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## APPENDIX

## APPENDIX

### Sensitivity on UV Spectrometric Method

The functional relationship between the quantity measured absorbance, A and the quantity sought (the analyte concentration, C) is known as Beer's Law and can be written as ;

$$A = abC$$

where

- a is a proportionality constant called specific absorptivity
- b is absorption path length, expressed in centimetres
- C is the concentration, in grams per litre

When

- C is expressed in moles per litre and
- b is in centimetres

the proportionality constant is called "the molar absorptivity", and

$$A = \epsilon bC$$

Where

$\epsilon$  has the unit of  $L \text{ cm}^{-1} \text{ mol}^{-1}$

Sensitivity on UV spectrometric method can be expressed in term of molar absorptivity,  $\epsilon$  ;

$$\epsilon (L \text{ cm}^{-1} \text{ mol}^{-1}) = \frac{A}{bC}$$

or specific absorptivity,  $a$  ;

$$a (mL \text{ g}^{-1} \text{ cm}^{-1}) = \frac{\epsilon}{\text{atomic weight} \times 1000}$$

or Sandell sensitivity,  $S$  ;

$$S (\mu\text{g cm}^{-2}) = 10^{-3}/a$$

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