CHAPTER III

DETAILS OF THE CALCULATIONS

3.1 Computational method

All structures, geometries, vibrational frequencies and transition state (TS) were determined using Becke three parameters exchange functional combined the Lee, Yang and Parr correlation functional (B3LYP) and second-order perturbation theory (MP2) with the 6-311++G(d,p) basis set. The calculations were performed with the Gaussian 03 program [25]. The molecular graphics of all molecular structures were generated with the MOLEKEL 4.3 program [26].

In this study, we use the quantum chemical method to predict the molecular mechanism of hydrogen release in the systems of borane amine [12-13], alane amine [14], borane phosphine and alane phosphine and the ability of borane, ammonia, alane and phosphine to serve as a catalyst for hydrogen release reaction. The B3LYP/6–311++G(d,p) optimized structures of studied molecules are shown in Figure 3.1. To ascertain the identity of the relevant transition structure (TS), intrinsic reaction coordinate (IRC) [28] calculations were also done at the same level. Mechanism of hydrogen release reactions of all studied compounds are listed as follows.

3.2 Synthetic and hydrogen release reactions of BH₃NH₃

System without catalyst:

$$BH_3 + NH_3 \rightarrow BH_3NH_3 \tag{3.1}$$

$$BH_3NH_3 \rightarrow tsba \rightarrow BH_2NH_2 + H_2$$
 (3.2)

Systems with catalyst:

$$BH_3NH_3 + BH_3 \rightarrow BH_3NH_3 \cdots BH_3$$
(3.3)

 $BH_3NH_3 \rightarrow tsba' \rightarrow BH_2NH_2 + H_2 + BH_3$ (3.4)

$BH_3NH_3 + NH_3 \rightarrow BH_3NH_3\cdots NH_3$	(3.5)
$BH_3NH_3 \cdots NH_3 \rightarrow tsba'' \rightarrow BH_2NH_2 + H_2 + NH_3$	(3.6)

$$BH_3NH_3 + AlH_3 \rightarrow BH_3NH_3\cdots AlH_3$$
(3.7)

$$BH_3NH_3 \cdots AlH_3 \rightarrow tsba''' \rightarrow BH_2NH_2 + H_2 + AlH_3$$
(3.8)

$$BH_{3}NH_{3} + PH_{3} \rightarrow BH_{3}NH_{3}\cdots PH_{3}$$

$$BH_{3}NH_{2}\cdots PH_{3} \rightarrow BH_{3}NH_{3} + H_{2} + PH_{2}$$

$$(3.9)$$

$$(3.9)$$

3.3 Synthetic and hydrogen release reactions of AlH₃NH₃

System without catalyst:

$AlH_3 + NH_3 \rightarrow AlH_3NH_3$	(3.11)
$AlH_3NH_3 \rightarrow tsala \rightarrow AlH_2NH_2 + H_2$	(3.12)

Systems with catalyst:

$AlH_3NH_3 + BH_3 \rightarrow AlH_3NH_3 \cdots BH_3$	(3.13)
$AlH_3NH_3\cdots BH_3 \rightarrow tsala' \rightarrow AlH_2NH_2 + H_2 + BH_3$	(3.14)
$AlH_3NH_3 + NH_3 \rightarrow AlH_3NH_3 \cdots NH_3$	(3.15)
$AlH_3NH_3NH_3 \rightarrow tsala'' \rightarrow AlH_2NH_2 + H_2 + NH_3$	(3.16)
$AlH_3NH_3 + AlH_3 \rightarrow AlH_3NH_3\cdots AlH_3$	(3.17)
$AlH_3NH_3\cdots AlH_3 \rightarrow tsala''' \rightarrow AlH_2NH_2 + H_2 + AlH_3$	(3.18)
$AlH_3NH_3 + PH_3 \rightarrow AlH_3NH_3 \cdots PH_3$	(3.19)

3.4 Synthetic and hydrogen release reactions of BH₃PH₃

System without catalyst:

$$BH_3 + PH_3 \rightarrow BH_3PH_3 \tag{3.21}$$

$$BH_3PH_3 \rightarrow tsbp \rightarrow BH_2PH_2 + H_2 \tag{3.22}$$

Systems with catalysts:

$$BH_3PH_3 + BH_3 \rightarrow BH_3PH_3\cdots BH_3$$
(3.23)

$$BH_3PH_3\cdots BH_3 \rightarrow tsbp' \rightarrow BH_2PH_2 + H_2 + BH_3$$
(3.24)

$$BH_3PH_3 + NH_3 \rightarrow BH_3PH_3\cdots NH_3$$
(3.25)

 $BH_3PH_3\cdots NH_3 \rightarrow tsbp'' \rightarrow BH_2PH_2 + H_2 + NH_3$ (3.26)

$$BH_{3}PH_{3} + AlH_{3} \rightarrow BH_{3}PH_{3} \cdots AlH_{3}$$

$$BH_{3}PH_{3} \cdots AlH_{3} \rightarrow tsbp''' \rightarrow BH_{2}PH_{2} + H_{2} + AlH_{3}$$

$$(3.27)$$

$$(3.28)$$

$$BH_{3}PH_{3} + PH_{3} \rightarrow BH_{3}PH_{3} \cdots PH_{3}$$

$$BH_{3}PH_{3} \cdots PH_{3} \rightarrow tsbp'''' \rightarrow BH_{2}PH_{2} + H_{2} + PH_{3}$$

$$(3.29)$$

$$(3.20)$$

3.5 Synthetic and hydrogen release reactions of AlH₃PH₃

System without catalyst:

$$AlH_3 + PH_3 \rightarrow AlH_3PH_3$$

$$AlH_3PH_3 \rightarrow tsalp \rightarrow AlH_2PH_2 + H_2$$

$$(3.31)$$

$$(3.32)$$

Systems with catalyst:

$$AlH_3PH_3 + BH_3 \rightarrow AlH_3PH_3 \cdots BH_3$$
(3.33)

 $AlH_3PH_3\cdots BH_3 \rightarrow tsalp' \rightarrow AlH_2PH_2 + H_2 + BH_3$ (3.34)

$AlH_3PH_3 + NH_3 \rightarrow AlH_3PH_3 \cdots NH_3$	(3.35)

 $AlH_3PH_3\cdots NH_3 \rightarrow tsalp'' \rightarrow AlH_2PH_2 + H_2 + NH_3$ (3.36)

 $AlH_{3}PH_{3} + AlH_{3} \rightarrow AlH_{3}PH_{3} \cdots AlH_{3}$ (3.37)

$$AlH_{3}PH_{3}\cdots AlH_{3} \rightarrow tsalp''' \rightarrow AlH_{2}PH_{2} + H_{2} + AlH_{3}$$
(3.38)

$$AlH_{3}PH_{3} + PH_{3} \rightarrow AlH_{3}PH_{3} \cdots PH_{3}$$

$$AlH_{3}PH_{3} \cdots PH_{3} \rightarrow tsalp'''' \rightarrow AlH_{2}PH_{2} + H_{2} + PH_{3}$$

$$(3.39)$$

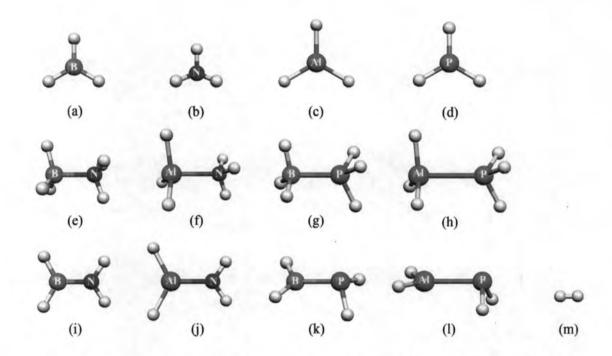


Figure 3.1 The reactant and product models: (a) BH₃, (b) NH₃, (c) AlH₃, (d) PH₃, (e) BH₃NH₃, (f) AlH₃NH₃, (g) BH₃PH₃, (h) AlH₃PH₃, (i) BH₂NH₂, (j) AlH₂NH₂, (k) BH₃PH₃, (l) AlH₃PH₃, (m) H₂ optimized with B3LYP/6-311++G(d,p).

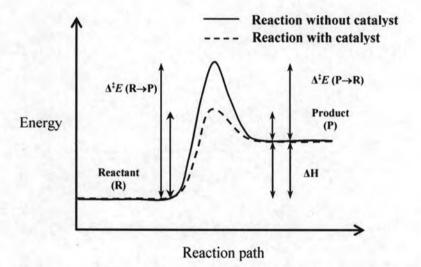
3.6 Transition state and thermodynamic properties calculations

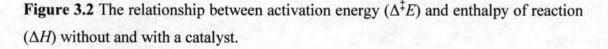
In this work, we have predicted the molecular mechanism of hydrogen release from borane amine, alane amine, borane phosphine and alane phosphine in systems without and with the borane, ammonia, alane and phosphine. Transition-state structures optimized at the B3LYP/6–311++G(d,p) and MP2/6–311++G(d,p) levels of theory. Their thermodynamic properties were derived from the vibrational frequency calculations at their corresponding levels of theory. Frequency calculations were carried out to confirm the single imaginary mode for all transition-state structures.

3.6.1 Activation energies

The activation energy $(\Delta^{\ddagger} E)$ is the energy difference between the reactants and the highest transition state barrier are shown in expression (3.41). In the catalysis reaction, the energy required to enter transition state decreases, thereby decreasing the energy required to initiate the reaction as shown in Figure 3.2.

$$\Delta^{\ddagger} E = E_{(TS)} - E_{(reactant)} \tag{3.41}$$





3.6.2 Rate constant and equilibrium constant

The rate constant k(T), in term of partition function (Q) and equilibrium constant, K derived from transition state theory were calculated from activation free energy, $\Delta^{\ddagger}G^{\circ}$ by use of the following equation:

$$k(T) = \frac{k_B T}{hc^0} e^{-\Delta^2 G^0 / RT}$$
(3.42)

$$k = \kappa \left[\frac{k_B T}{h}\right] \left[\frac{Q_{TS}}{Q_{Complex}}\right] e^{-\Delta^{\ddagger} E / RT}$$
(3.43)

$$k = \kappa A e^{-\Delta^{2} E / RT} \tag{3.44}$$

$$K = e^{-\Delta^2 G/RT} \tag{3.45}$$

where $\kappa = 1 + (1/24)(hv_i c/k_B T)^2$, $A = (k_B T/h)/(Q_{TS}/Q_{Complex})$, k_B is Boltzman constant, T is absolute temperature, Q_{TS} and $Q_{Complex}$ are the partition functions of transition state and complex, respectively, h is Plank constant, c^0 is concentration factor, c is speed of light and v_i is imaginary frequency of transition state.

3.6.3 Gibbs free energy and enthalpy

The energy difference between reactants and products is the enthalpy (ΔH) or the Gibbs free energy (ΔG) of the reaction as shown in equations (3.46) and (3.47). If the energy of the reactants is higher than the energy of the products, the reaction is exothermic. If the energy of the reactants is lower than the energy of the products, the reaction is endothermic.

$$\Delta H = H_{product} - H_{reactant}$$
(3.46)
$$\Delta G = G_{product} - G_{reactant}$$
(3.47)