

REFERENCES

- Adak, A., Pal, A., and Bandyopadhyay, M. (2006) Removal of phenol from water environment by surfactant-modified alumina through adsolubilization. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 277(1-3), 63-68.
- Chandar, P., Somasundaran, P., and Turro, N.J. (1987) Fluorescence probe studies on the structure of the adsorbed layer of dodecyl sulfate at the alumina-water interface. Journal of Colloid and Interface Science, 117(1), 31-46.
- Dickson, J. and O'Haver, J. (2002) Adsorption of naphthalene and α -naphthol in C_nTAB admicelles. Langmuir, 18(24), 9171-9176.
- Esumi, K. (2001) Adsorption and adsolubilization of surfactants on titanium dioxides with functional groups. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 176(1), 25-34.
- Esumi, K., Goino, M., and Koide, Y. (1996) Adsorption and adsolubilization by monomeric, dimeric, or trimeric quaternary ammonium surfactant at silica/water interface. Journal of Colloid and Interface Science, 183(2), 539-545.
- Esumi, K., Goino, M., and Koide, Y. (1996) The effect of added salt on adsorption and adsolubilization by a gemini surfactant on silica. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 118(1-2), 161-166.
- Esumi, K., Maedomari, N., and Torigoe, K. (2000) Mixed surfactant adsolubilization of 2-naphthol on alumina. Langmuir, 16(24), 9217-9220.
- Esumi, K., Maedomari, N., and Torigoe, K. (2001) Adsolubilization of 2-naphthol by binary mixtures of cationic and nonionic surfactants on silica. Langmuir, 17(23), 7350-7354.
- Esumi, K., Nagahama, T., and Meguro, K., (1991) Characterization of cationic surfactant adsorbed layer on silica. Colloids and Surfaces, 57(1), 149-160.
- Esumi, K., Sakai, K., and Torigoe, K. (2000) Reexamination of 2-naphthol adsolubilization on alumina with sodium dodecyl sulfate adsorption. Journal of Colloid and Interface Science, 224(1), 198-201.

- Esumi, K. and Yamamoto, S. (1998) Adsorption of sodium dodecyl sulfate on hydrotalcite and adsolubilization of 2-naphthol. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 137(1-3), 385-388.
- Esumi, K. and Yamanaka, Y. (1995) Interaction between sodium dodecyl poly(oxyethylene) sulfate and alumina surface in aqueous solution. Journal of Colloid and Interface Science, 172(1), 116-120.
- Esumi, K., Yoshida, K., Torigoe, K., and Koide, Y. (1999) Sorption of 2-naphthol and copper ions by cationic surfactant-adsorbed laponite. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 160(3), 247-250.
- Gao, Y., Yue, C., Lu, S., Gu, W., and Gu, T. (1984) Adsorption from mixed solutions of triton X-100 and sodium *n*-alkyl sulfates on silica gel. Journal of Colloid and Interface Science, 100(2), 581-583.
- Harwell, J.H., (1991) Factors effecting surfactant performance in groundwater remediation application. Sabatini, D.A., and Knox, R.C., (Eds.). (1991) Colloid Interfacial and Surfactant Phenomena. New York: Marcel Dekker.
- Huang, Z., Yan, Z., and Gu, T. (1989) Mixed adsorption of cationic and anionic surfactants from aqueous solution on silica gel. Colloids and Surfaces, 36(3), 353-358.
- Kitiyanan, B., O'Haver, J.H., Harwell, J.H., and Osuwan, S. (1996) Adsolubilization of styrene and isoprene in cetyltrimethylammonium bromide admicelle on precipitated silica. Langmuir, 12(9), 2162-2168.
- Levitz, P.E. and Damme, H.V. (1986) Fluorescence Decay study of the adsorption of nonionic surfactants at the solid-liquid interface 2. Influence of polar chain length. Physical Chemistry, 90(7), 1302-1310.
- Li, L., Wang, L., Du, X., Lu, Y., and Yang, Z. (2007) Adsolubilization of dihydroxybenzenes into CTAB layers on silica particles. Journal of Colloid and Interface Science, 315(2), 671-677.
- Malmsten, M., Linse, P., and Cosgrove, T. (1992) Adsorption of PEO-PPO-PEO block copolymers at silica. Macromolecules, 25(9), 2474-2481.
- Montarges, E., Moreau, A., and Michot, L.J. (1998) Removing of organic toxicants from water by Al₁₃-pluronic modified clay. Applied Clay Science, 13(3), 165-185.

- Myers, D. (2006) Surfactant Science and Technology. New Jersey: Wiley-Interscience.
- Nuysink, J. and Koopal, L.K. (1982) The Effect of Polyethylene Oxide Molecular Weight on Determination of Its Concentration in Aqueous Solutions. Talanta, 29, 495-501.
- Parida, S.K., Dash, S., Patel, S., and Mashra, B.K. (2006) Adsorption of organic molecules on silica surface. Advances in Colloid and Interface Science, 121(1-3), 77-110.
- Partyka, S., Zaini, S., Lindheimer, M., and Brun, B. (1984) The adsorption of non-ionic surfactants on a silica gel. Colloids and Surfaces, 12, 255-270.
- Pradubmook, T., O'Haver, J.H., Malakul, P., and Harwell, J.H. (2003) Effect of pH on adsolubilization of toluene and acetophenone into adsorbed surfactant on precipitated silica. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 224(1-3), 93-98.
- Rosen, M.J. (1989) Surfactant and Interfacial Phenomena. New York: Wiley-Interscience.
- Scamehorn, J.F., Schechter, R.S., and Wade, W.H. (1982) Adsorption of surfactants on mineral oxide surfaces from aqueous solutions. Colloid and Interface Science, 85(2), 463-477.
- Scamehorn, J.F., Harwell, J.H. Surfactant-based treatment of process streams. Wasan, D.T., Ginn, M.E., and Shah, D.O. (Eds.). (1988) Surfactant in Chemical/Process Engineering. New York: Marcel Dekker.
- Schroën, C.G.P.H., Stuart, M.A.C., Maarschalk, K.V., Padt, A., and Riet, K. (1995) Influence of preadsorbed block copolymers on protein adsorption: surface properties, layer thickness, and surface coverage. Langmuir, 11(8), 3068-3074.
- Shar, J.A., Obey, T.M., and Cosgrove, T. (1998) Adsorption studies of polyethers Part 1. Adsorption onto hydrophobic surfaces. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 136(1-2), 21-33.
- Shar, J.A., Obey, T.M., and Cosgrove, T. (1999) Adsorption studies of polyethers Part II. Adsorption onto hydrophilic surfaces. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 150(1-3), 15-23.

- Schick, M.J. (1987) Nonionic Surfactants Physical Chemistry. New York: Marcel Dekker Inc.
- Somasundaran, P. and Huang, L. (2000) Adsorption/aggregation of surfactants and their mixtures at solid-liquid interfaces. Advances in Colloid and Interface Science, 88(1-2), 179-208.
- Tadros, T.F. (2005) Applied Surfactants Principles and Applications. Mörlenbach: Wiley-VCH.
- Tan, Y. and O'Haver, J.H. (2004) Lipophilic linker impact on adsorption of and styrene adsolubilization in polyethoxylated octylphenols. Colloids and Surfaces A: Physicochemical and Engineering Aspects, 232(2-3), 101-111.
- Tsurumi, D., Sakai, K., Yoshimura, T., and Esumi, K. (2006) Adsolubilization of 2-naphthol into adsorbed layers of triblock PEO-PPO-PEO copolymers on hydrophobic silica particles. Journal of Colloid and Interface Science, 302(1), 82-86.
- Tsurumi, D., Yoshimura, T., and Esumi, K. (2006) Adsolubilization of 2-naphthol into adsorbed layer of PEO-PPO-PEO triblock copolymers on hydrophilic silica. Journal of Colloid and Interface Science, 297(2), 465-469.
- Wattanaphan, P. (2008) Removal of Aromatic Organic Compounds by Silica Modified with EO/PO-Based Block Copolymers. M.S. Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Yeskie, M.A. and Harwell, J.H. (1988) On the structure of aggregates of adsorbed surfactant: The surface charge density at the hemimicelle/admicelle transition. Physical Chemistry, 92(8), 2346-2352.
- Yu, C.C., Wong, D.W., and Lobban, L.L. (1992) Catalysis of the hydrolysis of trimethyl orthobenzoate by adsorbed sodium dodecyl sulfate. Langmuir, 8(10), 2582-2584.

APPENDICES

Appendix A Adsorption of Surfactants onto Hydrophobic Silica

The adsorption isotherms of EO/PO triblock copolymer surfactants, which are L31, P123, L64, 17R2, 25R4, and 10R5, onto hydrophobic silica at 29°C, are shown in table A1, A2, A3, A4, A5, and A6, respectively.

Weight of silica = 0.15 g

Volume of copolymer surfactant solution = 15 ml

Table A1 The adsorption isotherm of L31 onto hydrophobic silica

Initial concentration (mM)	Equilibrium concentration (mM)	Adsorbed surfactant (mmol/g of silica)
0.10	0.0001	0.0099
0.45	0.0904	0.0363
0.70	0.1446	0.0536
1.00	0.4607	0.0536
1.10	0.5444	0.0544
1.20	0.6151	0.0564
1.30	0.7122	0.0557
1.40	0.7500	0.0610
1.60	0.9621	0.0626
2.05	1.2670	0.0773
2.30	1.3709	0.0899
2.70	1.6961	0.1028
2.95	1.9634	0.0990
3.20	2.0885	0.1088

Table A2 The adsorption isotherm of P123 onto hydrophobic silica

Initial concentration (mM)	Equilibrium concentration (mM)	Adsorbed surfactant (mmol/g of silica)
0.10	0.0001	0.0099
0.50	0.0003	0.0499
1.00	0.0239	0.0974
1.10	0.0520	0.1047
1.20	0.1770	0.1023
1.25	0.2583	0.0989
1.30	0.3335	0.0963
1.40	0.4604	0.0939
1.50	0.6800	0.0818
1.75	0.9216	0.0826

Table A3 The adsorption isotherm of L64 onto hydrophobic silica

Initial concentration (mM)	Equilibrium concentration (mM)	Adsorbed surfactant (mmol/g of silica)
0.10	0.0001	0.0099
0.50	0.0002	0.0499
1.00	0.1590	0.0838
1.25	0.4911	0.0755
1.50	0.5923	0.0905
1.75	1.0102	0.0738
2.00	1.1834	0.0814
2.50	1.6185	0.0879
3.00	2.0386	0.0959
3.50	2.6267	0.0870

Table A4 The adsorption isotherm of 17R2 onto hydrophobic silica

Initial concentration (mM)	Equilibrium concentration (mM)	Adsorbed surfactant (mmol/g of silica)
0.05	0.0001	0.0049
0.10	0.0099	0.0090
0.30	0.0401	0.0259
0.50	0.2238	0.0276
0.75	0.3332	0.0415
1.00	0.5464	0.0451
1.25	0.7714	0.0478
1.50	0.9521	0.0547
1.75	1.2235	0.0525
2.00	1.4598	0.0536
2.25	1.6777	0.0571

Table A5 The adsorption isotherm of 25R4 onto hydrophobic silica

Initial concentration (mM)	Equilibrium concentration (mM)	Adsorbed surfactant (mmol/g of silica)
0.10	0.0458	0.0054
0.30	0.1213	0.0178
0.50	0.1845	0.0314
0.75	0.4350	0.0314
1.00	0.6754	0.0323
1.25	0.9139	0.0335
1.50	1.0856	0.0412
1.75	1.3843	0.0364

Table A6 The adsorption isotherm of 10R5 onto hydrophobic silica

Initial concentration (mM)	Equilibrium concentration (mM)	Adsorbed surfactant (mmol/g of silica)
0.10	0.0001	0.0099
0.30	0.0134	0.0285
0.75	0.4407	0.0310
1.25	0.8820	0.0367
1.75	1.3091	0.0440
2.25	1.7588	0.0491
2.75	2.2219	0.0529
3.00	2.4545	0.0544
3.25	2.6910	0.0557
3.50	2.9456	0.0552

Appendix B Adsolubilization of Organic Compounds

The adsolubilization of phenol in the adsorbed layer of block copolymer surfactants, which are L31, P123, L64, 17R2, 25R4, and 10R5, are shown in table B1, B2, B3, B4, B5, and B6 respectively.

Weight of silica = 0.15 g

Volume of phenol-surfactant solution = 15 ml

Aqueous solubility limit of phenol = 882.97 mM

Table B1 The adsolubilization of phenol in an adsorbed layer of hydrophobic silica modified with Pluronics L31

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
50	0.0541	0.2210	4.4200
100	0.1087	0.4027	8.0540
150	0.1605	0.8254	16.5077
200	0.2179	0.7562	15.1232
300	0.3257	1.2371	24.7414
400	0.4312	1.9145	38.2905
500	0.5399	2.3232	46.4641

Table B2 The adsolubilization of phenol in an adsorbed layer of hydrophobic silica modified with Pluronics P123

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
50	0.0354	1.8695	20.7722
100	0.0815	1.8005	20.0056
150	0.1256	3.8876	43.1952
200	0.1719	4.8086	53.4292
300	0.2655	6.5410	72.6778
400	0.3497	9.0933	101.0363
500	0.4320	11.8032	131.1469

Table B3 The adsolubilization of phenol in an adsorbed layer of hydrophobic silica modified with Pluronics L64

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
50	0.0426	1.2349	15.4363
100	0.0818	2.7701	34.6263
150	0.1382	2.7885	34.8558
200	0.1803	4.0645	50.8060
300	0.2759	5.6334	70.4174
400	0.3625	7.9545	99.4309
500	0.4441	10.7690	134.6120

Table B4 The adsolubilization of phenol in an adsorbed layer of hydrophobic silica modified with Pluronics 17R2

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
50	0.0529	0.3279	16.3950
100	0.1042	0.8256	41.2800
150	0.1607	0.8066	40.3283
200	0.2172	0.8202	41.0094
300	0.3246	1.3360	66.7983
400	0.4273	2.2553	112.7649
500	0.5326	2.9635	148.1741

Table B5 The adsolubilization of phenol in an adsorbed layer of hydrophobic silica modified with Pluronics 25R4

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
50	0.0484	0.7217	24.0567
100	0.0937	1.7205	57.3500
150	0.1517	1.5979	53.2644
200	0.1997	2.3658	78.8599
300	0.2953	3.9099	130.3308
400	0.3921	5.3578	178.5925
500	0.4862	7.0484	234.9452

Table B6 The adsolubilization of phenol in an adsorbed layer of hydrophobic silica modified with Pluronics 10R5

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
50	0.0528	0.3344	6.6880
100	0.0936	1.7262	34.5240
150	0.1414	2.5058	50.1153
200	0.1830	3.8251	76.5001
300	0.2783	5.4193	108.3863
400	0.3655	7.7136	154.2713
500	0.4533	9.9209	198.4181

The adsolubilization of 2-naphthol in the adsorbed layer of block copolymer surfactants, which are L31, P123, L64, 17R2, 25R4, and 10R5, are shown in table B7, B8, B9, B10, B11, and B12 respectively.

Weight of silica = 0.15 g

Volume of 2-naphthol-surfactant solution = 15 ml

Aqueous solubility limit of 2-naphthol = 5.1389 mM

Table B7 The adsolubilization of 2-naphthol in an adsorbed layer of hydrophobic silica modified with Pluronics L31

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.5	0.0612	0.0185	0.3709
1.0	0.1161	0.0402	0.8049
1.5	0.1686	0.0630	1.2604
2.0	0.2144	0.0896	1.7911
2.5	0.2555	0.1184	2.3675
3.0	0.2972	0.1468	2.9355
3.5	0.3204	0.1852	3.7045
4.0	0.3812	0.2032	4.0636
4.5	0.3930	0.2470	4.9406
5.0	0.4407	0.2722	5.4442
6.0	0.5333	0.3247	6.4949
6.9238	0.5971	0.3841	7.6816

Table B8 The adsolubilization of 2-naphthol in an adsorbed layer of hydrophobic silica modified with Pluronics P123

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.5	0.0172	0.0411	0.4570
1.0	0.0270	0.0859	0.9542
1.5	0.0351	0.1315	1.4615
2.0	0.0428	0.1774	1.9713
2.5	0.0498	0.2233	2.4811
3.0	0.0575	0.2699	2.9990
3.5	0.0648	0.3161	3.5125
4.0	0.0732	0.3617	4.0191
4.5	0.0841	0.4054	4.5045
5.0	0.0895	0.4516	5.0177
6.0	0.1040	0.5448	6.0535
7.0	0.1094	0.6424	7.1375
8.0	0.1297	0.7302	8.1139
8.9237	0.1439	0.8160	9.0666

Table B9 The adsolubilization of 2-naphthol in an adsorbed layer of hydrophobic silica modified with Pluronics L64

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.5	0.0270	0.0360	0.4499
1.0	0.0497	0.0743	0.9289
1.5	0.0636	0.1167	1.4587
2.0	0.0755	0.1603	2.0037
2.5	0.0939	0.2007	2.5085
3.0	0.1035	0.2457	3.0712
3.5	0.1240	0.2854	3.5678
4.0	0.1413	0.3257	4.0708
4.5	0.1581	0.3668	4.5853
5.0	0.1649	0.4132	5.1651
6.0	0.1763	0.5074	6.3421
7.0	0.1821	0.6040	7.5500
7.94	0.1910	0.6954	8.6926

Table B10 The adsolubilization of 2-naphthol in an adsorbed layer of hydrophobic silica modified with Pluronics 17R2

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.5	0.0516	0.0235	1.1727
1.0	0.0847	0.0563	2.8147
1.5	0.1292	0.0835	4.1764
2.0	0.1683	0.1131	5.6564
2.5	0.2002	0.1464	7.3207
3.0	0.2402	0.1764	8.8185
3.5	0.2741	0.2082	10.4117
4.0	0.3107	0.2396	11.9784
4.5	0.3517	0.2683	13.4141
5.0	0.3949	0.2963	14.8174
6.0	0.4285	0.3785	18.9260
7.0	0.5308	0.4260	21.3000
7.8567	0.5966	0.4776	23.8818

Table B11 The adsolubilization of 2-naphthol in an adsorbed layer of hydrophobic silica modified with Pluronics 25R4

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.5	0.0731	0.0124	0.4132
1.0	0.1149	0.0409	1.3618
1.5	0.1428	0.0765	2.5515
2.0	0.1840	0.1052	3.5081
2.5	0.2139	0.1398	4.6591
3.0	0.2551	0.1686	5.6207
3.5	0.2647	0.2131	7.1036
4.0	0.2946	0.2481	8.2712
4.5	0.3166	0.2864	9.5479
5.0	0.3453	0.3221	10.7357
6.0	0.3707	0.4084	13.6129
7.2466	0.3904	0.5342	17.8066

Table B12 The adsolubilization of 2-naphthol in an adsorbed layer of hydrophobic silica modified with Pluronics 10R5

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.5	0.0287	0.0351	0.7030
1.0	0.0598	0.0691	1.3816
1.5	0.0752	0.1109	2.2172
2.0	0.0865	0.1551	3.1011
2.5	0.1108	0.1925	3.8499
3.0	0.1298	0.2325	4.6505
3.5	0.1441	0.2751	5.5022
4.0	0.1712	0.3115	6.2299
4.5	0.1825	0.3549	7.0984
5.0	0.1911	0.4007	8.0146
6.0	0.2052	0.4933	9.8651
7.3571	0.2162	0.6119	12.2382

The adsolubilization of naphthalene in the adsorbed layer of block copolymer surfactants, which are L31, P123, L64, 17R2, 25R4, and 10R5, are shown in table B13, B14, B15, B16, B17, and B18 respectively.

Weight of silica = 0.15 g

Volume of naphthalene-surfactant solution = 15 ml

Aqueous solubility limit of naphthalene = 0.2344 mM

Table B13 The adsolubilization of naphthalene in an adsorbed layer of hydrophobic silica modified with Pluronics L31

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.0250	0.0318	0.0017	0.0350
0.0500	0.0677	0.0034	0.0681
0.0750	0.1065	0.0050	0.0997
0.1000	0.1280	0.0070	0.1395
0.1250	0.1800	0.0083	0.1651
0.1500	0.2236	0.0097	0.1944
0.1750	0.2483	0.0116	0.2327
0.2000	0.3204	0.0124	0.2486
0.2250	0.4013	0.0130	0.2608
0.3000	0.5285	0.0176	0.3510
0.3500	0.6863	0.0189	0.3772
0.4105	0.8311	0.0215	0.4303

Table B14 The adsolubilization of naphthalene in an adsorbed layer of hydrophobic silica modified with Pluronics P123

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.0250	0.0192	0.0020	0.0227
0.0500	0.0457	0.0039	0.0435
0.0750	0.0792	0.0056	0.0625
0.1000	0.1029	0.0076	0.0842
0.1250	0.1206	0.0097	0.1072
0.1500	0.1416	0.0117	0.1295
0.1750	0.1497	0.0140	0.1553
0.2000	0.1910	0.0155	0.1719
0.2250	0.2197	0.0173	0.1922
0.3000	0.2934	0.0231	0.2563
0.3000	0.3619	0.0314	0.3488
0.5000	0.4196	0.0401	0.4452
0.6000	0.5005	0.0481	0.5339
0.6715	0.5754	0.0535	0.5947

Table B15 The adsolubilization of naphthalene in an adsorbed layer of hydrophobic silica modified with Pluronics L64

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.0250	0.0355	0.0017	0.0208
0.0500	0.0558	0.0037	0.0461
0.0750	0.0672	0.0059	0.0738
0.1000	0.0708	0.0083	0.1041
0.1250	0.0930	0.0103	0.1287
0.1500	0.1056	0.0125	0.1560
0.1750	0.1255	0.0145	0.1815
0.2000	0.1508	0.0164	0.2051
0.2250	0.1620	0.0186	0.2328
0.3000	0.2526	0.0240	0.2998
0.4000	0.3106	0.0326	0.4079
0.5000	0.4583	0.0391	0.4892
0.6000	0.5076	0.0479	0.5983
0.6745	0.6248	0.0527	0.6584

Table B16 The adsolubilization of naphthalene in an adsorbed layer of hydrophobic silica modified with Pluronics 17R2

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.0250	0.0443	0.0015	0.0728
0.0500	0.0620	0.0035	0.1767
0.0750	0.0945	0.0053	0.2633
0.1000	0.1031	0.0076	0.3778
0.1250	0.1237	0.0096	0.4782
0.1500	0.1437	0.0116	0.5800
0.1750	0.1598	0.0137	0.6863
0.2000	0.1823	0.0157	0.7843
0.2250	0.2162	0.0174	0.8701
0.3000	0.2493	0.0241	1.2047
0.3500	0.3004	0.0278	1.3924
0.4250	0.3378	0.0345	1.7228

Table B17 The adsolubilization of naphthalene in an adsorbed layer of hydrophobic silica modified with Pluronics 25R4

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.0250	0.0202	0.0020	0.0675
0.0500	0.0422	0.0040	0.1333
0.0750	0.0699	0.0058	0.1949
0.1000	0.1066	0.0075	0.2490
0.1250	0.1108	0.0099	0.3300
0.1500	0.1342	0.0118	0.3937
0.1750	0.1642	0.0136	0.4539
0.2000	0.1766	0.0158	0.5271
0.2250	0.1839	0.0181	0.6044
0.3000	0.2237	0.0246	0.8212
0.3405	0.2547	0.0280	0.9346

Table B18 The adsolubilization of naphthalene in an adsorbed layer of hydrophobic silica modified with Pluronics 10R5

Initial concentration (mM)	Reduced bulk concentration	Adsolubilized amount (mmol/g of silica)	Adsolubilization ratio
0	0	0	0
0.0250	0.0183	0.0021	0.0413
0.0500	0.0310	0.0043	0.0851
0.0750	0.0499	0.0063	0.1264
0.1000	0.0684	0.0084	0.1676
0.1250	0.0888	0.0104	0.2075
0.1500	0.1024	0.0125	0.2510
0.1750	0.1260	0.0145	0.2901
0.2000	0.1405	0.0166	0.3329
0.2250	0.1545	0.0188	0.3770
0.3000	0.2233	0.0247	0.4939
0.4000	0.3673	0.0313	0.6266
0.5073	0.4827	0.0393	0.7854

CURRICULUM VITAE

Name: Mr. Pattarit Sahasyodhin

Date of Birth: May 14, 1987

Nationality: Thai

University Education:

2005-2009 Bachelor Degree of Industrial Chemistry (1st Class Honours), Faculty of Science, King Mongkut's Institute of Technology Ladkrabang, Bangkok, Thailand

Work Experience:

2008 Position: Chemist (Student Internship)
Company name: PTT Public Company Limited

Proceedings:

1. Sahasyodhin, P.; Malakul, P.; O'Haver, J. H.; Nithitanakul, M. (2011, April 26) Adsolubilization of Organic Compounds by Hydrophobic Silica Modified with EO/PO Triblock Copolymers. Proceedings of the 2nd Research Symposium on Petroleum, Petrochemicals, and Advanced Materials and The 17th PPC Symposium on Petroleum, Petrochemicals, and Polymers, Bangkok, Thailand.