

REFFERECES

- [1] Jiang, Q.; Qu, M.; Zhou, G.; and Zhang, B. A study of activated carbon nanotubes as electrochemical super capacitor electrode materials. Materials Letters 57 (2002): 988-991.
- [2] Chen, Q.; Xu, R.; and Yu, D. Multiwalled carbon nanotube/polybenzoxazine nanocomposites: Preparation, characterization and properties. Polymer 47 (2006): 7711-7719.
- [3] Catalytic Synthesis of Carbon Nanotubes [online]. (n.d.). Available from: <http://cobweb.ecn.purdue.edu/~catalyst/Carbon%20Nanotubes/Catalytic%20Nanotubes/Catalytic%20Synthesis%20of%20Carbon%20Nanotubes.htm> [2008, March 24].
- [4] Gabriel, G.B. Design considerations and characterization of origami electrochemical capacitors. Bachelor's Thesis, Department of Mechanical Engineering, Massachusetts Institute of Technology, 2004.
- [5] Jian, Y.; Xiao, L.; Hui, C.; Wei, D.; and Fwu, S. Electrochemical oxidation of multi-walled carbon nanotubes and its application to electrochemical double layer capacitors. Electrochemistry Communications 7 (2005): 249-255.
- [6] Yan, S.; and Lian, G. Formation and characterization of multi-wall carbon nanotubes/Co₃O₄ nanocomposites for supercapacitors. Materials Chemistry and Physics 103 (2007): 206-210.
- [7] Mukta, S.; and Girish, A. Preparation and characterization of composite electrodes of coconut-shell-based activated carbon and hydrous ruthenium oxide for capacitor. Journal of Power Sources 141 (2005): 198-203.
- [8] Frackowiak, E.; and Beguin, F. Carbon materials for the electrochemical storage of energy in capacitors. Carbon 39 (2001): 937-950.
- [9] Gurunathan, K.; Murugan, A.; and Mulik, U. Electrochemical synthesis conducting polymeric materials for applications towards technology in electronics, optoelectronics and energy storage devices. Mater chem. Phys 61 (1999): 173-191.

- [10] Kim, I.H.; and Kim, K.B. Ruthenium oxide thin film electrodes for supercapacitor. Electrochemical Solid-State Lett 4 (2001): 62-64.
- [11] Wang, Q.; Zhang, B.; Yu, Z.; and Qu, M. Manganese oxide/MWNTs composite electrodes for supercapacitors. Solid State Ionics 176 (2005): 1169-1174.
- [12] Leela, A.; and Ramaprabhu, S. Nanocrystalline metal oxide dispersed multiwalled carbon nanotubes as supercapacitor electrodes. Journal of Physical Chemistry 111 (2007): 7727-7734.
- [13] Sawyer, D.T.; Sobkowiak, A.; and Robert, J.L. Electrochemistry for Chemists. New York: Wiley Interscience, 1995.
- [14] Settle, F.A. Handbook of Instrumental Techniques for Analytical Chemistry. New Jersey: Prentice Hall, 1997.
- [15] Thomas, F.G.; and Henze, G. Introduction to Voltammetric Analysis. Collingwood, Australia: CSIRO Publishing, 2001.
- [16] Kim, J.H.; Nam, K.W.; and Kim, K.B. Fabrication and electrochemical properties of carbon nanotube film electrodes. 44 Carbon (2006): 1963-1968.
- [17] Camelia, G.; Joop, S.; and Martin, L. Porous indium oxide thin film deposited by electrostatic spray deposition technique. Ceramics International 34 (2006): 95-100.
- [18] Shu, D.; Chung, K.Y.; and Kim, K.B. Electrochemical investigations on electrostatic spray deposited LiMn_2O_4 films. Journal of Power Sources 114 (2003): 253-263.
- [19] Deng, G.; Xiao, X.; Chen, J.; He, D.; and Kuang, Y. A new method to prepare $\text{RuO}_2 \cdot x\text{H}_2\text{O}$ /carbon nanotube composite for electrochemical capacitors. Carbon 43 (2005): 1557-1583.
- [20] Kyung, N.W.; and Won, S. X-ray absorption spectroscopy studies of nickel oxide thin film electrodes for supercapacitors. Electrochimica Acta 47 (2002): 3201-3209.
- [21] Chang, Z.Y. ; and Gao, B. Electrochemical capacitance of $\text{NiO}/\text{Ru}_{0.35}\text{V}_{0.65}\text{O}_2$ asymmetric electrochemical capacitor. Journal of Power Sources 158 (2006): 1538-1543.

- [22] Christian, G.D.; and O'Reilly, J.E. Instrument Analysis Vol.2. (n.d.). Allyn and Bacon, 1986.
- [23] Hughes, M.; Chen, G.Z.; Fray, D.J.; and Windle, A.H. Electrochemical capacitance of a nanoporous composite of carbon nanotubes and polypyrrole. Chem Mater 14 (2000): 1610-1613.
- [24] Kotz, R.; and Carlen, M. Principle and application of electrochemical capacitors. Electrochim Acta 45 (2000): 2483-2488.
- [25] Kim, Y.J.; Tantrakarn, K.; and Abe, Y. Correlation between the capacitor performance and pore structure. Tanso 221 (2006): 31-39.
- [26] Gryglewicz, G.; and Machnikowski, J. Effect of pore size distribution of coal-based activated carbons on double layer capacitance. Electrochim Acta 50 (2005): 1197-1206.
- [27] X-Ray powder Diffractometer [online]. (n.d.). Available from: http://www.tint.or.th/adv/phys_oap/XRD.pdf [2008, March 24].
- [28] Kim, U.J.; Furtado, C.A.; Lui, X.; Chen, G.; and Eklund, P.C. Raman and IR spectroscopy of chemically processed single-walled nanotubes. Journal of Am Chem Soc 127 (2005): 15437-15445.
- [29] Kuznetsova, A.; Douglas, D.; Namneko, V.; Liu, J.; and Smalley, R.E. Enhancement of adsorption of single-walled carbon nanotubes : opening the entry ports. Chem Phys Letter 321 (2000): 292-296.
- [30] Saito, R.; Dreselhuas, G.; and Dreselhuas, M.S. Physical properties of carbon nanotubes. Imperial College Press 34 (1999): 814-816.
- [31] Kuzmang, H.; Simon, F.; and Holzweber, M. Functionalization of carbon nanotubes. Synthetic Metals 141 (2004): 113-122.
- [32] Jian, L.; Wei, D.; Li, P.; Xiao, L.; and Qing, H. Preparation of multi-walled nanotubes grafted with synthetic poly(L-lysine) through surface-initiated ring-opening polymerization. Polymer 48 (2007): 4352-4360.
- [33] Leela, A.; Ramaprabhu, S., Hydrogen storage properties of nanocrystalline Pt dispersed multi-walled carbon nanotubes., Journal of Hydrogen Energy., 2007, 32, 3998-4004.
- [34] Elzbieta, F. Carbon materials for the electrochemical storage of energy in capacitors. Carbon 39 (2001): 937-950.

- [35] Chen, G.; Hai, T.; Feng, L.; Min, L.; and Hui, M. Single-walled carbon nanotubes modified by electrochemical treatment for application in electrochemical capacitors. Journal of Power Sources 160 (2006): 758-761.
- [36] Chen, J.H.; Li, W.Z.; and Wang, D.Z. Electrochemical characterization of carbon nanotubes as electrode in electrochemical double-layer capacitors. Carbon 40 (2002): 1193-1197.
- [37] Eiki, I.; Sylwia, M.; and Takashi, N. Nanoporous carbons from cypress II. Application to electric double layer capacitors. New Carbon Materials 22 (2007): 321-326.
- [38] Frackowiak, E.; Mettenier, K.; and Beguin, F. Supercapacitor electrodes from multiwalled carbon nanotubes. Appl Phys Lett 77 (2000): 2421-2423.
- [39] Chen, L.; Xue, H.; Shen, W.; and Tao, F. Fabrication and electrochemical properties of carbon nanotube array electrode for supercapacitor. Electrochimica Acta 49 (2004): 4157-4161.
- [40] Emmenegger, C.; Mauron, P.; Sudan, P.; and Wenger, P. Investigation of electrochemical double-layer capacitor electrodes based on carbon nanotubes and activated carbon materials. Journal of Power Sources 124 (2003): 321-329.
- [41] Liu, Z.; Hong, L.; and Guo, B. Physicochemical and electrochemical characterization of anatase titanium dioxide nanoparticles. Journal of Power Sources 143 (2005): 231-235.
- [42] Ramani, M.; Haran, B.S.; and White, R.E. Studies on activated carbon capacitor materials loaded with different amounts of ruthenium oxide. Journal of Power Sources 93 (2001): 209-214.
- [43] Conway, B.E.; Birss, V.; and Wojtowicz, J. The role and utilization of pseudocapacitance for energy storage by supercapacitors. Journal of Power Sources 66 (1997): 1-14.
- [44] Yan, S.; and Lian, G. Formation and characterization of multi-wall carbon nanotubes/ Co_3O_4 nanocomposites for supercapacitors. Materials Chemistry and Physics 103 (2007): 206-210.

- [45] Deng, G.; Xiao, X.; Chen, J.; He, D.; and Kuang, Y. A new method to prepare RuO₂ xH₂O/carbon nanotube composite for electrochemical capacitors. Carbon 43 (2005): 1557-1583.
- [46] Vasile, V.N. On the performance of supercapacitors with electrodes based on carbon nanotubes and activated material. Physica E 40 (2008): 2596-2605.
- [47] Chuang, P.; Shengwen, Z.; and Daniel, J. Carbon nanotube and conducting polymer composites for supercapacitors. Progress in Natural Science 18 (2008): 777-788.
- [48] Jayalakshmi, M.; and Venugopal, N. Nano SnO₂-Al₂O₃ – carbon composite oxides as new and novel electrodes for supercapacitor applications. Journal of Power Sources 158 (2006): 1538-1543.
- [49] Kai, D.; and Lixi, S. Synthesis of silver nanoparticles on functional multi-walled carbon nanotubes. Material Science and Engineering A 465 (2007): 283-286.
- [50] Nae, L. Nanocrystalline oxide supercapacitors. Materials Chemistry and Physics 75 (2002): 6-11.
- [51] Jae, K.H.; Byung, C.W.; Young, L.H.; and Kwang, K.B. Synthesis and electrochemical characterization of vanadium oxide on carbon nanotube film substrate for pseudocapacitor applications. Journal of The Electrochemical Society 153 (2006): 989-996.
- [52] Jae, K.H.; Byung, C.W.; Young, L.H.; and Kwang, K.B. Synthesis and characterization of electrochemically prepared ruthenium oxide on carbon nanotube film substrate for pseudocapacitor applications. Journal of The Electrochemical Society 152 (2005): 2170-2178.
- [53] Wendong, W.; Philippe, S.; Philippe, K.; and Joaquim, L. Photocatalytic degradation of phenol on MWNT and titania composite catalysts prepared by a modified sol-gel method. Applied Catalysis 56 (2005): 305-312.
- [54] Nui, C.; Sichel, E.; Hoch, R.; and Moy, D. High power electrochemical capacitors based on carbon nanotube electrodes. Appl Phys Lett 70 (1997): 1480-1484.

- [55] Xie, Y.; Huang, C.; and Zhou, L. Supercapacitor application of nickel oxide-titania nanocomposites. Composites Science and Technology (2009).
- [56] Stephen, A.; Ronald, E.; and Sergej, Y. Structure, metal-insulator transitions and magnetic properties of FeO at high pressures. American Mineralogist 88 (2003): 257-261.

APPENDICES

APPENDIX A

Results from cyclic voltammetric technique

Table A Specific capacitance from cyclic voltammetry

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	1	2	3	
- Refluxed MWNTs with 6 hours	57.42	58.65	57.84	57.97
- Refluxed MWNTs with 12 hours	70.34	69.42	69.63	69.76
- Refluxed MWNTs with 18 hours	81.67	83.34	82.65	82.55
- 2 % w/w TiO ₂ /MWNTs	144.31	145.05	144.01	144.45
- 4 % w/w TiO ₂ /MWNTs	155.34	156.06	155.86	155.75
- 6 % w/w TiO ₂ /MWNTs	172.08	174.61	173.90	173.53
- 8 % w/w TiO ₂ /MWNTs	176.55	177.00	176.13	176.56
- 10 % w/w TiO ₂ /MWNTs	178.65	177.74	178.07	178.15
- 2 % w/w FeO/MWNTs	101.56	100.49	101.61	101.22
- 4 % w/w FeO/MWNTs	103.53	103.67	102.98	103.39
- 6 % w/w FeO/MWNTs	104.46	104.92	104.66	104.68
- 8 % w/w FeO/MWNTs	104.92	105.02	105.23	105.05
- 10 % w/w FeO/MWNTs	106.01	105.65	105.61	105.75
- 2 % w/w NiO/MWNTs	146.56	147.23	146.90	146.89
- 4 % w/w NiO/MWNTs	151.65	152.23	152.67	152.18
- 6 % w/w NiO/MWNTs	158.96	157.76	157.48	158.06
- 8 % w/w NiO/MWNTs	168.93	169.11	168.73	168.92
- 10 % w/w NiO/MWNTs	172.56	171.31	171.86	171.91

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	1	2	3	
- 2 : 8 % w/w FeO-TiO ₂ /MWNTs	123.88	124.54	124.03	124.15
- 4 : 6 % w/w FeO-TiO ₂ /MWNTs	123.44	124.31	123.46	123.78
- 6 : 4 % w/w FeO-TiO ₂ /MWNTs	122.56	122.43	123.77	122.92
- 8 : 2 % w/w FeO-TiO ₂ /MWNTs	121.67	121.52	122.06	121.75
- 2 : 8 % w/w NiO-TiO ₂ /MWNTs	179.08	180.88	179.21	179.72
- 4 : 6 % w/w NiO-TiO ₂ /MWNTs	180.67	180.84	181.77	181.09
- 6 : 4 % w/w NiO-TiO ₂ /MWNTs	179.56	180.49	180.66	180.23
- 8 : 2 % w/w NiO-TiO ₂ /MWNTs	180.42	179.34	180.43	180.06

APPENDIX B

Results from electrochemical impedance spectroscopic technique

Table B Specific capacitance from impedance spectroscopy

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	1	2	3	
- Refluxed MWNTs with 6 hours	55.63	55.34	55.78	55.58
- Refluxed MWNTs with 12 hours	67.90	66.43	66.55	66.96
- Refluxed MWNTs with 18 hours	79.42	80.47	80.06	79.98
- 2 % w/w TiO ₂ /MWNTs	139.33	140.51	140.99	140.27
- 4 % w/w TiO ₂ /MWNTs	150.37	150.58	149.57	150.17
- 6 % w/w TiO ₂ /MWNTs	170.03	171.07	171.60	170.90
- 8 % w/w TiO ₂ /MWNTs	176.48	175.08	175.32	175.62
- 10 % w/w TiO ₂ /MWNTs	176.89	176.83	177.68	177.13
- 2 % w/w FeO/MWNTs	98.56	99.09	99.80	99.15
- 4 % w/w FeO/MWNTs	102.67	103.56	102.21	102.81
- 6 % w/w FeO/MWNTs	104.24	103.39	103.18	103.60
- 8 % w/w FeO/MWNTs	104.46	104.55	105.00	104.67
- 10 % w/w FeO/MWNTs	105.44	104.63	105.03	105.03
- 2 % w/w NiO/MWNTs	144.62	145.34	144.25	144.73
- 4 % w/w NiO/MWNTs	150.04	149.66	149.68	149.97
- 6 % w/w NiO/MWNTs	154.51	155.47	155.62	155.20
- 8 % w/w NiO/MWNTs	167.30	166.93	166.31	166.84
- 10 % w/w NiO/MWNTs	170.22	169.78	170.05	170.01

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	1	2	3	
- 2 : 8 % w/w FeO-TiO ₂ /MWNTs	123.09	122.52	122.21	122.60
- 4 : 6 % w/w FeO-TiO ₂ /MWNTs	122.47	121.68	122.04	122.06
- 6 : 4 % w/w FeO-TiO ₂ /MWNTs	121.42	121.49	121.66	121.52
- 8 : 2 % w/w FeO-TiO ₂ /MWNTs	121.34	120.08	120.51	120.64
- 2 : 8 % w/w NiO-TiO ₂ /MWNTs	177.45	177.49	177.03	177.32
- 4 : 6 % w/w NiO-TiO ₂ /MWNTs	180.42	180.03	179.01	179.82
- 6 : 4 % w/w NiO-TiO ₂ /MWNTs	179.54	179.22	179.06	179.27
- 8 : 2 % w/w NiO-TiO ₂ /MWNTs	179.00	178.07	178.15	178.40

APPENDIX C

Results from galvanostatic charge-discharge technique

Table C Specific capacitance from galvanostatic charge-discharge

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	1	2	3	
- Refluxed MWNTs with 6 hours	53.38	54.04	53.17	55.53
- Refluxed MWNTs with 12 hours	66.42	67.74	67.09	67.08
- Refluxed MWNTs with 18 hours	80.78	81.53	80.99	81.10
- 2 % w/w TiO ₂ /MWNTs	141.75	140.03	140.04	140.60
- 4 % w/w TiO ₂ /MWNTs	149.95	149.92	149.81	149.89
- 6 % w/w TiO ₂ /MWNTs	172.56	170.08	170.70	171.11
- 8 % w/w TiO ₂ /MWNTs	175.63	175.33	176.02	175.66
- 10 % w/w TiO ₂ /MWNTs	177.80	176.07	176.19	176.68
- 2 % w/w FeO/MWNTs	97.46	97.57	99.84	98.29
- 4 % w/w FeO/MWNTs	103.74	102.51	102.94	103.06
- 6 % w/w FeO/MWNTs	103.85	103.61	103.00	103.48
- 8 % w/w FeO/MWNTs	104.26	104.39	104.88	104.51
- 10 % w/w FeO/MWNTs	105.19	104.95	105.01	105.05
- 2 % w/w NiO/MWNTs	143.69	144.84	143.09	143.87
- 4 % w/w NiO/MWNTs	150.65	149.16	149.01	149.60
- 6 % w/w NiO/MWNTs	154.51	155.49	155.04	155.01
- 8 % w/w NiO/MWNTs	166.86	167.06	167.91	167.27
- 10 % w/w NiO/MWNTs	170.08	170.65	170.55	170.42

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	1	2	3	
- 2 : 8 % w/w FeO-TiO ₂ /MWNTs	122.37	122.18	122.32	122.29
- 4 : 6 % w/w FeO-TiO ₂ /MWNTs	122.06	122.14	122.04	122.08
- 6 : 4 % w/w FeO-TiO ₂ /MWNTs	121.55	120.01	121.48	121.01
- 8 : 2 % w/w FeO-TiO ₂ /MWNTs	120.35	119.76	120.02	120.04
- 2 : 8 % w/w NiO-TiO ₂ /MWNTs	175.63	177.52	177.92	177.02
- 4 : 6 % w/w NiO-TiO ₂ /MWNTs	180.88	180.69	179.33	180.30
- 6 : 4 % w/w NiO-TiO ₂ /MWNTs	179.47	179.09	179.56	179.37
- 8 : 2 % w/w NiO-TiO ₂ /MWNTs	175.76	178.46	178.48	177.56

APPENDIX D

Results from three electrochemical techniques

Table D Specific capacitance from three electrochemical techniques

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	CV	EIS	Galvano static	
- Refluxed MWNTs with 6 hours	57.97	55.58	55.53	56.36
- Refluxed MWNTs with 12 hours	69.76	66.96	67.08	67.93
- Refluxed MWNTs with 18 hours	82.55	79.98	81.10	81.21
- 2 % w/w TiO ₂ /MWNTs	144.45	140.27	140.60	141.77
- 4 % w/w TiO ₂ /MWNTs	155.75	150.17	149.89	151.93
- 6 % w/w TiO ₂ /MWNTs	173.53	170.90	171.11	171.84
- 8 % w/w TiO ₂ /MWNTs	176.56	175.62	175.66	175.94
- 10 % w/w TiO ₂ /MWNTs	178.15	177.13	176.68	177.32
- 2 % w/w FeO/MWNTs	101.22	99.15	98.29	99.55
- 4 % w/w FeO/MWNTs	103.39	102.81	103.06	103.08
- 6 % w/w FeO/MWNTs	104.68	103.60	103.48	103.92
- 8 % w/w FeO/MWNTs	105.05	104.67	104.51	104.74
- 10 % w/w FeO/MWNTs	105.75	105.03	105.05	105.27
- 2 % w/w NiO/MWNTs	146.89	144.73	143.87	145.16
- 4 % w/w NiO/MWNTs	152.18	149.97	149.60	150.52
- 6 % w/w NiO/MWNTs	158.06	155.20	155.01	156.09
- 8 % w/w NiO/MWNTs	168.92	166.84	167.27	167.67
- 10 % w/w NiO/MWNTs	171.91	170.01	170.42	170.78

Electrode Material	Specific Capacitance (F/g)			Average Capacitance (F/g)
	CV	EIS	Galvano static	
- 2 : 8 % w/w FeO-TiO ₂ /MWNTs	124.15	122.60	122.29	123.01
- 4 : 6 % w/w FeO-TiO ₂ /MWNTs	123.78	122.06	122.08	122.64
- 6 : 4 % w/w FeO-TiO ₂ /MWNTs	122.92	121.52	121.01	121.81
- 8 : 2 % w/w FeO-TiO ₂ /MWNTs	121.75	120.64	120.04	120.81
- 2 : 8 % w/w NiO-TiO ₂ /MWNTs	179.72	177.32	177.02	178.02
- 4 : 6 % w/w NiO-TiO ₂ /MWNTs	181.09	179.82	180.30	180.40
- 6 : 4 % w/w NiO-TiO ₂ /MWNTs	180.23	179.27	179.37	179.62
- 8 : 2 % w/w NiO-TiO ₂ /MWNTs	180.06	178.40	177.56	178.67

VITA

Mr. Narabhandhu Laohawich was born on March 26, 1984 in Nonthaburi, Thailand. He received his Bachelor's degree of Engineering (Chemical Engineering) from Srinakharinwirot University in 2006.

Since 2006, he has become a graduate student of Program in Petrochemistry and Polymer Science, Faculty of Science, Chulalongkorn University and worked under the supervision of Associate Professor Dr. Orawon Chailapakul. Then in 2008, he received the scholarship from Chulalongkorn University Graduate School Thesis Grant. He graduated the Master of Science degree from Chulalongkorn University in the academic year 2008.

His present address is 25/560 Changwatthana Road, Bangtalard, Pakkred, Nonthaburi, Thailand 11120, Tel. 02-5840874, 086-5108382.