

REFERENCES

- Anderson, J.A., and Garcia, M.F. (2005). Supported Metals in Catalysis. London: Imperial College Press.
- Beyer, H.K. (2002). Molecular Sieves Vol.3. Berlin: Springer-Verlag.
- Canizares, P., Lucas, A., Dorado, F., Duran, A., and Asencio, I. (1997). Characterization of Ni and Pd supported on H-mordenite catalysts: Influence of the metal loading method. Applied Catalysis A: General, 169, 137-150.
- Chankham, O. (2007). Hydrogen production via steam reforming of methane over Ni supported-NaY zeolite catalyst. M.S. Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Chan, J.-S., Park, S.-E., and Chon, H. (1996). Catalytic activity and coke resistance in the CO₂ reforming of methane to synthesis gas over zeolite-supported Ni catalysts. Applied Catalysis A: General, 145, 111-124.
- Chen, Y.-G., Tomishige, K., Yokoyama, K., and Fujimoto, K. (1999). Catalytic performance and catalyst structure of Ni-Magnesia catalysts for CO₂ reforming of methane. Journal of Catalysis, 184, 479-490.
- Comas, J., Dieuzeide, M.L., Baronetti, G., Laborde, M., and Amadeo, N. (2006). Methane steam reforming and ethanol steam reforming using a Ni(II)-Al(III) catalyst prepared from lamellar double hydroxides. Chemical Engineering Journal, 118, 11-15.
- Craciun, R., Daniell, W., and Knozinger, H. (2002). The effect of CeO₂ structure on the activity of supported Pd catalysts used for methane steam reforming. Applied Catalysis A: General, 230, 153-168.
- Crisafulli, C., Scire, S., Minico, S., Solarino, L. (2002). Ni-Ru bimetallic catalysts for the CO₂ reforming of methane. Applied Catalysis A: General, 225, 1-9.
- Damyanova, S., Perez, C.A., Schmal, M., and Bueno, J.M.C. (2002). Characterization of ceria-coated alumina carrier. Applied Catalysis A: General, 234, 271-282.
- Dong, W.-S., Roh, H.-S., Jun, K.-W., Park, S.-E., and Oh, Y.-S. (2002). Methane reforming over Ni/Ce-ZrO₂ catalysts: effect of Ni content. Applied Catalysis A: General, 226, 63-72.

- Feio, L.S.F., Hori, C.E., Damyanova, S., Noronha, F.B., Cassinelli, W.H., Marques, C.M.P., and Bueno, J.M.C. (2007). The effect of ceria content on the properties of Pd/CeO₂/Al₂O₃ catalysts for steam reforming of methane. Applied Catalysis A: General, 316, 107-116.
- Gates, B.C. (1992). Catalytic Chemistry. New York: John Wiley & Sons, Inc.
- Gordeeva, L.G., Aristov, Y.I., Moroz, E.M., Rudina, N.A., Zaikovskii, V.I., Tanashev, Y.Y., and Parmon, V.N. (1994). Preparation and study of porous uranium oxides as supports for new catalysts of steam reforming of methane. Journal of Nuclear Materials, 218, 202-209.
- Goud, S.K., Whittenberger, W.A., Chattopadhyay, S., and Abraham, M.A. (2007). Steam reforming of n-hexadecane using a Pd/ZrO₂ catalyst : kinetics of catalyst deactivation. International Journal of Hydrogen Energy, 32, 2868-2874.
- Hegarty, M.E.S., O' Conner, A.M., and Ross, J.R.H. (1998). Syngas production from natural gas using ZrO₂-supported metals. Catalysis Today, 42, 225-232.
- Kusakabe, K., Sotowa, K.I., Eda, T., and Iwamoto, Y. (2004). Methane steam reforming over Ce-ZrO₂-supported noble metal catalysts at low temperature. Fuel Processing Technology, 86, 319-326.
- Levent, M., Gunn, D.J., Ali El-Bousiffi, M. (2002). Production of hydrogen-rich gases from steam reforming of methane in an automatic catalytic microreactor. International Journal of Hydrogen Energy, 1-15.
- Le Page, J.-F., Cosyns, J., Courty, P., Freund, E., Franck, J.-P., Jacquin, Y., Juguin, B., Marcilly, C., Martino, G., Miquel, J., Montarnal, R., Sugier, A., and Van Landegham, H. (1978). Applied Heterogeneous Catalysis. Paris: Editions Technip.
- Laosiripojana, N., Sutthisripok, W., Assabumrungrat, S. (2005). Synthesis gas production from dry reforming of methane over CeO₂ doped Ni/Al₂O₃ : Influence of the doping ceria on the resistance toward carbon formation. Chemical Engineering Journal, 112, 13-22.
- Matar, S., and Hatch, L. (1994). Chemistry of petrochemical processes. Texas: Gulf Publishing Company.

- Matsumura, Y., and Nakamori, T. (2004). Steam reforming of methane over Ni catalysts at low reaction temperature. *Applied Catalyst A: General*, 258, 107-114.
- Oh, Y.-S., Roh, H.-S., Jun, K.-W., and Baek, Y.-S. (2003). A highly active catalyst, Ni/Ce-ZrO₂/θ-Al₂O₃, for on-site H₂ generation by steam methane reforming: pretreatment effect. *International Journal of Hydrogen Energy*, 1-6.
- Querini, C.A. (2004). *Catalysis, volume 17*. The Royal Society of Chemistry.
- Quincoces, C.E., Vargas, S.P., Grange, P., Gonzalez, M.G. (2002). Role of Mo in CO₂ reforming of CH₄ over Mo promoted Ni/Al₂O₃ catalysts. *Materials Letters*, 56, 698-704.
- Roh, H.-S., Jun, K.-W., and Park, S.-E. (2003). Methane-reforming reactions over NiO/Ce-ZrO₂/θ-Al₂O₃ catalysts. *Applied Catalysis A: General*, 251, 275-283.
- Rostrup-nielsen. (1984). *Catalysis science and technology*. New York: Splinkger.
- Satterfield, C.N. (1991). *Heterogeneous Catalysis in Industrial Practice*. New York: McGraw-Hill, Inc.
- Senathipbodee, N. (2004). Hydrogen production from steam reforming of methane using Ni-supported KL zeolite catalysts. *M.S. Thesis*, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Sturzenegger, M., D'Souza, L., Struis, R.P.W.J., and Stucki, S. (2006). Oxygen transfer and catalytic properties of Ni iron oxides for steam reforming of methane. *Fuel*, 85, 1599-1602.
- Tosheva, L. (1999). Zeolite macrostructure. *Licentiate Thesis*, Chemical technology department of chemical and metallurgical engineering, Lulea University, Sweden.
- Tosiri, S. (2006). Hydrogen production from CO₂ reforming of methane using Ni-supported KL zeolite catalysts. *M.S. Thesis*, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Viswanathan, B., Sivasanker, S., and Ramaswamy, A.V. (2002). *Catalysis Principles and Applications*. New Delhi: Narosa Publishing House.

- Wang, S., and Lu, G.Q.M. (1997). CO₂ reforming of methane on Ni catalysts: effects of the support phase and preparation technique. Applied Catalysis B: Environmental, 16, 269-277.
- Yamazaki, O., Tomishige, K., and Fujimoto, K. (1996). Development of highly stable Ni catalyst for methane-steam reaction under low steam to carbon ratio. Applied Catalysis A: General, 136, 49-56.
- http://ec.europa.eu/research/rtdinfo/42/01/article_1315_en.html
- <http://www.3dchem.com/molecules.asp>
- <http://www.uni-leipzig.de/~pdfhome/zsm1.html>

APPENDICES

Appendix A Calculations

1. Catalyst Preparation

1.1 Amount of Ni loading

Example Prepared 1 g of 3 wt% Ni/ZSM-5 catalyst;

- Amount of Ni (MW = 58.70 g/mole)

$$\begin{aligned} \text{Ni} &= 1 * (3/100) \text{ g} \\ &= 0.03 \text{ g} \end{aligned}$$

- Amount of Ni $(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (MW = 290.81 g/mole)

$$\begin{aligned} \text{Ni } (\text{NO}_3)_2 \cdot 6\text{H}_2\text{O} &= 0.03 * (290.81/58.70) \text{ g} \\ &= 0.1486 \text{ g} \end{aligned}$$

- Amount of ZSM-5 Zeolite

$$\text{ZSM-5 Zeolite} = 1 - 0.03 = 0.97 \text{ g}$$

1.2 Amount of CeO₂ loading

Example Prepared 1g of 3%Ce-11% Ni/ZSM-5 catalyst;

- Amount of Ni (MW = 58.70 g/mole)

$$\begin{aligned} \text{Ni} &= 1 * (11/100) \text{ g} \\ &= 0.11 \text{ g} \end{aligned}$$

- Amount of Ni $(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (MW = 290.81 g/mole)

$$\begin{aligned} \text{Ni } (\text{NO}_3)_2 \cdot 6\text{H}_2\text{O} &= 0.11 * (290.81/58.70) \text{ g} \\ &= 0.5450 \text{ g} \end{aligned}$$

- Amount of Ce (MW = 140.12 g/mole)

$$\begin{aligned} \text{Ce} &= 1 * (3/100) \text{ g} \\ &= 0.03 \text{ g} \end{aligned}$$

- Amount of Ce(NO_3)₃ · 6H₂O (MW = 434.23 g/mole)

$$\begin{aligned} \text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O} &= 0.03 * (434.23/140.12) \text{ g} \\ &= 0.0930 \text{ g} \end{aligned}$$

- Amount of ZSM-5 Zeolite

$$\text{ZSM-5 Zeolite} = 1 - 0.11 - 0.03 = 0.86 \text{ g}$$

2. Steam- to-Methane Ratio for Feed

For mixture system;

Assume: 1:1 H₂O: CH₄ molar ratio;

Where Total flow = 100 ml/min

CH₄ flow = 20 ml/min

Steam flow = 20 ml/min

He balance = 60 ml/min

Finding the flow rate of water in feed;

From thermodynamic steam table;

At 160°C, V_g = 0.3068 m³/kg (D_g = 0.00326 g/cm³)

At 25°C, V_f = 0.001003 m³/kg (D_f = 0.997 g/cm³)

At evaporator (water feed is in the steam phase)

$$\begin{aligned} \text{Then, } M_g &= D_g * V_g \\ &= 0.00326 \text{ g/cm}^3 * 20 \text{ ml} \\ &= 0.0652 \text{ g} \end{aligned}$$

Thus, water in liquid phase as a feed is

$$\begin{aligned} V_l &= M_l / D_l \\ &= 0.0652 \text{ g} / 0.997 \text{ g/cm}^3 \\ &= 0.0654 \text{ ml/min.} = 3.9 \text{ ml/hr.} \end{aligned}$$

3. Metal Crystallite Size from XRD

The thickness of crystallite (L) calculated from Scherrer Equation;

$$L_{k\bar{h}\bar{l}} = k\lambda / (\beta \cos \theta_0)$$

Where λ = the x-ray wavelength

B = the peak width (expressed in radian)

θ₀ = the angle between the beam and the normal on the reflecting plane
(expressed in radian)

k = a constant (or shape factor) (often take as 1)

4. Average Particle Size from TEM

The average particle size (d_s) is calculated on the basis of counting the diameter of numerous particles, according to the equation;

$$d_s = \sum n_i d_i^3 / \sum n_i d_i^2$$

Where n_i = number of particles in each size range

d_i = diameter of particle in each size range

Moreover, the particle size distribution is also calculated from:

$$\frac{\text{Number of particles in range}}{\text{Total number of particles}} \times 100\%$$

Appendix B Effect of Ni loading on reduced catalysts for XRD

These XRD patterns show that there are only Ni peaks for the reduced catalysts and the Ni peak intensity is increased with the increasing of Ni loading. It confirms the complete reduction of the catalysts before the reforming reaction was started.

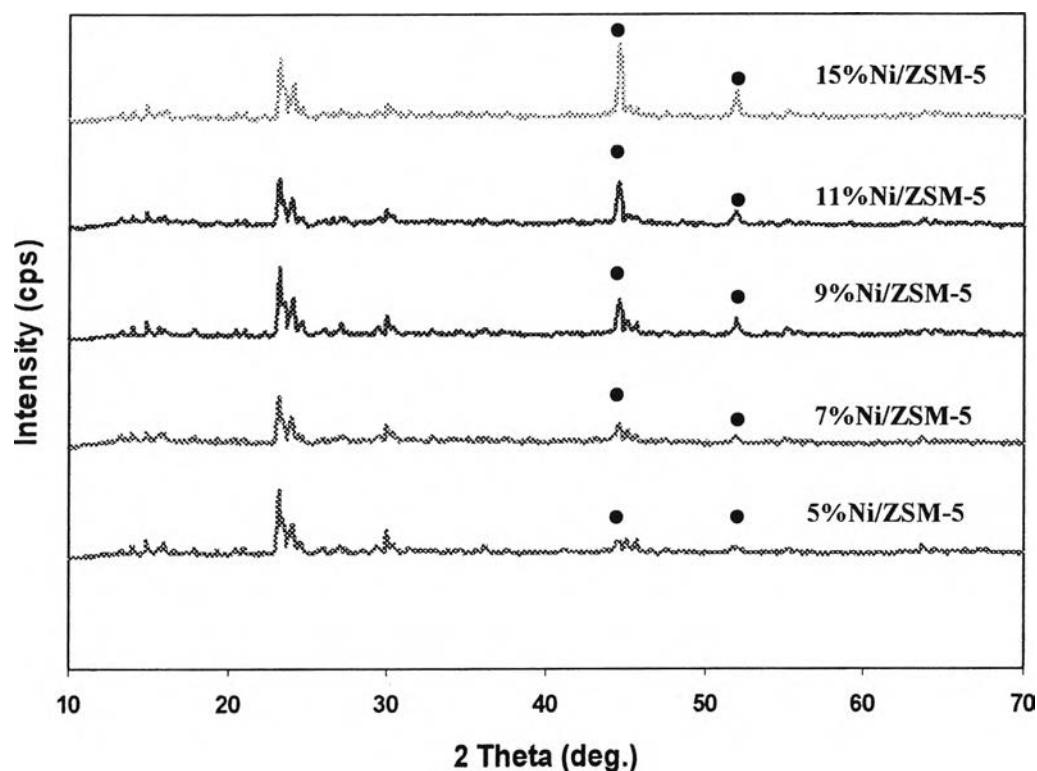


Figure B1 XRD patterns of the reduced Ni/ZSM-5 catalysts with various loadings of Ni, which operated on steam reforming reaction at 700°C and atmospheric pressure for 5 hours; (●), Ni metal phase.

Appendix C Particle size distribution from TEM

1. Effect of Ni loading

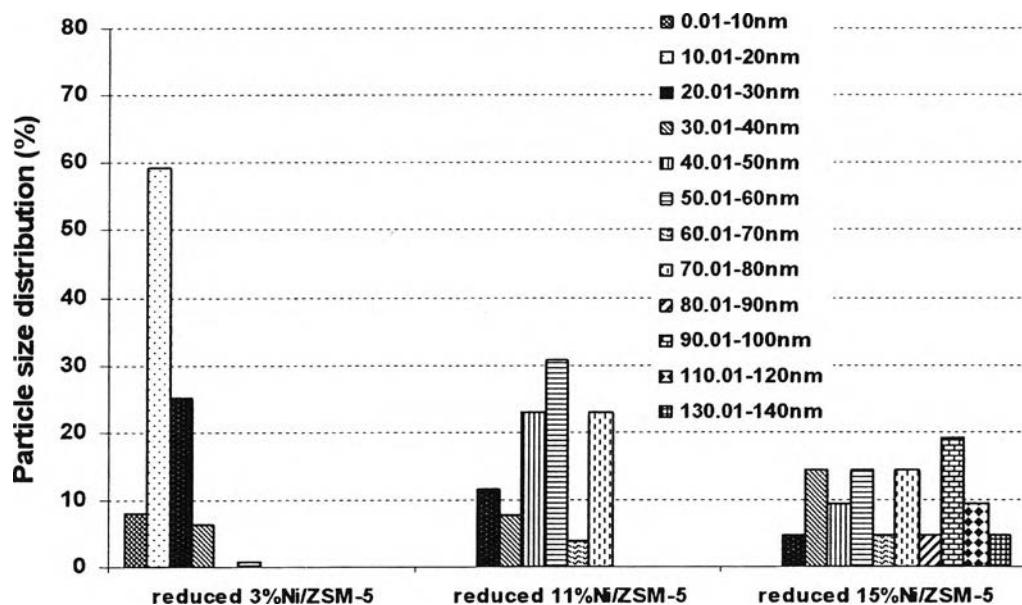


Figure C1 Particle size distribution of the reduced Ni/ZSM-5 catalysts with various loadings of Ni which were calculated from TEM results.

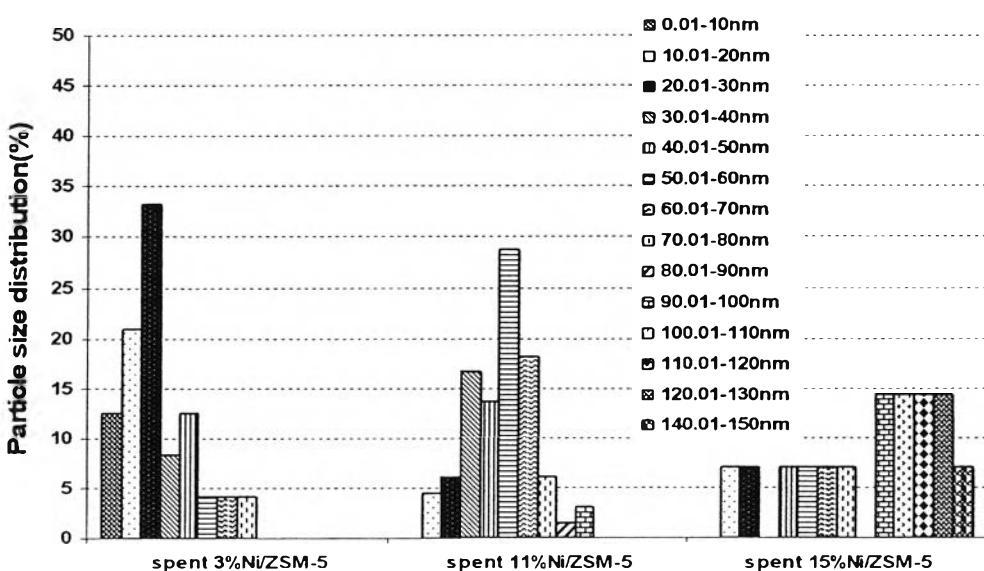


Figure C2 Particle size distribution of the spent Ni/ZSM-5 catalysts with various loadings of Ni which were calculated from TEM results.

2. Effect of CeO₂ addition

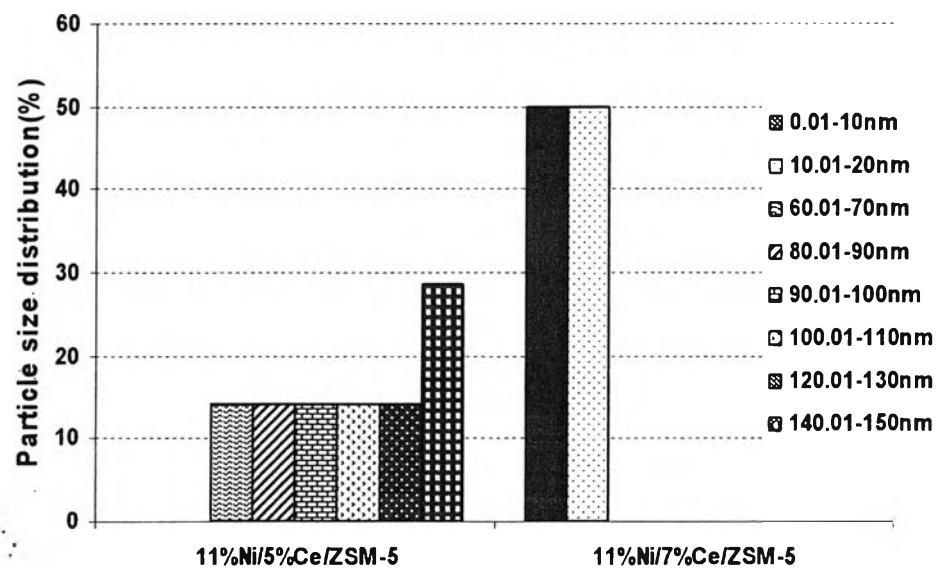


Figure C3 Particle size distribution of the spent Ni/ZSM-5 catalysts with various loadings of CeO₂ which were calculated from TEM results.

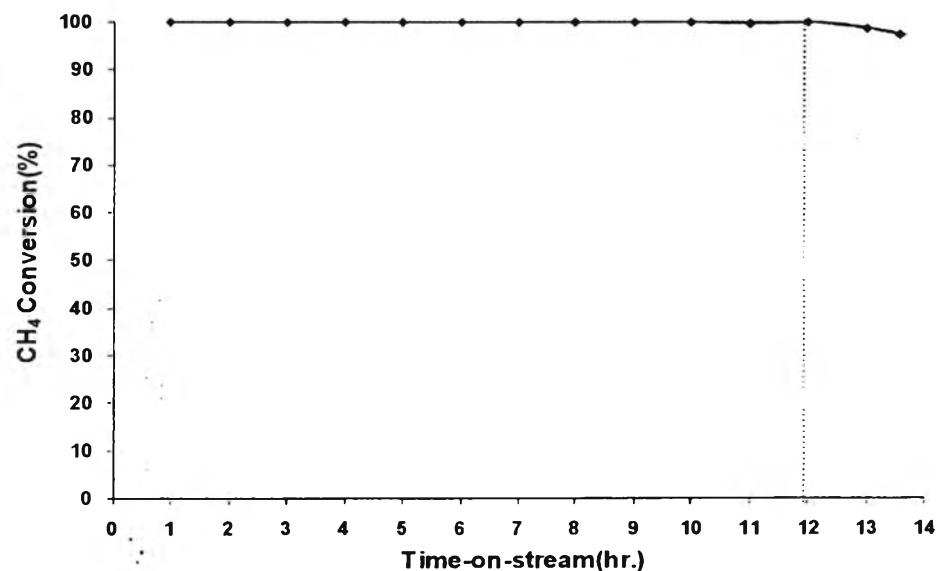
Appendix D Stability test of the 11%Ni/5%Ce/ZSM-5 catalyst

Figure D1 CH₄ conversion of the 11%Ni/5%Ce/ZSM-5 catalyst operated on the reforming reaction at 700°C and a H₂O/CH₄ ratio of 0.8 for the stability test.

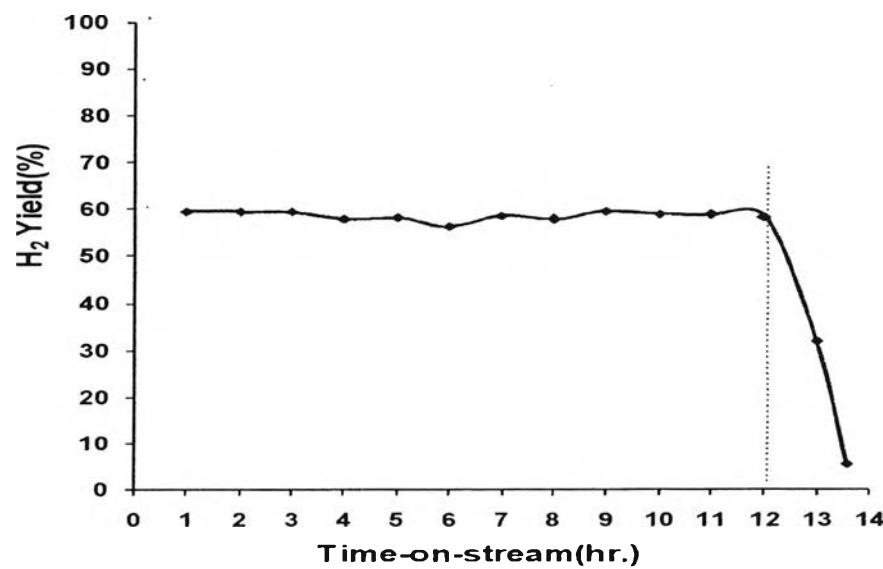


Figure D2 H₂ yield of the 11%Ni/5%Ce/ZSM-5 catalyst operated on the reforming reaction at 700°C and a H₂O/CH₄ ratio of 0.8 for the stability test.

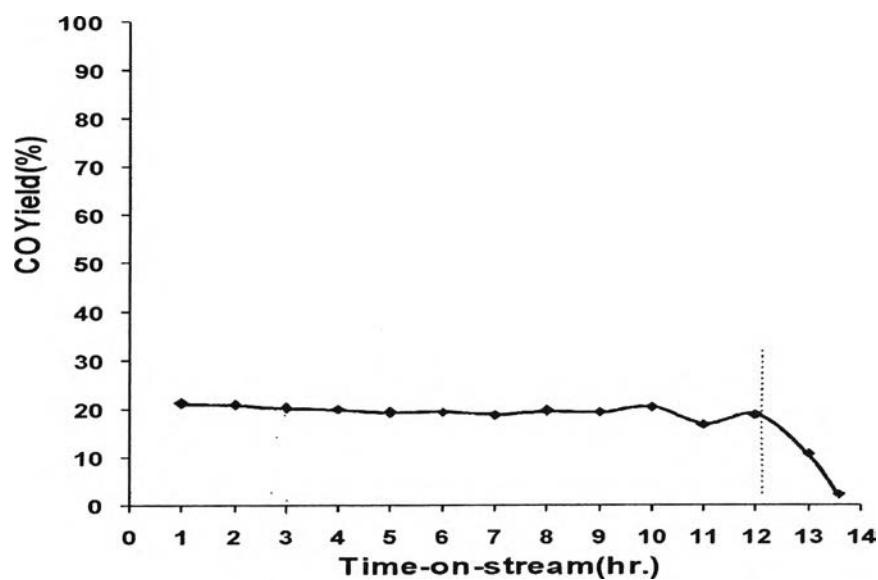


Figure D3 CO yield of the 11%Ni/5%Ce/ZSM-5 catalyst operated on the reforming reaction at 700°C and a H₂O/CH₄ ratio of 0.8 for the stability test.

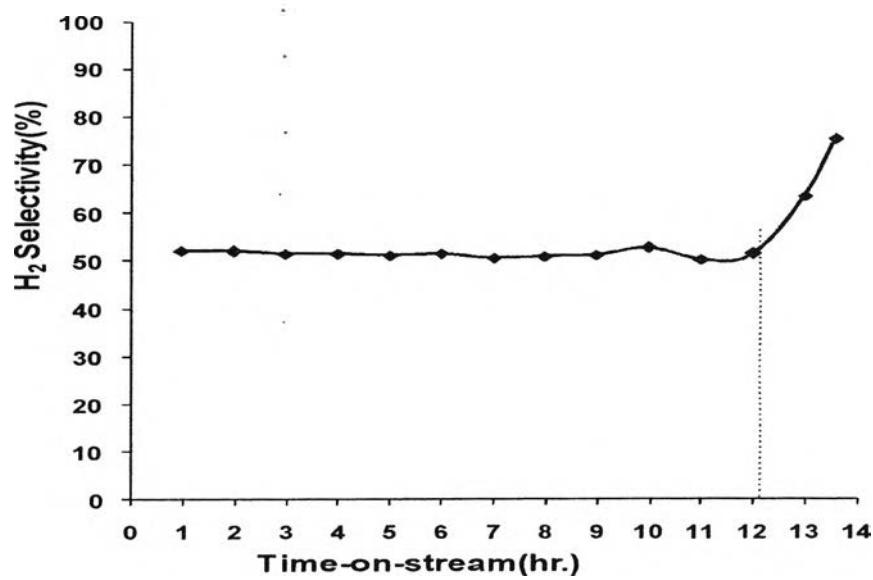


Figure D4 H₂ Selectivity of the 11%Ni/5%Ce/ZSM-5 catalyst operated on the reforming reaction at 700°C and a H₂O/CH₄ ratio of 0.8 for the stability test.

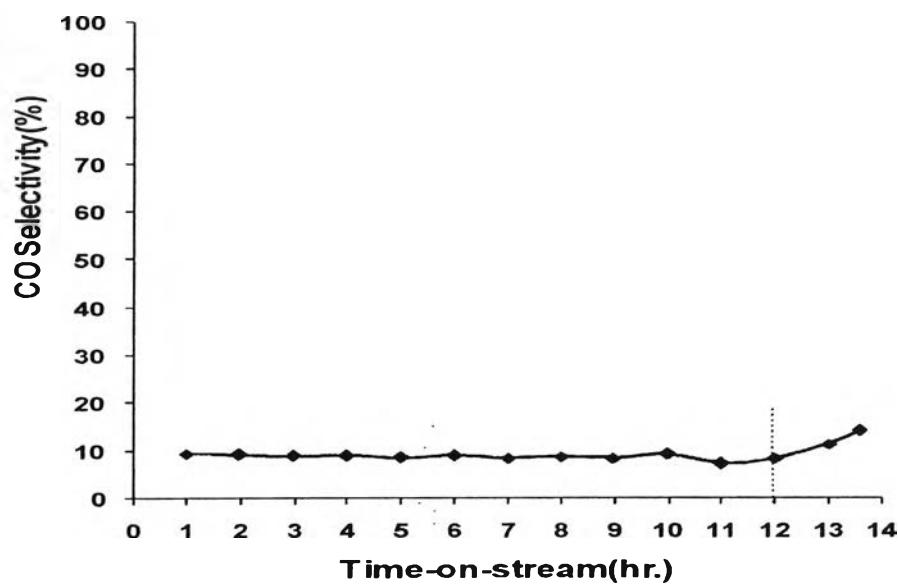


Figure D5 CO selectivity of the 11%Ni/5%Ce/ZSM-5 catalyst operated on the reforming reaction at 700°C and a H₂O/CH₄ ratio of 0.8 for the stability test.

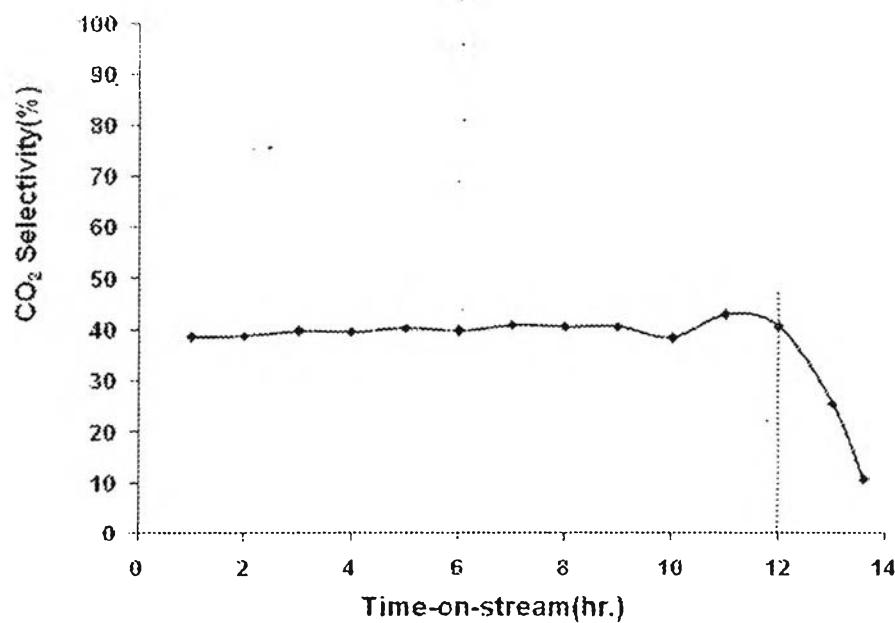


Figure D6 CO₂ selectivity of the 11%Ni/5%Ce/ZSM-5 catalyst operated on the reforming reaction at 700°C and a H₂O/CH₄ ratio of 0.8 for the stability test.

Appendix E CO₂ selectivity of the catalysts in all effect studies

1. Effect of Ni loading

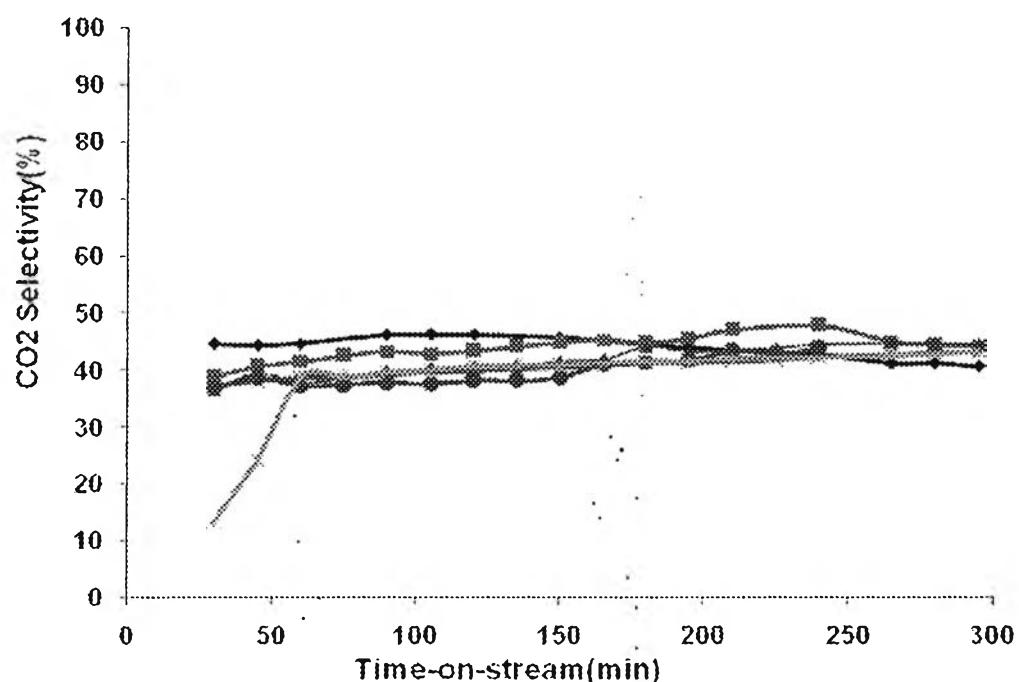


Figure E1 CO₂ selectivity of the Ni/ZSM-5 catalysts with various loadings of Ni, which operated on steam reforming reaction at 700°C and atmospheric pressure for 5 hours; (-◆-), 3%Ni/ZSM-5; (-■-), 5%Ni/ZSM-5; (-▲-), 7%Ni/ZSM-5; (-×-), 9%Ni/ZSM-5; (-●-), 11%Ni/ZSM-5; (-✗-), 15%Ni/ZSM-5.

2. Effect of H₂O/CH₄ ratio

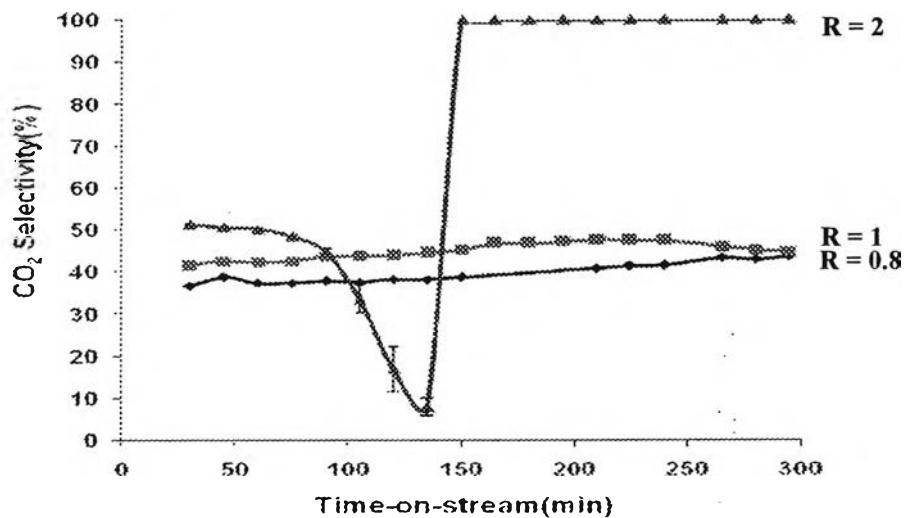


Figure E2 CO₂ selectivity of the 11%Ni/ZSM-5 catalysts with different H₂O/CH₄ ratios, which operated on steam reforming reaction at 700°C and atmospheric pressure for 5 hours.

3. Effect of CeO₂ loading

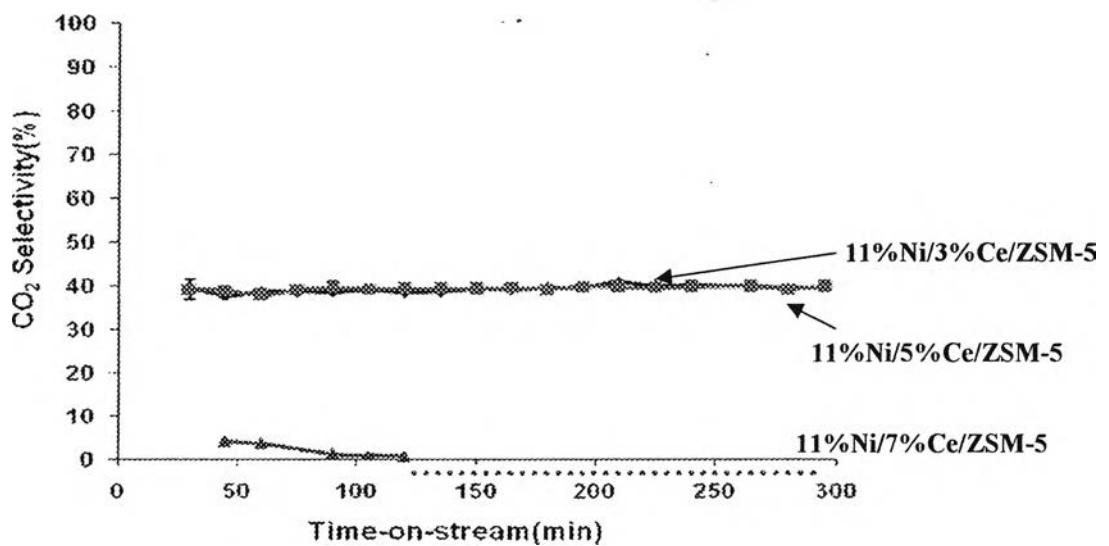


Figure E3 CO₂ selectivity of the 11%Ni/ZSM-5 catalysts with various CeO₂ loadings, which operated on steam reforming reaction at 700°C, H₂O/CH₄ ratio of 0.8, and atmospheric pressure for 5 hours.

Appendix F Experimental data

Table F1 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

3%Ni loading

Table F2 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

5%Ni loading

Table F3 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

7%Ni loading

Table F4 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

9%Ni loading

Table F5 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni loading

Table F6 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

15%Ni loading

Table F7 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni loading at R = 1

Table F8 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni loading at R = 2

Table F9 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni loading at R = 2 (repeat)

Table F10 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni/3%Ce/ZSM-5

Table F11 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

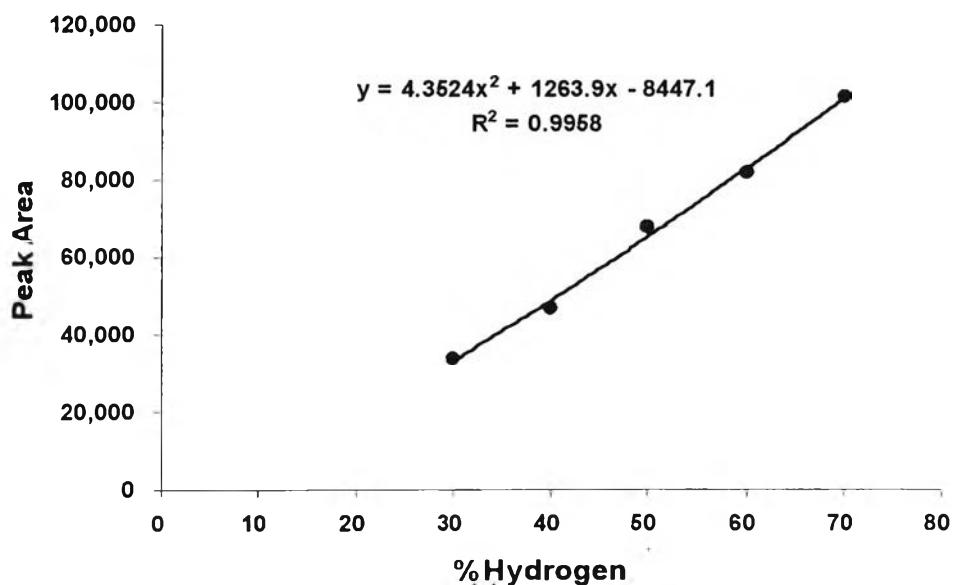
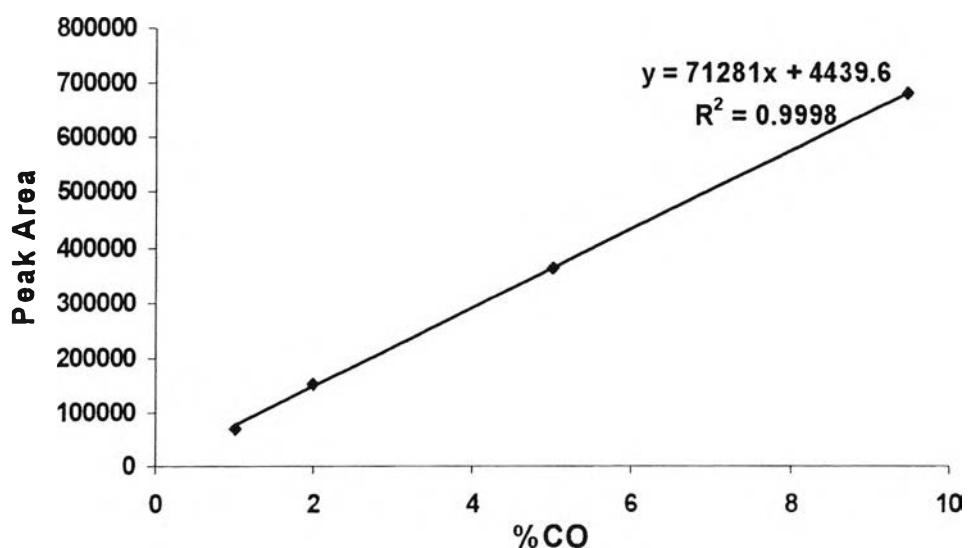
11%Ni/5%Ce/ZSM-5

Table F12 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni/5%Ce/ZSM-5 (repeat)

Table F13 Effect of Ni loading: H₂ selectivity, CO selectivity, CO₂ selectivity, H₂ yield, CO yield, CH₄ conversion

11%Ni/7%Ce/ZSM-5

Appendix G Calibration curves for product gases**Figure G1** H₂ calibration curve.**Figure G2** CO calibration curve.

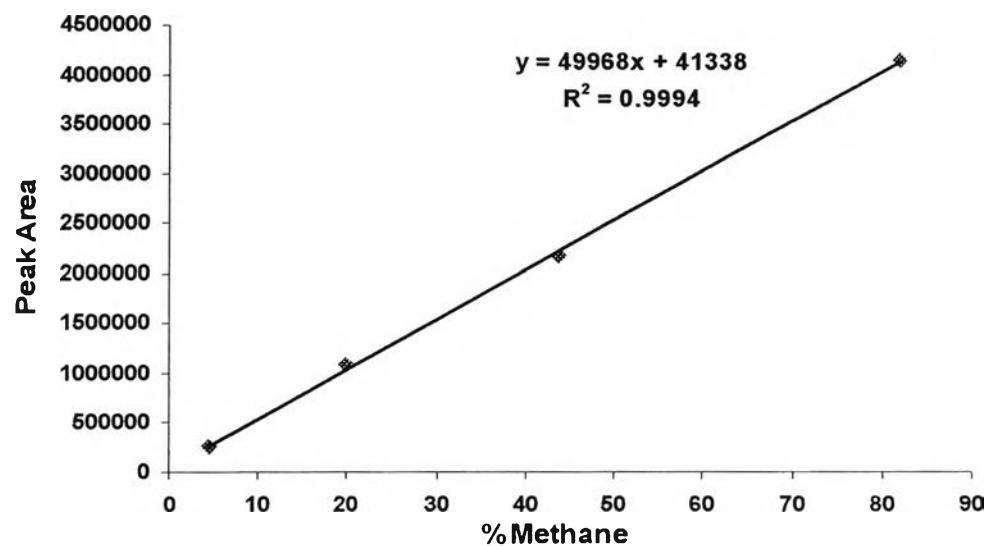


Figure G3 CH₄ calibration curve.

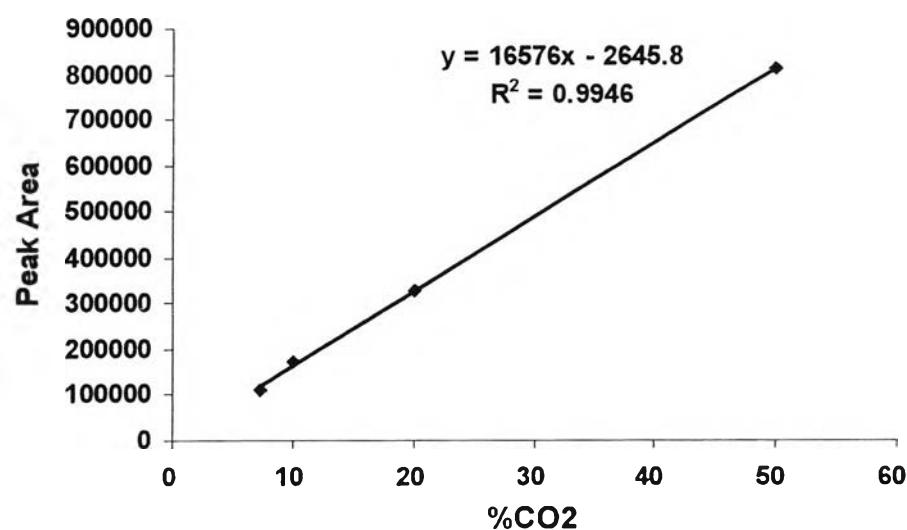


Figure G4 CO₂ calibration curve.

CURRICULUM VITAE

Name: Ms. Wanwanat Noisra

Date of Birth: March 6, 1983

Nationality: Thai

University Education:

2002-2006 Bachelor Degree of Chemical Engineering, Faculty of Engineering, Mahidol University, Nakornprathom, Thailand

Working Experience:

2005-2005	Position:	Quality control intern
	Company name:	PTT Co.Ltd.

Presentations:

1. Noisra, W., Luengnaruemitchai, A., and Jitkarnka, S. (2007, October 29-30) Hydrogen Production from the Steam Reforming of Methane over Ni Supported on ZSM-5 Zeolite Catalysts at Thailand Chemical Engineering & Applied Chemistry Conference 2007, Chiang Mai, Thailand.

