

REFERENCES

- ASTM F739-91. (1991). Standard test method for resistance of protective clothing materials to permeation by liquids of gases under conditions of continuous contact. <u>Annual Book of ASTM Standards</u>. Philadelphia: The American Society for Testing and Materials.
- Berthiaume, S., Anderson, E., and Yoshida, Y. (2000). Chlorine/Sodium Hydroxide. Chemical Economics Handbook Report.
- Bitter, J.G.A. (1991). <u>Transport Mechanisms in Membrane Separation Processes</u>. New York: Plenum Press.
- Centre for Nuclear Energy Research. (1998). Development of a solid state sensor for monitoring hydrogen in humid chlorine gas. <u>Proposal to the Chlorine Institute</u>.
- Centre for Nuclear Energy Research. (1999). Development of a solid state sensor for monitoring hydrogen in humid chlorine gas. <u>Final report of phase I to the Chlorine Institute</u>.
- Centre for Nuclear Energy Research. (2000). Development of a solid state sensor for monitoring hydrogen in humid chlorine gas. <u>Supplementary report and</u> research proposal to the Chlorine Institute.
- Centre for Nuclear Energy Research. (2000). Development of a sensor for Monitoring hydrogen in humid chlorine gas. <u>Final report of phase II to the</u> <u>Chlorine Institute.</u>
- Centre for Nuclear Energy Research. (2000). Development of a sensor for Monitoring hydrogen in humid chlorine gas. <u>Status report to the Chlorine</u> <u>Institute</u>
- Centre for Nuclear Energy Research. (2000). Development of a sensor for Monitoring hydrogen in humid chlorine gas. <u>Progress report to the Chlorine</u> <u>Institute</u>.
- Crank, J. and Park, G.S. (1968). <u>Diffusion in Polymers</u>. London and New York: Academic Press.

- Dhingra, S.S., and Marand, E. (1998). Mixed gas transport study through polymeric membranes. Journal of Membrane Science, 141, 45-63.
- Freeman, B., and Pinnau, I. (1997). Separation of gases using solubility-selective polymers. <u>TRIP</u>, 5(5), 167-173.
- George, S.C., and Thomas, S. (2001). Transport phenomena through polymeric systems. <u>Progress in Polymer Science</u>, 26, 291-1017.
- Hagg, M-B. (2000). Membrane purification of Cl₂ gas I. Permeabilities as function of temperature for Cl₂, O₂, N₂, and H₂ in two types of PDMS membranes. Journal of Membrane Science, 170, 173-190.
- Hagg, M-B. (2000). Membrane purification of Cl₂ gas II. Permeabilities as function of temperature for Cl₂, O₂, N₂, H₂ and HCl in perfluorinated, glass and carbon molecular sieve membranes. <u>Journal of Membrane Science</u>, 177, 109-128.
- Hagg, M-B. (2001). Purification of chlorine gas with membranes-an integrated process solution for magnesium production. <u>Separation and Purification</u> <u>Technology</u>, 21, 261-278.
- Ismail, A.F., and David, L.I.B. (2001). Review-A review on the latest development of carbon membranes for gas separation. Journal of Membrane Science, 193, 1-18.
- Kesting, R.E., and Fritzsche, A.K. (1993). <u>Polymeric gas separation Membranes</u>. New York : John Wiley.
- Kham Sa-ang, K (2001). <u>Development of a probe for measuring hydrogen</u> <u>concentration in the presence of moist chlorine</u>. M.S. Thesis for The Petroleum and Petrochemical college, Chulalongkorn University.
- Kinemuchi, Y., Suzuki, T., Jiang, W., and Yatsui, K. (2001). Ceramic membrane filter using ultrafine powders. Journal of American Ceramic Society, 84(9), 2144-2146.
- Kitjaroenvong, P. (2000). <u>Development of a probe for measuring hydrogen</u> <u>Concentration in the presence of moist chlorine</u>. M.S. Thesis for The Petroleum and Petrochemical college, Chulalongkorn University.

- Klopffer, M.H. and Flaconneche, B. (2001). Transport Properties of Gases in Polymers: Bibliographic Review. <u>Oil & Gas Science and Technology</u>, 56(3), 223-244.
- Koros, W.J., and Chern, R.T. (1987). Separation of Gaseous Mixtures Using Polymer Membranes. <u>Handbook of Separation Process Technology</u>, 20, 862-953.
- Kulprathipanja, S., Nenzil, R.W., and Li, N.N. (1988). U.S. Patent 4740219.
- Lokhandwala, K.A., Segelke, S., Nguyen, P., Baker, R.W., Su, T.T., and Pinnau, I. (1999). A membrane process to recover chlorine from chloralkali plant tail gas. <u>Industrial Engineering Chemistry Research</u>, 38, 3606-3613.
- Matsuura, T. (1994). <u>Synthetic Membranes and Membrane Separation Processes</u>. Florida: CRC Press, Inc.
- Mulder, M.H.V. (1991). <u>Basic Principles of Membrane Technology</u>. Dordrecht: Kluwer Academic Publisher.
- Naylor, T.V. (1989). Permeation Properties. <u>Comprehensive Polymer Science</u>, 2, 643-668.
- Ohashi, H., Ohya, H. Aihara, M., Negishi, Y, and Semenova, S.I. (1998). Hydrogen production from hydrogen sulfide using membrane reactor integrated with porous membrane having thermal and corrosion resistance. <u>Journal of</u> <u>Membrane Science</u>, 146, 39-52.
- Pinto, C.G., Laespada, M.E.F., Pavon, J.L.P., and Cordero, B.M. (1999). Review analytical applications of separation techniques through membranes. <u>Laboratory Automation and Information Management</u>, 34, 115-130.
- Rakotoarevelo, J. (1995). Literature Search for Diffusion Coefficient of Permeability of Hydrogen through Teflon and Palladium Membranes at Elevated Temperatures. <u>Paper presented at Centre for Nuclear Energy Research</u>, Canada.
- Raymond, K.E. (1998). Alkali and chlorine products. <u>Encyclopedia of Chemical</u> <u>Technology</u>, 1, 984-989.

- Sarner, S.F., and Henry III, N.W. (1989). The use of detector tubes following ASTM method F-739-85 for measuring permeation resistance of clothing. <u>American Industrial Hygiene Association Journal</u>, 50(6), 298-302.
- Tufts, L.E. (1992). Reaction Characteristics of Hydrogen Admixed with Chlorine, Oxygen and Inert gases. <u>Explosive Properties of Gaseous Mixtures</u> <u>Containing Hydrogen and Chlorine</u>. Washington DC: The Chlorine Institute Inc.
- Tuwiner, S.B., Miller, L.P., and Brown, W.E. (1962). <u>Diffusion_and_Membrane</u> <u>Technology</u>. New York: Reinhold Publishing Corporation.
- Vieth, W.R. (1991). Diffusion in and through polymers. Munich: Hanser.
- Zimmerman, C.M., Singh, A., and Koros, W.J. (1997). Tailoring mixed matrix composite membranes for gas separations. Journal of Membrane Science, 137, 145-154.

APPENDICES

Appendix A Properties of Materials

The following materials were used for the determination of hydrogen and chlorine permeability coefficients.

A.1 Teflon

Common Names: Polytetrafluoroethylene, PTFE Density = 2.280-2.290 2 g/cm³ Viscosity (at 25°C) = 3000 centipoises Upper Working Temperature = 260°C Limiting Oxygen Index = 95% Molecular Weight = 99.96

A.2 Derakane® 470-300

Common Name: Epoxy vinyl ester resin Viscosity, cps at 25°C = 300 Specific gravity = 1.08 Tensile strength, psi at room temperature = 12,400 Tensile elongation, % at room temperature = 3.5 Heat Distortion Temperature, °F at 1.82Mpa applied stress = 152-305

A.3 Derakane® 8084

Common Name: Epoxy vinyl ester resin Viscosity, cps at 25°C = 350 Specific gravity = 1.02 Tensile strength, psi at room temperature = 10,500 Tensile elongation, % at room temperature = 8.0/11.0 Heat Distortion Temperature, °F at 1.82Mpa applied stress = 85-175

Appendix B Properties of normal-hydrogen gas

Hydrogen gas was tested to determine the permeation of hydrogen to the test materials.

Normal-Hydrogen gas Physical properties Formula: H₂ Atomic number: Z=1 Molecular weight: 2.016 g/mole Ionization potential: 15.427 eV Density S.T.P: 0.08989 kg/m³ @ Operational condition: 80°C and 1atm Density: 0.06956 kg/m³ Heat capacity at constant pressure (Cp): 3.452 kcal/kg K Heat capacity at constant volume (Cv): 2.463 kcal/kg K Viscosity: 9.92×10-5 poises Thermal conductivity (k): 49.1276×10⁻⁵ cal/cm sec K Melting point: -259.1°C Boiling point: -252.7°C

Appendix C Permeability coefficients of Teflon

Teflon was used as supporting sheets for Derakane® sandwich membranes so its hydrogen and chlorine permeability coefficients must be determined. The resistance model described in chapter 2 (section 2.2.4) was used in order to calculate the permeability coefficients of Derakane®. Table A3 shows the hydrogen and chlorine permeability coefficients of different types of Teflon at 25, 50 and 80°C.

Table C3 H_2 and Cl_2 permeability coefficients of different types of Teflon® at different temperatures.

Gas Mixture	T (°C)	P _{H2}					P _{Cl2}	
		0.001 in. (1)*	0.001 in. (2)*	0.002 in.	Etched 0.002 in. (#1)**	Etched 0.002 in (#2)**	0.15 mm	0.002 in.
5% H ₂ in Ar	25	4.98E-6	3.83E-6	3.49E-6	2.59E-6	2.66E-6	-	
	50	-	6.06E-6	6.20E-6	-	4.97E-6	-	
	80	2.18E-5	1.15E-5	1.20E-5	-	9.29E-6	-	
100 % Cl ₂	25	-	-	-	-	-	1.06E-7	-
	50	-	-	-	-	-	1.45E-7	-
	80	-	-	-	-	-	1.86E-7	-
5% H ₂ in Cl ₂	80	-	-	3.50E-5	-	-	-	5.85E-8

Note: Unit of the permeability coefficients is cm³(STP).cm.cm⁻².min⁻¹.atm.

*There were two lots of 0.001 in. thick Teflon® used and tested in this study.

** The etched Teflon® of 0.002 in. thick were tested twice (first in August and then in January) to verify the stability of the system.

CURRICULUM VITAE

Name: Ms. Suchada Passararat

Date of Birth: January 8, 1979

Nationality: Thai

University Education:

1996-1999 Bachelor Degree of Science in Chemical Engineering,Department of Chemical Technology, Faculty of Science,Chulalongkorn University, Bangkok, Thailand

