

REFERENCES

1. Altieri, Paul A. Biodegradable shaped products and the method of preparation thereof. *U.S. Patent. Appl. No. 5,153,037*, 1992.
2. Lacourse, Norman L., Altieri, Paul A. Biodegradable packaging material and the method of preparation thereof. *U.S. Patent. Appl. No. 4,863,655*, 1989.
3. Randal L. Shogren., John W. Lawton., Karl F. Tiefenbacher., Liang Chen. Starch-Poly (vinyl alcohol) Foamed Articles Prepared by baking Process. *J. of Applied Polymer Science*, 68, (1998): 2129-2140.
4. Kakinoki, Hideo., Sato, Hisaya., Foam-expanded material. *U.S. Patent. Appl. No. 5,766,749*, 1998.
5. Jeffs, Hyrum J. Biodegradable packaging foam and method of preparation. *U.S. Patent Appl. No. 5,252,271*, 1993.
6. C. L. Swanson., R.L. Shogren., G. F. Fanta., S. H. Imam. Starch-Plastic Materials-Preparation, Physical Properties, and Biodegradability (A Review of Recent USDA Research). *J. of environmental Polymer Degradation*. 1(2), 1993, pp 155-166.
7. Catia Bastioli. Starch polymer composites. *Degradable Polymers*. London: Chapman & Hall, 1995, pp 112-137.
8. Gerald O. Aspinall. *Polysaccharides*. Oxford: Pergamon Press Ltd., Headington Hill Hall, 1970, pp 54-68.
9. Randal L. Shogren., George F. Fanta., William M. Doane., Peoria, IL. Development of Starch Based Plastics-A Reexamination of Selected Polymer Systems in Historical Perspective. *Starch/Starke*. 45 (1993): 276-280.

10. Gazeley KF., Gorton ADT., Pendle TD. Latex concentrates: properties and composition. In: Roberts AD, editor. *Natural rubber science and technology*. Oxford: Oxford University Press, 1988.
11. Blackley DC. Latices. In: Mark HF, Bikales NM, Overberger CG, editors. *Encyclopedia of polymer science and engineering*. 2nd ed. Canada: John Wiley & Sons, 1987; 8: 647-677.
12. Ottewill RH. Effect of nonionic surfactants on the stability of dispersions. In: Schick MJ, editor. *Nonionic surfactant*. 2nd ed. New York: Marcel Dekker, 1967: 627-679.
13. Ottewill RH. Colloidal properties of latex particles. In: candau F, Ottewil RH, editors. *An introduction to polymer collids*. Dordrecht: Kluwer Academic Publishers, 1990: 129-157.
14. Barlow FW. *Rubber compounding: principle, materials, and techniques*, New York, Marcel Dekker; 1988.
15. Ultracki, L A. *Polymer alloys and blends*. New York: Hanser, 1990.
16. Daniel Klempner., Kurt C. Frisch. *Hanbook of Polymeric Foams and Foam Technology*. New York: Hanser, 1991.
17. Baird, Ronald J. *Industrial plastics*. The goodheart-willcox company. INC. 1986.
18. Tomka, Ivan. Foamed starch polymer. *U.S. Patent Appl. No. 5,705,536*, 1998.
19. Boehmer, Edward W., Hanlon, Daniel L. Biodegradable expanded foam material. *U.S. Patent Appl. No. 5,272,181*, 1993.
20. Akamatu, Yoshimi., Tomori, Masashiro. Process for preparing biodegradable resin foam. *U.S. Patent Appl. No. 5,308,879*, 1994.

21. Kustner, Franz. Process for producing a foamed product or foam made of unmodified starch. *U.S. Patent Appl. No. 5,476,621*, 1995.
22. Franke, Hans., Bittner, Donald R. Resilient biodegradable packaging materials. *U.S. Patent Appl. No. 5,766,529*, 1998.
23. Neumann, Paul E, Seib, Paul A. Starch-based, biodegradable packaging filler and method of preparing same. *U.S. Patent Appl. No. 5,208,267*, 1993.
24. Knight, Adrian T. Starch derived shaped articles. *U.S. Patent Appl. No. 5,314,754*, 1994.
25. ASTM D 1076-88. Standard specification for rubber-concentrated, ammonia preserved, creamed and centrifuged natural latex. In: *Annual book of ASTM standards*. 1997; 9.01: section 9.
- 26 M. A. Rodriguez-perez, S. Rodriguez-llorente, and J. A. De saja. Dynamic Mechanical Properties of Polyolefin Foams Studied by DMA Techniques. *Polymer Engineering and Science*, 37 (6), (1997): 959-965.
27. Y. Vodovotz and P. Chinachoti. Thermal Transitions in Gelatinized Wheat Starch at Different Moisture Contents by Dynamic Mechanical Analysis. *J. of Food Science*, 61 (5), (1996): 932-941.
28. Olagoke Olabisi., lloyd M. Robeson., Montgomery T. Shaw. *Polymer-Polymer Miscibility*. New York: Academic Press, 1979.
29. Malvern Instruments Ltd. Mastersizer reference manual: 1-10.
30. Sanchez J, Myers TN. Peroxide initiators. *Polymeric material encyclopedia* Vol.7. CRC Press; 1996.

APPENDICES

APPENDIX A

Particle Size Measurement

Particle size of starch and the concentrated NR latex was measured by using Mastersizer S. The diagram of Mastersizer S is schematically presented in Figure A. The instrument is composed of an optical unit which includes the basic particle size sensor interfaced with a system to aid computer to manage the measurement and performs result analysis and presentation.

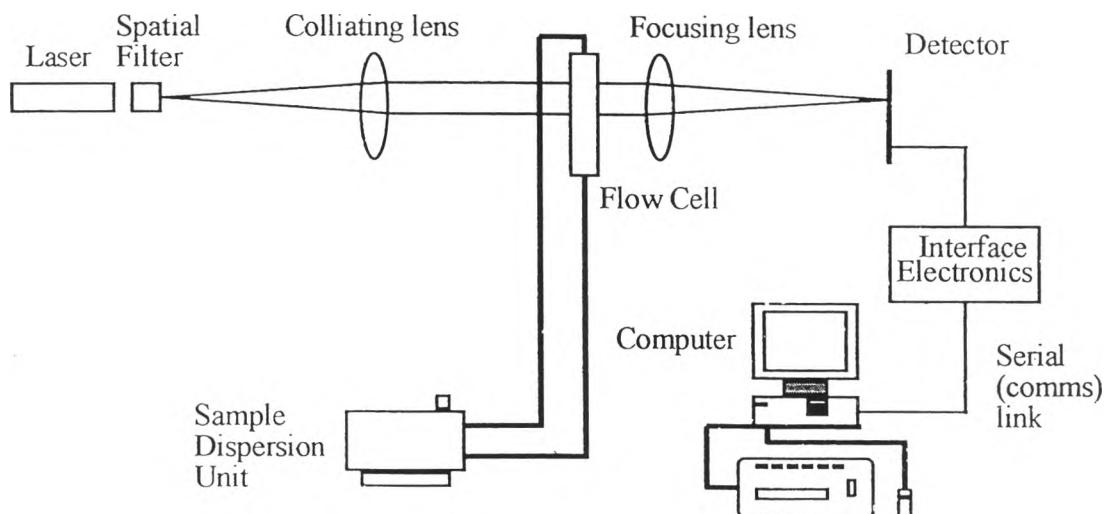


Figure A: Schematic diagram of Mastersizer S [29]

The principle of Mastersizer is based on laser light scattering both diffusion and diffraction. The refractive index of the material must be different from the medium in which it is supported. The Mastersizer S employs two forms of optical configuration to provide its unique specification. The first is the well-known optical method, called “Conventional Fourier Optics”. The second is called “Reverse Fourier Optics”, used in order to allow the measurement size range to be extended down to 0.05 μm .

APPENDIX B

Table B.1: Half-life temperature for commercial peroxides [30]

Chemical Name	Solvent ^d	E_a^b kJ/ mol	Half-life Temperature, °C			
			ln(A)	10-hr	1-hr	1-min
Class: Diacyl Peroxides						
Di(2,4-dichlorobenzoyl) peroxide	Benzene	117.2	32.3	54	73	111
Duisomannoyl peroxide	TCE	129.7	35.9	61	78	114
Dibenzoyl peroxide	Benzene	128.9	35.9	73	92	131
Di(decanoyl) peroxide	TCE	128.4	34.9	65	83	120
Dilauroyl peroxide	TCE	130.1	35.6	62	81	117
Succinic acid peroxide	Acetone	97.5	23.7	66	90	143
Class: Dialkyl Peroxides						
1,3-Di(<i>t</i> -butylperoxy)-2,5-dimethylhexane	Dodecane	154.0	36.4	119	139	181
2,5-Di(<i>t</i> -butylperoxy)-2,5-dimethylhexane	Dodecane	159.4	36.6	131	152	194
2,5-Di(<i>t</i> -butylperoxy)-2,5-dimethylhexane	Dodecane	155.6	36.7	120	140	181
Di- <i>a</i> -cumyl peroxide	Decane	154.0	36.6	117	137	178
Di- <i>tert</i> -amyl peroxide	Dodecane	159.4	37.5	123	143	182
Di- <i>tert</i> -butyl peroxide	Decane	164.8	38.4	129	149	189
<i>tert</i> -Butyl <i>a</i> -cumyl peroxide	Dodecane	158.2	37.0	124	143	185
Class: Dperoxyketals						
1,1-Di(<i>tert</i> -amylperoxy)cyclohexane	Dodecane	144.8	36.6	93	112	150
1,1-Di(<i>tert</i> -butylperoxy)-2,5-dimethylcyclohexane	Dodecane	148.5	37.5	96	115	153
1,1-Di(<i>tert</i> -butylperoxy)cyclohexane	Dodecane	144.8	36.2	97	116	155
2,2-Di(<i>tert</i> -amylperoxy)propane	Dodecane	144.3	34.8	108	128	170
2,2-Di(<i>tert</i> -butylperoxy)butane	Dodecane	143.5	34.5	107	127	169
Ethyl 3,3-di(<i>tert</i> -amylperoxy)butyrate	Dodecane	148.5	35.6	112	132	173
Ethyl 3,3-di(<i>tert</i> -butylperoxy)butyrate	Dodecane	151.5	36.2	114	134	175
<i>n</i> -Butyl 4,4-di(<i>tert</i> -butylperoxy)valerate	Dodecane	147.7	35.6	109	129	170
Class: Peroxydicarbonates						
Di-2-ethylhexyl peroxydicarbonate	TCE	127.6	36.7	49	66	99
Di-2-phenoxyethyl peroxydicarbonate	TCE	128.0	36.8	50	67	101
Di-2-ethylhexyl peroxydicarbonate	TCE	123.8	35.3	50	67	102
Dihexadecyl peroxydicarbonate	TCE	125.1	35.8	50	67	101
Diisopropyl peroxydicarbonate	TCE	123.8	35.2	50	67	102
Di-n-propyl peroxydicarbonate	TCE	128.4	37.1	50	66	99
Di- <i>tert</i> -butyl peroxydicarbonate	TCE	116.7	32.5	51	69	102
Class: Peroxesters						
2,5-Di(2-ethylhexanoylperoxy)-2,5-dimethylhexane	Decane	131.4	32.8	73	91	129
2,5-Di(<i>tert</i> -butylperoxy)-2,5-dimethylhexane	Benzene	152.3	38.2	100	118	156
3-Hydroxy-1,1-dimethylbutyl 2-ethylperoxyhexanoate	AMS	118.0	31.2	65	82	125
3-Hydroxy-1,1-dimethylbutyl peroxyneodecanoate	TCE	111.3	32.3	37	52	91
3-Hydroxy-1,1-dimethylbutyl peroxyneohexanoate	AMS	115.1	31.2	41	58	92
<i>tert</i> -Cumyl peroxyneodecanoate	TCE	111.3	32.0	38	56	93
<i>tert</i> -Cumyl peroxyneohexanoate	TCE	115.9	33.4	43	60	96
Di- <i>tert</i> -butyl peroxyphthalate	Benzene	125.2	39.6	104	122	159
OO- <i>tert</i> -amyl O-(2-ethylhexyl) monoperoxycarbonate	Dodecane	150.6	37.8	99	117	155
OO- <i>tert</i> -butyl O-(2-ethylhexyl) monoperoxycarbonate	Dodecane	151.5	31.7	100	121	166
OO- <i>tert</i> -butyl O-isopropyl monoperoxycarbonate	Benzene	142.3	35.2	99	118	159
<i>tert</i> -Amyl 2-ethylperoxyhexanoate	TCE	121.8	38.4	73	96	125
<i>tert</i> -Amyl peroxyacetate	Dodecane	139.3	34.1	100	120	162
<i>tert</i> -Amyl peroxybenzoate	Dodecane	138.5	33.9	100	120	162
<i>tert</i> -Amyl peroxyneodecanoate	TCE	120.5	32.8	46	64	99
<i>tert</i> -Amyl peroxyisopivalate	TCE	117.6	32.2	55	74	112
<i>tert</i> -Butyl 2-ethylperoxyhexanoate	Dodecane	122.7	38.1	77	95	130
<i>tert</i> -Butyl 3,5,5-trimethylperoxyhexanoate	Benzene	138.9	33.7	101	122	164
<i>tert</i> -Butyl peroxymaleate	Acetone	113.0	26.9	87	111	161
<i>tert</i> -Butyl peroxyacetate	Decane	137.7	33.3	102	123	166
<i>tert</i> -Butyl peroxybenzoate	Dodecane	132.2	31.4	104	125	171
<i>tert</i> -Butyl peroxyisobutyrate	Decane	123.8	31.1	82	102	146
<i>tert</i> -Butyl peroxyneodecanoate	TCE	119.2	33.7	48	66	102
<i>tert</i> -Butyl peroxyneohexanoate	TCE	113.4	31.0	53	72	112
<i>tert</i> -Butyl peroxyisopivalate	TCE	117.6	32.0	58	76	116

^bTCE = trichloroethylene, AMS = α -methylstyrene.

^cTo change from kilojoules to kilocalories, divide by 4.184.

VITA

Ms Piyawan Surunchanajirasakul was born on October 25, 1971 in Mahasarakham. She graduated with a Bachelor Degree of Science (Chemistry) from Burapha University in 1994. She has joined the National Metal and Material Technology Center (MTEC) since 1996. In 1998, she was accepted as a graduate student in the Program of Petrochemistry and Polymer Science, Faculty of Science, Chulalongkorn University. She received a Master's Degree of Science in Polymer Science, in October 2000.

