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ศูนย์วิทยทรัพยากร
จุฬาลงกรณ์มหาวิทยาลัย

APPENDIX

1. Calculation of Biogas at STP

In this study the total pressure of biogas collected on the top of the UASB reactor operating under 15 cm. water pressure was 780.6 mm Hg. The collecting of biogas was made over water of 38°C and the temperature of biogas was 40°C in average.

The calculation was based on the Gas Law as

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

or
$$V_2 = V_1 \left(\frac{P_1}{P_2} \right) \left(\frac{T_2}{T_1} \right)$$

where V_1 = volume of biogas collected
 V_2 = final volume of gas at STP
 P_1 = pressure of biogas collected
 P_2 = standard pressure = 760 mm Hg
 T_1 = temperature of biogas collected (°K)
 T_2 = standard temperature = 273°K

The biogas collected was estimated to compose of CH₄ & CO₂ (ignoring traces of N₂, H₂S, and NH₃) and water vapor. The partial pressure of water vapor at 38°C is 49.7 mm Hg (see Table I).

Table I
Vapor pressure of water of various temperatures

Temperature (°C)	30	35	38	40	50	60
Vapor Pressure (mm Hg)	31.8	42.2	49.7	55.3	92.5	149.4

Source : *Handbook of Chemistry and Physics, 40th ed., CRC Publishing Co.*

Hence the partial pressure of the biogas is the total pressure less 49.7 mm Hg.

$$\begin{aligned}
 \text{Pressure of dry biogas} &= \text{total pressure} - \text{vapor pressure of water} \\
 &= 780.6 \text{ mm Hg} - 49.7 \text{ mm Hg} \\
 &= 730.9 \text{ (or } 731 \text{ mm Hg)}
 \end{aligned}$$

Example

Exactly 300 l of biogas is collected per day from the UASB reactor. Compute the standard volume of the dry biogas.

$$\begin{aligned}
 \text{Volume of biogas at STP (V}_2\text{)} &= 300 \text{ l} \times \frac{731 \text{ mm Hg}}{760 \text{ mm Hg}} \times \frac{273^\circ\text{K}}{273+40^\circ\text{K}} \\
 &= 251.7 \text{ l/d}
 \end{aligned}$$

2. Calculation of CO₂ Solubility

The solubility of CO₂ is based on Henry's Law as

$$X = \frac{P}{K}$$

where X = mole fraction of CO₂
P = partial pressure of CO₂ (mm Hg)
K = Henry's Constant for CO₂

In this study, the UASB reactor was always found in a normal performance. Hence the given CO_2 and CH_4 containing in the biogas collected were 25% and 75% by volume, respectively.

With the above pressure of dry biogas computed (in item 1), it equals to 731 mm Hg; the partial pressure of CO_2 and CH_4 are therefore asd follow :

$$\text{Partial pressure of } \text{CO}_2 = 731 \times \frac{25}{100} = 182.75 \text{ mm Hg}$$

$$\text{Partial pressure of } \text{CH}_4 = 731 \times \frac{75}{100} = 548.25 \text{ mm Hg}$$

And Henry's Constant for carbon dioxide at various temperature are listed in Table II.

Table II

Henry's Constant for carbon dioxide at various temperatures

Temperature ($^{\circ}\text{C}$)	30	38	40	50	60
$K \times 10^{-7}$	0.139	0.168	0.173	0.217	0.258

Source : *Handbook of Chemistry and Physics, 40th ed., CRC Publishing Co.*

We can find CO_2 solubility in water 1 liter from Henry's Law as

$$X_{\text{CO}_2} = \frac{P_{\text{CO}_2}}{K_{\text{CO}_2}(38^{\circ}\text{C})} = \frac{\text{moles of } \text{CO}_2}{\text{moles of } \text{CO}_2 + \text{moles of } \text{H}_2\text{O} + \text{moles of } \text{CH}_4}$$

$$\text{where moles of } \text{H}_2\text{O} = \frac{1000\text{g}}{18.02\text{g}} = 55.49$$

moles of CO_2 , moles of CH_4 \lll moles of H_2O \dots neglected

$$\begin{aligned} \text{Hence moles of } \text{CO}_2 &= \frac{182.75 \text{ mm Hg} \times 55.49}{0.168 \times 10^7} \\ &= 6.04 \times 10^{-3} \end{aligned}$$

$$\text{Molecular Weight of } \text{CO}_2 \text{ 1M} = 44 \text{ g}$$

$$\text{Volume of } \text{CO}_2 \text{ 1M} = 22.4 \text{ l}$$

$$\dots \text{ The } \text{CO}_2 \text{ solubility} = 6.04 \times 10^{-3} \times 44 = 0.26 \text{ g/l}$$

$$\text{or} = 0.04 \times 10^{-3} \times 22.4 = 0.14 \text{ l/l}$$

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VITA

The author, Ms. Arunee Wajanavijai, was born on October 21, 1960 here in Bangkok, Thailand. She received a bachelor degree in general science from Chulalongkorn University in the education year 1982. She then worked as a staff of the Boon Rawd Brewery Co., Ltd ever since the year 1983. After 4 years working, she was sent to further her postgraduate study in Environmental Science at Graduate School, Chulalongkorn University. At present, she is now in charge of chief of wastewater analysis laboratory which is a unit in the Boon Rawd Brewery Wastewater Treatment Section.

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