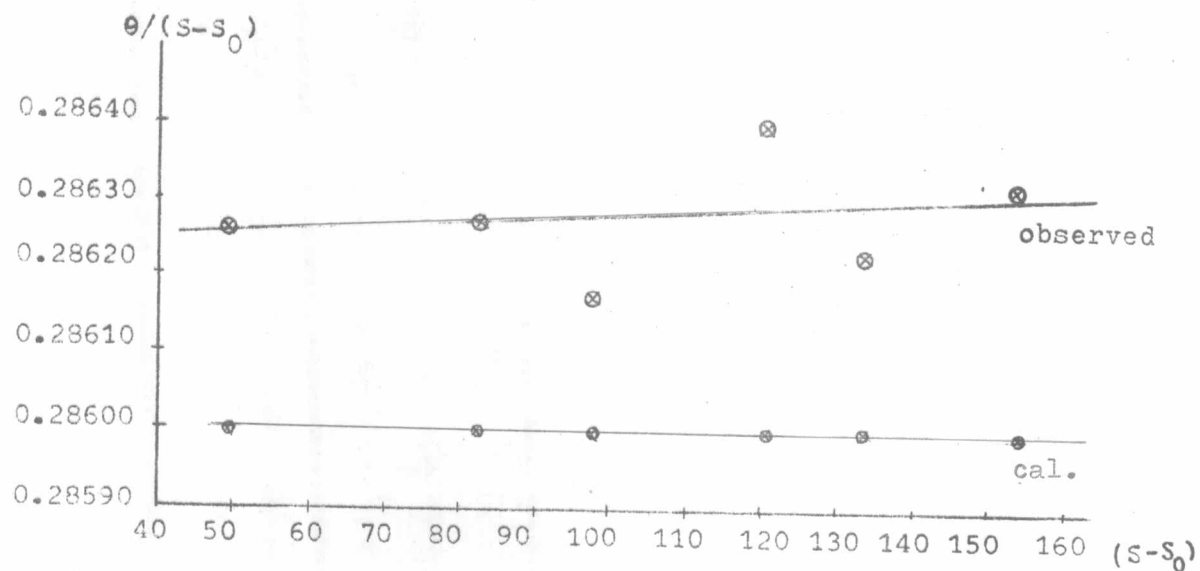


Fig.4.4.1 The graph of the observed values and those obtained from the calculation of  $\theta/(s-s_0)$  against  $(s-s_0)$  of silicon.

The data leading to the d-spacing as obtained from the powder pattern were shown in Table 4.4.2.

Table 4.4.2 The powder data obtained from powder pattern of  $\text{Nb}_3\text{As}$  where  $S_0 = 26.05 \text{ mm.}$ ,  $\lambda = 1.54051 \text{ \AA.}$

S (mm.)	S-S <sub>0</sub> (mm.)	$\theta$ (degree)	$\theta / (S-S_0)$	d = $\lambda / 2 \sin \theta$ d-spacings identified as those of lines, ( $\text{\AA}$ )		
				Unidentifiable	$\text{Nb}_7\text{As}_4$	$\text{Nb}_3\text{As}$
61.29	35.24	10.0875	0.28625		4.398	
62.88	36.83	10.5428	0.28626			4.210
78.45	52.40	15.0000	0.28626		2.976	
79.54	53.49	15.3121	0.28626		2.917	
81.56	55.51	15.8903	0.28626		2.813	
82.71	55.66	16.2195	0.28626			2.758
83.02	56.97	16.3082	0.28626		2.743	
83.24	57.19	16.3712	0.28626		2.733	
85.97	59.92	17.1527	0.28626		2.612	
87.19	61.14	17.5019	0.28626		2.561	
88.66	62.61	17.9227	0.28626			2.503
88.93	62.88	18.0003	0.28626		2.483	
89.36	63.31	18.1234	0.28627		2.476	
90.13	64.08	18.3439	0.28627			2.447
90.67	64.62	18.4984	0.28627			2.428
92.68	66.63	19.0738	0.28627		2.357	

Table 4.4.2

(continued)

S (mm.)	S-S <sub>0</sub> (mm.)	θ (degree)	θ/(S-S <sub>0</sub> )	d = λ/2sinθ d-spacings identified as those of lines , (Å)		
				Unidentifi- able	Nb <sub>7</sub> As <sub>4</sub>	Nb <sub>3</sub> As
93.59	67.54	19.3343	0.28627			2.327
93.90	67.85	19.4231	0.28627		2.316	
94.29	68.24	19.5347	0.28627			2.304
94.51	68.46	19.5977	0.28627		2.297	
94.80	68.75	19.6807	0.28627		2.287	
95.29	69.24	19.8210	0.28627		2.272	
95.95	69.90	20.0099	0.28627			2.251
97.11	71.06	20.3420	0.28627		2.216	
97.59	71.54	20.4794	0.28627		2.202	
98.95	72.90	20.8687	0.28627			2.162
99.46	73.41	21.0147	0.28627		2.143	
100.65	74.60	21.3554	0.28627			2.115
101.02	74.97	21.4613	0.28627			2.105
101.62	75.57	21.6331	0.28627		2.089	
102.94	76.89	22.0109	0.28627		2.055	
103.33	77.28	22.1226	0.28627		2.045	
103.91	77.86	22.2889	0.28627			2.031
104.38	78.33	22.4235	0.28627			2.019
105.99	79.94	22.8844	0.28627		1.981	

Table 4.4.2

(continued)

S (mm.)	S-S <sub>0</sub> (mm.)	θ (degree)	θ/(S-S <sub>0</sub> )	d = λ/2sin θ d-spacings identified as those of lines , (Å)		
				Unidentifi- able	Nb <sub>7</sub> As <sub>4</sub>	Nb <sub>3</sub> As
107.52	81.47	23.3224	0.28627		1.946	
108.07	82.02	23.4399	0.28627		1.933	
108.99	82.94	23.7432	0.28627			1.913
110.50	84.45	24.1747	0.28626			1.881
111.88	85.83	24.5688	0.28625		1.853	
113.12	87.07	24.9229	0.28625			1.828
114.39	88.34	25.2860	0.28624		1.803	
114.94	88.89	25.4429	0.28623			1.793
115.25	89.20	25.5313	0.28623		1.787	
116.00	89.95	25.7455	0.28622			1.773
116.63	90.58	25.9258	0.28622		1.762	
117.08	91.03	26.0546	0.28622		1.754	
118.84	92.79	26.5570	0.28621			1.723
119.10	93.05	26.6310	0.28620			1.718
124.44	98.39	28.1563	0.28617		1.632	
125.98	99.93	28.5990	0.28619			1.609
127.51	101.46	29.0379	0.28620		1.587	
128.40	102.35	29.2936	0.28621		1.574	
132.16	106.11	30.3740	0.28625		1.523	

Table 4.4.2

(continued)

S (mm.)	S-S <sub>0</sub> (mm.)	θ (degree)	θ / (S-S <sub>0</sub> )	d = λ / 2 sin θ d-spacings identified as those of lines (Å)		
				Unidentifiable	Nb <sub>7</sub> As <sub>4</sub>	Nb <sub>3</sub> As
132.63	106.58	30.5091	0.28626		1.517	
134.38	108.33	31.0116	0.28627		1.495	
135.01	108.96	31.1931	0.28628		1.487	
135.58	109.53	31.3568	0.28628		1.480	
137.26	111.21	31.8394	0.28630			1.460
137.69	111.64	31.9631	0.28631			1.455
139.36	113.31	32.4429	0.28632		1.436	
140.26	114.21	32.7018	0.28633		1.426	
140.60	114.55	32.7997	0.28634		1.422	
141.04	114.99	32.9262	0.28634		1.417	
141.99	115.94	33.1994	0.28635		1.407	
142.69	116.64	33.4005	0.28636		1.399	
143.05	117.00	33.5041	0.28636		1.395	
143.58	117.53	33.6565	0.28637		1.390	
144.01	117.96	33.7802	0.28637		1.385	
144.83	118.78	34.0156	0.28638		1.377	
145.53	119.48	34.2173	0.28639		1.370	
146.11	120.06	34.3840	0.28639			1.364
147.30	121.25	34.7248	0.28639			1.352



Table 4.4.2

(continued)

S (mm.)	S-S <sub>0</sub> (mm.)	θ (degree)	θ/(S-S <sub>0</sub> )	d = λ/2sinθ d-spacings identified as those of lines , (Å)		
				Unidentifi- able	Nb <sub>7</sub> As <sub>4</sub>	Nb <sub>3</sub> As
147.95	121.90	34.9085	0.28637		1.346	
149.20	123.15	35.2640	0.28635		1.334	
150.30	124.25	35.5778	0.28634		1.324	
151.25	125.20	35.8500	0.28634			1.315
151.74	125.69	35.9888	0.28633		1.311	
152.16	126.11	36.1085	0.28633		1.307	
153.06	127.01	36.3642	0.28631			1.299
156.46	130.41	37.3325	0.28627			1.270
159.96	133.91	38.3291	0.28623			1.242
160.61	134.56	38.5158	0.28624			1.237
161.60	135.55	38.8012	0.28625		1.229	
163.70	137.65	39.4023	0.28625			1.213
169.30	143.25	41.0089	0.28628		1.174	
170.21	144.16	41.2694	0.28628		1.166	
170.48	144.43	41.3474	0.28628		1.134	
172.50	146.45	41.9316	0.28632		1.153	
173.70	147.65	42.2766	0.28633			1.145
175.50	149.45	42.7875	0.28630	1.134		
176.00	149.95	42.9307	0.28630		1.131	

Table 4.4.2  
(continued)



S (mm.)	S-S <sub>0</sub> (mm.)	θ (degree)	θ/(S-S <sub>0</sub> )	d = λ/2sinθ d-spacings identified as those of lines , (Å)		
				Unidentifi- able	Nb <sub>7</sub> As <sub>4</sub>	Nb <sub>3</sub> As
177.01	150.96	43.2274	0.28635		1.125	
177.70	151.65	43.4189	0.28631			1.121
180.37	154.32	44.1849	0.28632	1.105		
181.10	155.05	44.3947	0.28632			1.101
182.72	156.67	44.8593	0.28633			1.092

The unit cell dimensions were refined by the least-squares method using CELSP program. This program was written by J. Tegenfeldt and N-O Ersson, Institute of Chemistry, University of Uppsala, Sweden, and modified for the NEAC 2200 computer by the Crystallography Group, Department of Physics, Chulalongkorn University, under the supervision of R-Liminga, Institute of Chemistry, University of Uppsala, Sweden. For this program, the maximum number of input reflections is 50.

From the diffraction lines of the photograph in Fig.3.6.3, it was found that there were more than two compounds other than the standard silicon in the sample. In the refinement, the data of the Nb<sub>3</sub>As compound was chosen to be refined. Ten diffraction lines were used to be refined first, the cell parameters obtained were :

$$a = 10.294 \pm 0.001 \text{ \AA} , \quad c = 5.1965 \pm 0.0004 \text{ \AA}$$

The other diffraction lines of  $\text{Nb}_3\text{As}$  were refined further and the results of refinement are shown in Table 4.4.3.

The final cell parameters obtained were :

$$a = 10.294 \pm 0.001 \text{ \AA} , \quad c = 5.1958 \pm 0.0007 \text{ \AA}$$

and  $v = 550.58 \text{ \AA}^3$

Table 4.4.3 The comparison of the  $\sin^2 \theta_{\text{obs}}$  and  $\sin^2 \theta_{\text{cal}}$   
( $a = 10.294 \text{ \AA}$  ,  $c = 5.1958 \text{ \AA}$  ) from powder  
diffraction data of  $\text{Nb}_3\text{As}$

hkl	$\sin^2 \theta_{\text{o}} \times 10^5$	$\sin^2 \theta_{\text{c}} \times 10^5$	d (Å)
311	7802	7797	2.758
321	9470	9477	2.503
112	9905	9911	2.447
411	11709	11717	2.251
222	13261	13270	2.115
312	14385	14390	2.031
510	14550	14559	2.019
402	17758	17750	1.828
332	18868	18870	1.773
422	19989	19990	1.723
611	22913	22916	1.609
532	27830	27829	1.460
004	35155	35163	1.299
722	38462	38468	1.242

#### 4.5 Interatomic distances and angles

The distance between two atoms in fractional co-ordinates  $(x_1, y_1, z_1)$  and  $(x_2, y_2, z_2)$ , which are at the points A and B in Fig.4.5.1, is given by the law of cosines in three dimensions,

$$l = \sqrt{(\Delta xa)^2 + (\Delta yb)^2 + (\Delta zc)^2 + 2ab \Delta x \Delta y \cos \gamma + 2ac \Delta x \Delta z \cos \beta + 2bc \Delta y \Delta z \cos \alpha} \quad \dots (4.5.1)$$

where  $a, b, c, \alpha, \beta, \gamma$  are the unit cell parameters,

$\Delta x$  is  $(x_2 - x_1)$ ,  $\Delta y$  is  $(y_2 - y_1)$  and  $\Delta z$  is  $(z_2 - z_1)$

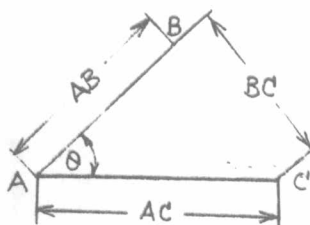


Fig. 4.5.1 A set of three atoms showing the interatomic distances AB, BC, AC and the bond angles  $\theta$ .

In Fig.4.5.1., the angle  $\theta$  subtended by bonds AB and AC can be calculated by the law of cosines,

$$\theta = \cos^{-1} \left[ \frac{(AB)^2 + (AC)^2 - (BC)^2}{2(AB)(AC)} \right] \dots (4.5.2)$$

For  $Nb_3As$ , all of the positional parameters were obtained from the last refinement so the interatomic distances and interatomic angles could be solved by using the DISTAN program as shown in Table 4.5.1

Table 4.5.1 The interatomic distances ( $\llcorner 4\text{\AA}$ ) and angles of  $\text{Nb}_3\text{As}$ .

***** * DISTANCES * *****									
ATOM1 TO ATOM2		X,Y,Z OF ATOM1			X,Y,Z OF ATOM2			DISTANCE	S.D.
NB1	NB1	-0.0568	.4639	-0.2308	.0568	.5361	.2388	2.8411	.0125
NB1	NB1	-0.0568	.4639	.7612	.0568	.5361	.2388	3.0474	.0127
NB2	NB1	.1651	.6544	-0.2698	.0568	.5361	.2388	3.1158	.0088
NB2	NB1	.1651	.6544	.7302	.0568	.5361	.2388	3.0411	.0087
NB2	NB1	.3456	.6651	.2302	.0568	.5361	.2388	3.2567	.0088
NB2	NB1	.1544	.8349	.2302	.0568	.5361	.2388	3.2371	.0054
NB3	NB1	-0.1651	.3456	.2698	.0568	.5361	.2388	3.0139	.0051
NB3	NB1	.1024	.2621	.5161	.0568	.5361	.2388	3.2017	.0080
NB3	NB1	-0.2621	.0024	.0161	.0568	.5361	.2388	3.5465	.0063
NB3	NB1	-0.1024	.7379	.4839	.0568	.5361	.2388	2.9306	.0069
NB3	NB1	.2621	.3976	-0.0161	.0568	.5361	.2388	2.8727	.0077
AS	NB1	.0419	.2714	.0167	.0568	.5361	.2388	2.9624	.0083
AS	NB1	-0.2714	.5419	.5167	.0568	.5361	.2388	3.6744	.0071
AS	NB1	-0.0419	.7286	-0.0167	.0568	.5361	.2388	2.5925	.0083
AS	NB1	.2714	.4581	.4833	.0568	.5361	.2388	2.6725	.0087
ATOM1 TO ATOM2		X,Y,Z OF ATOM1			X,Y,Z OF ATOM2			DISTANCE	S.D.
NB1	NB2	.0568	.5361	.2388	.1651	.6544	.7302	3.0411	.0087
NB1	NB2	.0568	.5361	1.2308	.1651	.6544	.7302	3.1158	.0088
NB1	NB2	.4639	.5568	.7388	.1651	.6544	.7302	3.2371	.0055
NB1	NB2	.0361	.9432	.7388	.1651	.6544	.7302	3.2567	.0056
NB1	NB2	-0.0568	.4639	.7612	.1651	.6544	.7302	3.0139	.0051
NB2	NB2	.3349	.8456	.7302	.1651	.6544	.7302	2.6329	.0084
NB2	NB2	.3456	.6651	.2302	.1651	.6544	.7302	3.1961	.0035
NB2	NB2	.3456	.6651	1.2302	.1651	.6544	.7302	3.1961	.0035
NB2	NB2	.1544	.8349	.2302	.1651	.6544	.7302	3.1961	.0035
NB2	NB2	.1544	.8349	1.2302	.1651	.6544	.7302	3.1961	.0035
NB3	NB2	-0.1651	.7379	.4839	.1651	.6544	.7302	3.1554	.0047
NB3	NB2	.2621	.3976	.9839	.1651	.6544	.7302	3.1180	.0068
AS	NB2	-0.0419	.7286	.9833	.1651	.6544	.7302	2.6177	.0062
AS	NB2	.2714	.4581	.4833	.1651	.6544	.7302	2.6322	.0078
ATOM1 TO ATOM2		X,Y,Z OF ATOM1			X,Y,Z OF ATOM2			DISTANCE	S.D.
NB1	NB3	.0568	.5361	.2388	.1024	.2621	.5161	3.2017	.0080
NB1	NB3	.0361	-0.0568	.7388	.1024	.2621	.5161	3.5465	.0061
NB1	NB3	-0.0568	.4639	.7612	.1024	.2621	.5161	2.9366	.0069
NB1	NB3	-0.0361	.0568	.2613	.1024	.2621	.5161	2.8727	.0071
NB2	NB3	.1651	.6544	.2698	.1024	.2621	.5161	3.1554	.0047
NB2	NB3	.1544	.8349	.7698	.1024	.2621	.5161	3.1180	.0082
NB3	NB3	.3976	.2379	.5161	.1024	.2621	.5161	3.0493	.0069
NB3	NB3	.2621	.3976	-0.0161	.1024	.2621	.5161	3.5066	.0113
NB3	NB3	.2621	.3976	.9839	.1024	.2621	.5161	3.2490	.0108
NB3	NB3	.2379	.1024	-0.0161	.1024	.2621	.5161	3.5066	.0113
NB3	NB3	.2379	.1024	.9839	.1024	.2621	.5161	3.5066	.0113
AS	NB3	.0419	.2714	.0167	.1024	.2621	.5161	2.8784	.0110
AS	NB3	.0419	.2714	1.0167	.1024	.2621	.5161	3.6759	.0110
AS	NB3	.2714	.4581	.4833	.1024	.2621	.5161	2.6696	.0069
AS	NB3	.2286	.0419	.4833	.1024	.2621	.5161	2.6179	.0063
ATOM1 TO ATOM2		X,Y,Z OF ATOM1			X,Y,Z OF ATOM2			DISTANCE	S.D.
NB1	AS	.0568	.5361	.2388	.0419	.2714	.0167	2.9624	.0083
NB1	AS	.0361	-0.0568	-0.2612	.0419	.2714	.0167	3.6744	.0068
NB1	AS	-0.0568	.4639	-0.2308	.0419	.2714	.0167	2.5925	.0083
NB1	AS	-0.0361	.0568	.2613	.0419	.2714	.0167	2.8725	.0074
NB2	AS	.1651	.6544	.2698	.0419	.2714	.0167	2.6177	.0062
NB2	AS	.1544	.8349	-0.2302	.0419	.2714	.0167	2.6322	.0086
NB3	AS	.1024	.2621	.5161	.0419	.2714	.0167	2.8759	.0110
NB3	AS	.1024	.2621	.5161	.0419	.2714	.0167	2.6704	.0110
NB3	AS	.2621	.3976	-0.0161	.0419	.2714	.0167	2.6179	.0073
NB3	AS	.2379	.1024	-0.0161	.0419	.2714	.0167	2.6696	.0069
AS	AS	.2714	.4581	.4833	.0419	.2714	.0167	3.8927	.0117
AS	AS	.2286	.0419	.4833	.0419	.2714	.0167	3.8927	.0115









Table 4.5.1

(continued)

Table with columns: ATJ1, ATJ2, ATJ4, ANGLE, SD, DIST1, DIST2, DIST3, XYZ OF ATOM1, XYZ OF ATOM2, XYZ OF ATOM3. It lists atom coordinates and distances for various ATJ (Atom Type) and SD (Site Distance) values.