

REFERENCES

- BABCOCK, R.H. (1950) Iron and Manganese Removal by Spaulding Precipitator. J. of The New Engl. Water Ass.
V. 64 pp 138 - 147
- GHOSH, N.O'CONNOR, J.T. and ENGELBRECHT, R.S. (1966) Precipitation of Iron in Aerated Ground Water. J. of San. Eng. Div. Proc. Am. Soc. Civ. Eng. V. 92. No.11 pp 81-89
- GHOSH, M.M. (1965) Filtration as Related to the Removal of Iron from Ground Water. Dep. of Civil Engng. University of Illinois.
- HAKER, E.A. & LYLE, W. Electrophoretic Clarification of Water. Water Resource. Inst. Tex. A and IA University Colleges Station Tex. U.S. Clearing house Fed. Sci. Tech. Inferr., P.B. Rep. 1970 No 194566
140 pp.
- HAUER (1950) Iron and Carbondioxide Removal. J. Am. Wat. Wks. Ass. V.42 pp.555 - 561
- JOHN, D.HEM. Stability Field Diagrams as Aids in Iron Chemistry Studies. J. of Am. Wat. Wks. Ass. V.53
pp.211 - 229
- JOHN, D.HEM & CROPPER, W.H. Chemistry of Iron in Natural Water- A Survey of Ferrous-Ferric Chemical : -

Equilibrium and Redox Potential. USGS Water Supply Paper No 1459-A pp.3 - 17

JOHN, D. HEM. Chemistry of Iron in Natural Water-Constraints on Dissolved Ferrous Iron Imposed by Bicarbonate, Redox Potential and pH. USGS Water Supply Paper No 1459-B pp.33 - 55

LONGLEY, J. M. & ENGELBRECHT, B. S. & MARGRAVE, G. E. (1961) Laboratory and Field Studies on Treatment of Iron-Bearing Water. J. of Am. Wat. Wks. Ass. V.54 pp.731 - 745

PHILIP, C. SINGER & WENER STUMM. (1970) The Solubility of Ferrous Iron in Carbonate Bearing Waters. J. of. AWWA V.62 pp.198 - 202

STANDARD METHOD (1968) American Public Health Ass. Inc. 1790 Broadway, New York 19, N.Y.

SUVANARAT LIMRAT (1968) An Investigation into the Use of Tray Aerator for Iron Removal. Thesis of AIT. Bangkok, Thailand. No 244

THOMAS, M. RIDDICK. Zeta Potential and Its Application to Difficult Waters. J. of Am. Wat. Wks. Ass. V.53 pp.1007 - 1030

THAMNOON SITTICHAIMANEE (1973) Water Clarification by Electrical Means. Thesis of Department of Sanitary Engineering, Chulalongkorn University, Thailand

UNITED STATES PUBLIC HEALTH SERVICE (1962) Drinking Water
Standard. J. AWWA. V.53 pp. 925 - 945

WERNER STUMM & JAMES.J.MORGAN (1970) Aquatic Chemistry - An
Introduction Emphasizing Chemisical Equilbium in
Natural Water. John Wiley and Sons. Inc. New York
London. Sydney. Toronto. Chapters 2. 5. 6. 7. 10.

WESTON,R.S. (1914) Some Experience in The Deferization and
Demanganization of Water. J. New Engl. Wat. Wks.
Ass. V.28 No 1 pp. 27 - 59

WEART J.G., & KARGRAVE,G.E., (1957) Oxidation-Reduction
potential Measurement Applied to Iron Removal J.
AWWA. V.49 pp.1223

Appendix

Procedure of Iron Testing by Delta Scientific Company
Photometric Analyzer, Model 260

Test No. 260 - 14-----Iron

The Regents and Accessories requires are as follows:

	R-101	Ammonium Acetate Buffer
	R-39	Hydrochloric Acid, Concentrated
	R-1000	Hydroxylamine Reagent
	R-36	Iron Reagent Powder
	R-102	Phenanthroline, 0.1%
2	103-A	Cylinder, Graduated, 50 ml
2	105-A	Flask, Erlenmeyer, 125 ml
2	107-A	Flask, Volumetric, 50 ml
2	110-A	Pipette, Mohr Measuring, 5 ml in 1/10 ml
	112-A	Scoop

Procedure

This is based on the standard A.P.H.A. method, 12th edition, using Phenanthroline. Where the sample is TREATED TO REMOVE INTERFERING SUBSTANCES. The calibration graph is used.

1. For total Iron, mix the sample thoroughly and measure 50 ml of mixed sample into a 125 ml Erlenmeyer flask.
2. Simultaneously prepare a reagent by adding 50 ml of distilled water to a second 125 ml Erlenmeyer flask.
3. To each Flask, add 2 ml of No R-39 Hydrochloric Acid, Concentrated. And 1 ml of No. R-100 Hydroxylamine Reagent.
4. Heat both flasks to boiling on a hot plate and reduce to volume to approximately 20 ml.
5. Cool to room temperature and transfer each to No.107-A ml Volumetric Flasks.
6. To each, add 10 ml of No.R-101 Ammonium Acetate Buffer and 3 ml of No.R-102 Phenanthroline, 0.1%

7. Dilute each flask to volume with distilled water, and mix. Let stand for five minutes.

8. Turn on the Analyzer and rotate filter selector to No.520.

9. Fill one No.260-T Square 26 mm Test Cell with untreated filtered sample.

10. Insert the test cell containing the blank into the Analyzer, making certain that one of the frosted sides faces to the front. Adjust the meter needle to 100.

11. Replace the blank with the test cell containing the treated sample, and read the meter.

12. Refer to the graph for Iron to convert the meter reading to ppm Iron (Fe)

13. The Model 260 Water Analyzer.

Table 2 Characteristics of Influent and Effluent Ground Water Sample.

Test No I B

Loading 1.50 1/min
 Energy 20 Watts
 Conductivity 3.43 micro-mho

	Raw sample	Treated sample
pH	6.35	6.5
Temperature	28°C	43°C
Iron	1.7 ppm	0.025 ppm
Hardness As CaCO ₃	380 ppm	370 ppm
Alkalinity	480 ppm	-
Dissolved O ₂	2.8 ppm	2 ppm

Percent Iron Removal = 98.55 %

Table 3 Characteristics of Influent
and Effluent Ground Water

Test No I C

Loading 2.31 1/min
 Energy 20 Watts
 Conductivity 3.43 micro-mho

	Raw sample	Treated sample
pH	6.35	6.55
Temperature	28°C	40°C
Iron	1.7 ppm	0.0625 ppm
Hardness	380 ppm	330 ppm
Alkalinity	450 ppm	-
Dissolved O ₂	2.8 ppm	2.6 ppm

Percent Iron Removal = 96.32 %



Table 4 Characteristics of Influent and Effluent Ground Water.

Test No I D

Loading 2.85 1/min
Energy used 20 Watts
Conductivity 3.43 micro-mho

	Raw sample	Treated sample
pH	6.35	6.55
Temperature	28°C	32°C
Iron	1.7 ppm	0.14 ppm
Hardness	380 ppm	380 ppm
Alkalinity	450 ppm	-
Dissolved O ₂	2.6 ppm	-

Percent Iron Removal = 91.76 %

Table 5 Characteristics of Influent and Effluent Ground Water Sample.

Test No I E

Loading 3.48 1/min
 Energy used 20 Watts
 Conductivity 3.43 micro-mho



	Raw sample	Treated sample
pH	6.35	6.45
Temperature	28°C	30°C
Iron	1.7 ppm	0.235 ppm
Hardness	380 ppm	-
Alkalinity	450 ppm	-
Dissolved O ₂	2.6 ppm	-

Percent Iron Removal = 85.88 %

Table 6 Characteristics of Influent and Effluent Ground Water Sample.

Test No II A

Loading 2.4 1/min
 Energy used 40 Watts
 Conductivity 3.96 micro-mho

	Raw sample	Treated sample
pH	7.1	7.3
Temperature	28°C	36.5°C
Iron	1.5 ppm	0.15 ppm
Hardness As CaCO ₃	405 ppm	-
Alkalinity	380 ppm	180 ppm
Dissolved O ₂	2.6 ppm	1.4 ppm

Percent Iron Removal = 90 %

Table 7 Characteristics of Influent and Effluent Ground Water Sample.

Test No II B

Loading	3.12	1/min
Energy used	40	Watts
Conductivity	3.96	micro-mho

	Raw sample	Treated sample
pH	7.1	7.2
Temperature	28°C	35°C
Iron	1.5 ppm	0.46 ppm
Hardness As CaCO ₃	405 ppm	340 ppm
Alkalinity	380 ppm	180 ppm
Dissolved O ₂	2.6 ppm	-

Percent Iron Removal = 69.33 %

Table 9 Characteristics of Influent
and Effluent Ground Water
Sample.

Test No III A

Loading 5.04 1/min
Energy used 60 Watts
Conductivity 3.96 micro-mho

	Raw sample	Treated sample
pH	7.1	7.2
Temperature	27°C	36.5°C
Iron	1.5 ppm	0.55 ppm
Hardness As CaCO ₃	405 ppm	-
Alkalinity	380 ppm	230 ppm
Dissolved O ₂	2.6 ppm	2 ppm

Percent Iron Removal = 63.33 %

Table 10 Characteristics of Influent and Effluent Ground Water Sample.

Test No III B

Loading	2.7	1/min
Energy used	60	Watts
Conductivity	3.96	micro-mho

	Raw sample	Treated sample
pH	7.1	7.05
Temperature	27°C	38°C
Iron	1.5 ppm	0.025 ppm
Hardness	405 ppm	340 ppm
Alkalinity	380 ppm	140 ppm
Dissolved O ₂	2.6 ppm	-

Percent Iron Removal = 98.33 %

Table 11 Characteristics of Influent and Effluent Ground Water Sample.

Test No III C

Loading 4.2 1/min
 Energy used 60 Watts
 Conductivity 3.96 micro-mho

	Raw sample	Treated sample
pH	7.1	6.8
Temperature	27°C	37.5°C
Iron	1.5 ppm	0.52 ppm
Hardness As CaCO ₃	405 ppm	360 ppm
Alkalinity	380 ppm	180 ppm
Dissolved O ₂	2.6 ppm	-

Percent Iron Removal = 65.33 %

Table 12 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV A

Loading 0.698 1/min
 Energy used 80 Watts
 Conductivity 4.5 micro-mho

	Synthetic sample	Treated sample
pH	7.1	7.2
Temperature	27°C	27°C
Alkalinity	106.6 ppm	103.3 ppm
Hardness	45 ppm	35 ppm
Iron	2.4 ppm	0.12 ppm

Percent Iron Removal = 95.0 %

Table 13 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV B

Loading	0.698	1/min
Energy used	80	Watts
Conductivity	4.5	micro-mho

	Synthetic sample	Treated sample
pH	7.1	7.25
Temperature	27°C	27°C
Alkalinity	106.6 ppm	81.6 ppm
Hardness	45 ppm	35 ppm
Iron	2.4 ppm	0.1 ppm

Percent Iron Removal = 95.85 %

Table 14 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV C

Loading	0.698	1/min
Energy used	100	Watts
Conductivity	4.5	micro-mho

	Synthetic sample	Treated sample
pH	7.1	7.1
Temperature	27°C	27.5°C
Alkalinity	106.6 ppm	91.75 ppm
Hardness	45 ppm	40 ppm
Iron	2.4 ppm	0.1 ppm

Percent Iron Removal = 95.83 %

Table 15 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV D

Loading 0.698 1/min
 Energy used 160 Watts
 Conductivity 4.5 micro-mho

	Synthetic sample	Treated sample
pH	7.1	6.9
Temperature	27°C	27.5°C
Alkalinity	106.6 ppm	91.75 ppm
Hardness	45 ppm	42.5 ppm
Iron	2.4 ppm	0.032 ppm

Percent Iron Removal = 98.6 %

Table 16 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV E

Loading 0.698 1/min
 Energy used 160 Watts
 Conductivity 4.5 micro-mho

	Synthetic sample	Treated sample
pH	7.1	6.9
Temperature	27°C	28°C
Alkalinity	106.6 ppm	82.5 ppm
Hardness	45 ppm	40.0 ppm
Iron	2.4 ppm	0.05 ppm

Table 16 Characteristics of Effluent
Percent Iron Removal Effluent = 97.9 %

Test No IV E

Loading 0.698 1/min
 Energy used 160 Watts
 Conductivity 4.5 micro-mho

	Synthetic sample	Treated sample
pH	7.1	6.9

Table 17 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV G

Loading	0.698	1/min
Energy	200	Watts
Conductivity	4.5	micro-mho

	Synthetic sample	Treated sample
pH	7.1	6.95
Alkalinity	106.6 ppm	90.0 ppm
Hardness	45 ppm	40 ppm
Iron	2.4 ppm	0.025 ppm
Temperature	27°C	28°C

Percent Iron Removal = 98.9 %

Table 18 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV H

Loading	0.698	1/min
Energy used	200	Watts
Conductivity	4.5	micro-mho

	Synthetic sample	Treated sample
pH	7.1	6.8
Alkalinity	106.6 ppm	77.5 ppm
Hardness	45 ppm	42.5 ppm
Iron	2.4 ppm	0.0125 ppm
Temperature	27°C	28°C

Percent Iron Removal = 99.4 %

Table 19 Characteristics of Influent and Effluent Synthetic Water Sample.

Test No IV I

Loading	0.698	1/min
Energy used	240	Watts
Conductivity	4.5	micro-mho

	Synthetic sample	Treated sample
pH	7.1	6.8
Alkalinity	106.6 ppm	36.0 ppm
Hardness	45 ppm	42.5 ppm
Iron	2.4 ppm	0.01 ppm
Temperature	27°C	28°C

Percent Iron Removal = 99.5 %

Cost Estimated

For Iron Removal by Electrical Means.

1. Natural Water

Flow rate 2.7 1/min = 162 1/hr.
 Energy used 60 Watts = 60 W-hr.
 percent of iron removal 98.33 %
 Alkalinity 380 ppm
 a unit of electrical energy is 0.75 ¢
 Cost of electrical energy per
 cu.m. is = $0.06 \frac{0.75}{0.162} = 0.77$ ¢
 approxinatty = 0.30 ¢/cu.m.

2. For a Synthetin Water

Flow rate 0.698 1/min = 41.88 1/hr.
 Energy used 160 Watts = 160 W-hr.
 Percent of iron remove 98.6 %
 Alkalinity 106.6 ppm
 41.88 1/min energy used = 0.16 k-W
 cost of electrical energy
 per cu.m. = $0.75 \frac{0.16}{0.04188}$
 = 2.926 ¢
 approxinetty 3.00 ¢/cu.m.

Note. The lower the alkalinity, the higher the energy cost.

For Iron Removal by Spaulding Precipitator by

R.H.BABCOCK (1949)

Percentage of Iron removal 95 %
 Alum used = 2.5 gpg
 Lime used = 2.5 gpg
 For one cu.m. of water used = $2.5 \frac{1,000}{3.787} = 661.5$ g.
 = 0.66 g.
 cost of alum = .20 ¢/cu.m.
 cost of lime = .05 ¢/cu.m.
 total = .25 ¢/cu.m.

Table 20 Estimated cost at various flow rate

Flow rate l/min	40 Watts		60 Watts	
	Cost ₹/cm.	% Fe removal	Cost ₹/cm.	% Fe removal
2.0	0.249	97	-	-
2.5	0.199	88	0.300	99
3.0	0.166	75	0.249	95.25
3.5	0.142	66.25	0.217	81
4.0	-	-	0.187	68

Table 21 Estimated cost at various percentage of iron removal.

% Fe removal	40 Watts		60 Watts	
	Cost ₹/cm.	Flowrate l/min	Cost ₹/cm.	Flowrate l/min
95	.23	2.15	.243	3.05
90	.204	2.40	.230	3.25
85	.192	2.60	.220	3.40
80	.178	2.80	.212	3.55
75	.166	3.0	.202	3.70

Table VI Summary of Result

	Flow Rate l/min.	Energy watts.	Alk. ppm.	Hardness ppm.	Iron ppm.	%Removal	pH	Tem. °C
Natural Raw Sample No I	-	-	480	380	1.71.7	-	6.35	28
	0.42	20	-	330	0.0125	99.26	6.6	48
	1.50	20	-	370	0.025	98.55	6.5	43
	2.31	20	-	330	0.0625	96.32	6.55	40
	2.85	20	-	380	0.14	91.76	6.55	32
	3.13	20	-	-	0.235	85.88	6.45	30
Natural Raw Sample No II	-	-	380	495	1.5	-	7.1	28
	2.1	40	180	-	-	-	-	-
	2.4	40	180	-	0.15	90	7.3	36.5
	3.12	40	180	340	0.46	69.33	7.2	35
	4.8	40	260	395	0.57	62	7.1	33
	2.7	60	140	340	0.025	98.33	7.05	38
	4.2	60	180	360	0.52	65.33	6.8	37.5
	5.04	60	230	-	0.55	63.33	7.2	36.5
Synthetic Sample.	-	-	106.6	45	2.4	-	7.1	27
	0.689	80	103.3	35	0.12	95	7.2	27
	0.689	100	91.75	40	0.10	95.85	7.1	27.5
	0.689	160	82.50	40	0.05	47.90	6.9	28
	0.689	200	77.5	42.5	0.0125	99.4	6.8	28.5

VITA

The author, Montri Uthaiphattrakul, was born on May 23, 1949 in Bangkok, Thailand. He received a Bachelor's Degree of Engineering in Sanitary Engineering, Chulalongkorn University in 1971.