REFERENCES

- Abdolahi, A., Hamzah, E., Ibrahim, Z. and Hashim, S. (2013). Synthesis and characterization of high-quality polyaniline nanofibres. <u>High Performance</u> <u>Polymers.</u> 25(2), 236-242.
- Benjamin S., Jingxian Yu., Joseph G. and Jamine S. (2008). Patterned polyaniline
 & carbon nanotube-polyaniline composites on silicon. <u>The Royal Society</u> of Chemistry, 5, 164-172.
- Bhadra, S., Kim, N.H. and Lee, J.H. (2010). Synthesis of water soluble sulfonated polyaniline and determination of crystal structure. <u>Journal of Applied</u> <u>Polymer Science</u>, 117(4), 2025-2035.
- Campos, T.L.A., Kersting, D.F. and Ferreira, C.A. (1999). Chemical synthesis of polyaniline using sulphanilic acid as dopant agent into the reactional medium. <u>Surface and Coatings Technology</u>, 122(1), 3-5.
- Cho, W., Park, S.J. and Kim, S. (2011). Effect of monomer concentration on interfacial synthesis of platinum loaded polyaniline nanocomplex using poly(styrene sulfonic acid). <u>Synthetic Metals.</u> 161(21-22), 2446-2450.
- Chu, C.C., Wang, Y.W., Wang, L. and Ho, T.I. (2005). Synthesis and characterization of novel conductive star polymers containing PSS/PANI arms. <u>Synthetic Metals</u>, 153(1-3), 321-324.
- Detsri, E. and Dubas, S.T. (2013). Interfacial polymerization of polyaniline and its layer-by-layer assembly into polyelectrolytes multilayer thin-films. Journal of Applied Polymer Science, 128(1), 558-565.
- Gheno, G., De Souza Basso, N.R. and Hübler, R. (2011). Polyaniline/graphite nanocomposites: Synthesis and characterization. <u>Macromolecular</u> <u>Symposia.</u> 299-300(1), 74-80.
- Hopkins, A.R., Sawall, D.D., Villahermosa, R.M. and Lipeles, R.A. (2004).
 Interfacial synthesis of electrically conducting polyaniline nanofiber composites. <u>Thin Solid Films</u>, 469-470(SPEC. ISS.), 304-308.
- Hussain, S.T., Abbas, F., Kausar, A. and Khan, M.R. (2013). New polyaniline/polypyrrole/polythiophene and functionalized multiwalled

carbon nanotube-based nanocomposites: Layer-by-layer in situ polymerization. <u>High Performance Polymers</u>, 25(1), 70-78.

- Kuo, C.W. and Wen, T.C. (2008). Dispersible polyaniline nanoparticles in aqueous poly(styrenesulfonic acid) via the interfacial polymerization route.
 <u>European Polymer Journal</u>, 44(11), 3393-3401.
- Li, L., Ferng, L., Wei, Y., Yang, C. and Ji, H.F. (2012). Effects of acidity on the size of polyaniline-poly(sodium 4-styrenesulfonate) composite particles and the stability of corresponding colloids in water. Journal of Colloid and Interface Science, 381(1), 11-16.
- Mu, S. (2010). Nanostructured polyaniline synthesized using interface polymerization and its redox activity in a wide pH range. <u>Synthetic</u> <u>Metals.</u> 160(17-18), 1931-1937.
- Potyrailo, R.A., Nagraj, N., Tang, Z., Mondello, F.J., Surman, C. and Morris, W. (2012). Battery-free radio frequency identification (RFID) sensors for food quality and safety. Journal of Agricultural and Food Chemistry, 60(35), 8535-8543.
- Rubinger, C.P.L., Costa, L.C., Faez, R., Martins, C.R. and Rubinger, R.M. (2009). Hopping conduction on PAni/PSS blends. <u>Synthetic Metals</u>, 159(5-6), 523-527.
- Shaktawat, V., Saxena, N.S. and Sharma, K. (2011). Study of the structure and mechanical properties of pure and doped polyaniline. <u>Phase Transitions</u>, 84(3), 215-224.
- Srivastava, S., Sharma, S.S., Kumar, S., Agrawal, S., Singh, M. and Vijay, Y.K. (2009). Characterization of gas sensing behavior of multi-walled carbon nanotube polyaniline composite films. <u>International Journal of Hydrogen</u> <u>Energy</u>, 34(19), 8444-8450.
- Sun, L., Shi, Y., Chu, L., Xu, X. and Liu, J. (2011). Preparation of polyaniline coated polystyrene-poly(styrene-co-sodium 4-styrenesulfonate) microparticles and the further fabrication of hollow polyaniline microspheres. Journal of Applied Polymer Science, 126(3), 870-876.

- Tang, Q., Wu, J., Sun, X., Li, Q. and Lin, J. (2009). Layer-by-layer self-assembly of conducting multilayer film from poly(sodium styrenesulfonate) and polyaniline. Journal of Colloid and Interface Science, 337(1), 155-161.
- Wang, Y. and Levon, K. (2012). Influence of dopant on electroactivity of polyaniline. <u>Macromolecular Symposia</u>, 317-318(1), 240-247.
- Zhang, F., Halverson, P.A., Lunt, B. and Linford, M.R. (2006). Wet spinning of pre-doped polyaniline into an aqueous solution of a polyelectrolyte. <u>Synthetic Metals.</u> 156(14-15), 932-937.

APPENDICIES

Appendix A Effect of APS Concentration on PANI Polymerization



Figure A1 UV-vis spectrum of PANI-PSS at various APS concentrations.



Figure A2 UV-vis spectrum of PANI-CoPSS 3:1 at various APS concentrations.



Figure A3 UV-vis spectrum of PANI-CoPSS 3:1 at various APS concentrations.





Figure B1 UV-vis spectrum of PANI with various PSS concentrations.



Figure B2 UV-vis spectrum of PANI with various CoPSS 3:1 concentrations.



Figure B3 UV-vis spectrum of PANI with various CoPSS 1:1 concentrations.





Figure C1 UV-vis spectrum of PANI-PSS with various pH.



Figure C2 UV-vis spectrum of PANI-CoPSS 3:1 with various pH.



Figure C3 UV-vis spectrum of PANI-CoPSS 1:1 with various pH.



Appendix D Effect of pH, Salt and Dipping Time on PANI monolayer

Figure D1 UV-vis spectrum of PANI-PSS monolayer at pH 4 and 0.5 M NaCl.



Figure D2 UV-vis spectrum of PANI-PSS monolayer at pH 4 and 1 M NaCL



Figure D3 UV-vis spectrum of PANI-PSS monolayer at pH 4 and 2 M NaCl.



Figure D4 UV-vis spectrum of PANI-PSS monolayer at pH 7 and 0.5 M NaCl.



Figure D5 UV-vis spectrum of PANI-PSS monolayer at pH 7 and 1 M NaCl.



Figure D6 UV-vis spectrum of PANI-PSS monolayer at pH 7 and 2 M NaCl.



Figure D7 UV-vis spectrum of PANI-PSS monolayer at pH 9 and 0.5 M NaCl.



Figure D8 UV-vis spectrum of PANI-PSS monolayer at pH 9 and 1 M NaCl.



Figure D9 UV-vis spectrum of PANI-PSS monolayer at pH 9 and 2 M NaCl.



Figure D10 UV-vis spectrum of PANI-PSS monolayer at pH 12 and 0.5 M NaCl.



Figure D11 UV-vis spectrum of PANI-PSS monolayer at pH 12 and 1 M NaCl.



Figure D12 UV-vis spectrum of PANI-PSS monolayer at pH 12 and 2 M NaCl.





Figure E1 UV-vis spectrum of PANI-PSS multilayer at pH 4 and 1 M NaCl.



Figure E2 UV-vis spectrum of PANI-PSS multilayer at pH 7 and 1 M NaCl.



Figure E3 UV-vis spectrum of PANI-PSS multilayer at pH 9 and 1 M NaCl.



Figure E4 UV-vis spectrum of PANI-PSS multilayer at pH 12 and 1 M NaCl.



Figure E5 UV-vis spectrum of PANI-CoPSS 3:1 multilayer at pH 4 and 1 M NaCl.



Figure E6 UV-vis spectrum of PANI-CoPSS 3:1 multilayer at pH 7 and 1 M NaCl.



Figure E7 UV-vis spectrum of PANI-CoPSS 3:1 multilayer at pH 9 and 1 M NaCl.



Figure E8 UV-vis spectrum of PANI-CoPSS 3:1 multilayer at pH 12 and 1 M NaCl.





Figure F1 UV-vis spectrum of Ag with various PANI-PSS dilutions.



Figure F2 UV-vis spectrum of Ag with various PANI-CoPSS 3:1 dilutions.



Figure F3 UV-vis spectrum of Ag with various PANI-CoPSS 1:1 dilutions.



Figure F4 UV-vis spectrum of Ag with various PSS dilutions.



Figure F5 UV-vis spectrum of Ag with various CoPSS 3:1 dilutions.



Figure F6 UV-vis spectrum of Ag with various CoPSS 1:1 dilutions.





Figure G1 UV-vis spectrum of Ag-PANI-PSS 0.001 ml with dipping time.



Figure G2 UV-vis spectrum of Ag-PANI-PSS 0.005 ml with dipping time.



Figure G3 UV-vis spectrum of Ag-PANI-PSS 0.01 ml with dipping time.



Figure G4 UV-vis spectrum of Ag-PANI-PSS 0.05 ml with dipping time



Figure G5 UV-vis spectrum of Ag-PANI-PSS 0.1 ml with dipping time.



Figure G6 UV-vis spectrum of Ag-PANI-PSS 0.3 ml with dipping time.



Figure G7 UV-vis spectrum of Ag-PANI-PSS 0.6 ml with dipping time.



Figure G8 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.001 ml with dipping time.



Figure G9 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.005 ml with dipping time.



Figure G10 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.01 ml with dipping time.



Figure G11 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.05 ml with dipping time.



Figure G12 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.1 ml with dipping time.



Figure G13 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.3 ml with dipping time.



Figure G14 UV-vis spectrum of Ag-PANI-CoPSS 3:1 0.6 ml with dipping time.



Figure G15 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.001 ml with dipping time.



Figure G16 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.005 ml with dipping time.



Figure G17 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.01 ml with dipping time.



Figure G18 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.05 ml with dipping time.



Figure G19 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.1 ml with dipping time.

Figure G20 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.3 ml with dipping time.

Figure G21 UV-vis spectrum of Ag-PANI-CoPSS 1:1 0.6 ml with dipping time.

Figure G22 UV-vis spectrum of Ag-PSS 0.001 ml with dipping time.

Ag-PSS 0.005 ml for 5 minutes — Ag-PSS 0.005 ml for 10 minutes Ag-PSS 0.005 ml for 15 minutes — Ag-PSS 0.005 ml for 20 minutes Ag-PSS 0.005 ml for 25 minutes

Figure G23 UV-vis spectrum of Ag-PSS 0.005 ml with dipping time.

Figure G24 UV-vis spectrum of Ag-PSS 0.01 ml with dipping time.

Ag-PSS 0.05 ml for 25 minutes

Figure G25 UV-vis spectrum of Ag-PSS 0.05 ml with dipping time.

Figure G26 UV-vis spectrum of Ag-PSS 0.1 ml with dipping time.

Ag-PSS 0.3 ml for 15 minutes — Ag-PSS 0.3 ml for 20 minutes — Ag-PSS 0.3 ml for 20 minutes — Ag-PSS 0.3 ml for 25 minutes

Figure G27 UV-vis spectrum of Ag-PSS 0.3 ml with dipping time.

Figure G28 UV-vis spectrum of Ag-PSS 0.6 ml with dipping time

Figure G29 UV-vis spectrum of Ag-CoPSS 3:1 0.001 ml with dipping time.

Figure G30 UV-vis spectrum of Ag-CoPSS 3:1 0.005 ml with dipping time.

Figure G31 UV-vis spectrum of Ag-CoPSS 3:1 0.01 ml with dipping time.

Figure G32 UV-vis spectrum of Ag-CoPSS 3:1 0.05 ml with dipping time.

Figure G33 UV-vis spectrum of Ag-CoPSS 3:1 0.1 ml with dipping time.

Figure G34 UV-vis spectrum of Ag-CoPSS 3:1 0.3 ml with dipping time.

Figure G35 UV-vis spectrum of Ag-CoPSS 3:1 0.6 ml with dipping time.

Figure G36 UV-vis spectrum of Ag-CoPSS 1:1 0.001 ml with dipping time.

Figure G37 UV-vis spectrum of Ag-CoPSS 1:1 0.005 ml with dipping time.

Figure G38 UV-vis spectrum of Ag-CoPSS 1:1 0.01 ml with dipping time.

Figure G39 UV-vis spectrum of Ag-CoPSS 1:1 0.05 ml with dipping time.

Figure G40 UV-vis spectrum of Ag-CoPSS 1:1 0.1 ml with dipping time.

Figure G41 UV-vis spectrum of Ag-CoPSS 1:1 0.3 ml with dipping time.

Figure G42 UV-vis spectrum of Ag-CoPSS 1:1 0.6 ml with dipping time.

CURRICULUM VITAE

Name:	Ms. Nuntaporn Jeerawitkhajom	
Date of Birth:	July 7, 1990	
Nationality:	Thai	
University Educati	on:	
2008-2012	Bachelor Degree of Engineering, Faculty of Engineering and industrial technology, Silpakorn University,	
	Nakompathom, Thailand	
Work Experience:		

2011	Position	Student Trainee
	Company name:	Deestone

Proceeding:

 Jeerawitkhajorn, N.; and Dubas, S.(2014, April 22) Development of Polyaniline/Metallic Nanoparticles Hybrid Films for Food Packaging Sensors. <u>Proceedings of the 5th Research Symposium on Petrochemical and Materials</u> <u>Technology and The 20th PPC Symposium on Petroleum, Petrochemicals and Polymers</u>, Bangkok, Thailand.

Presentation:

 Jeerawitkhajorn, N.; and Dubas, S. (2013, December 2) Development of Polyaniline/metallic nanoparticles hybrid films for food packaging sensors, Paper presented at <u>the 2nd Seminar on Future Vision of Green Mobility and Advance</u> <u>Technologies for Realizing the vision</u>, Bangkok, Thailand.